

Name _____

Solution Stoichiometry Worksheet

Solve the following solutions Stoichiometry problems:

1. How many grams of silver chromate will precipitate when 150. mL of 0.500 M silver nitrate are added to 100. mL of 0.400 M potassium chromate?



$$\frac{0.150 \text{ L AgNO}_3}{1 \text{ L}} \left| \frac{0.500 \text{ moles AgNO}_3}{1 \text{ L}} \right| \left| \frac{1 \text{ moles Ag}_2\text{CrO}_4}{2 \text{ moles AgNO}_3} \right| \left| \frac{331.74 \text{ g Ag}_2\text{CrO}_4}{1 \text{ moles Ag}_2\text{CrO}_4} \right| = 12.4 \text{ g Ag}_2\text{CrO}_4$$

$$\frac{0.100 \text{ L K}_2\text{CrO}_4}{1 \text{ L}} \left| \frac{0.400 \text{ moles K}_2\text{CrO}_4}{1 \text{ L}} \right| \left| \frac{1 \text{ moles Ag}_2\text{CrO}_4}{1 \text{ moles K}_2\text{CrO}_4} \right| \left| \frac{331.74 \text{ g Ag}_2\text{CrO}_4}{1 \text{ moles Ag}_2\text{CrO}_4} \right| = 13.3 \text{ g Ag}_2\text{CrO}_4$$

2. How many mL of 0.280 M barium nitrate are required to precipitate as barium sulfate all the sulfate ions from 25.0 mL of 0.350 M aluminum sulfate? (93.8 mL barium nitrate)



$$\frac{0.0250 \text{ L Al}_2(\text{SO}_4)_3}{1 \text{ L}} \left| \frac{0.350 \text{ moles Al}_2(\text{SO}_4)_3}{1 \text{ L}} \right| \left| \frac{3 \text{ moles Ba}(\text{NO}_3)_2}{1 \text{ moles Al}_2(\text{SO}_4)_3} \right| \left| \frac{1 \text{ L}}{0.280 \text{ moles Ba}(\text{NO}_3)_2} \right| = 0.0938 \text{ L Ba}(\text{NO}_3)_2$$

3. 25.0 mL of 0.350 M NaOH are added to 45.0 mL of 0.125 M copper (II) sulfate. How many grams of copper (II) hydroxide will precipitate?



$$\frac{0.0250 \text{ L NaOH}}{1 \text{ L NaOH}} \left| \frac{0.350 \text{ moles NaOH}}{1 \text{ L NaOH}} \right| \left| \frac{1 \text{ moles Cu}(\text{OH})_2}{2 \text{ moles NaOH}} \right| \left| \frac{97.57 \text{ g Cu}(\text{OH})_2}{1 \text{ mole Cu}(\text{OH})_2} \right| = 0.427 \text{ g Cu}(\text{OH})_2$$

$$\frac{0.0450 \text{ L CuSO}_4}{1 \text{ L NaOH}} \left| \frac{0.125 \text{ moles CuSO}_4}{1 \text{ L NaOH}} \right| \left| \frac{1 \text{ moles Cu}(\text{OH})_2}{1 \text{ moles CuSO}_4} \right| \left| \frac{97.57 \text{ g Cu}(\text{OH})_2}{1 \text{ mole Cu}(\text{OH})_2} \right| = 0.549 \text{ g Cu}(\text{OH})_2$$

4. What volume of 0.415 M silver nitrate will be required to precipitate as silver bromide all the bromide ion in 35.0 mL of 0.128 M calcium bromide?



$$\frac{0.0350 \text{ L CaBr}_2}{1 \text{ L CaBr}_2} \left| \frac{0.128 \text{ moles CaBr}_2}{1 \text{ L CaBr}_2} \right| \left| \frac{2 \text{ moles AgNO}_3}{1 \text{ moles CaBr}_2} \right| \left| \frac{1 \text{ L AgNO}_3}{0.415 \text{ mole AgNO}_3} \right| = 0.0216 \text{ L AgNO}_3$$

5. What volume of 0.496 M HCl is required to neutralize 20.0 mL of 0.809 M sodium hydroxide?



$$\frac{0.0200 \text{ L NaOH}}{1 \text{ L NaOH}} \left| \frac{0.809 \text{ mole NaOH}}{1 \text{ L NaOH}} \right| \left| \frac{1 \text{ mole HCl}}{1 \text{ mole NaOH}} \right| \left| \frac{1 \text{ L HCl}}{0.496 \text{ mole HCl}} \right| = 0.0326 \text{ L HCl}$$

Chapter 121 Stoichiometry Answer Key

Michael Brown



Chapter 121 Stoichiometry Answer Key

Chemistry 121 Worksheet Answers Steve DeRoy,2012 **General Chemistry for Engineers** James O. Glanville,2002 Emphasizing problem solving and engineering approximation this chemistry book provides engineers with an understanding of the entities atoms molecules and ions that are relevant to their lives and professional careers Throughout the book internet key word searching and graphing exercises take advantage of users existing computer skills and encourages them to acquire new ones in designing preparing and interpreting graphs Chapter topics cover atoms elements and measurements nuclides molecules and ions chemical reaction and stoichiometry gases quantum mechanics and the periodic table chemical bonding and chemical structure chemical energy and the first law of thermodynamics the second law of thermodynamics and chemical equilibrium gas and solution equilibria liquids and their mixtures solids phase diagrams and solutions the periodic table and redox chemistry electrochemistry and rate processes For engineers preparing for the professional certification exam *A Problem-Solving Approach to Aquatic Chemistry* James N. Jensen,2003-03-13 This text provides a detailed introduction to aquatic equilibrium chemistry calculation methods for systems at equilibrium applications of aquatic chemistry and chemical kinetics Software designed especially for the text allows the reader to build complex models by applying equilibrium calculation principles Important features include material specific and integrated case studies thought provoking questions key ideas and historical sketches **Modern Analytical Chemistry** David Harvey,2000 This introductory text covers both traditional and contemporary topics relevant to analytical chemistry Its flexible approach allows instructors to choose their favourite topics of discussion from additional coverage of subjects such as sampling kinetic method and quality assurance **Investigating Chemistry** Matthew Johll,2006-03-17 Matthew Johll s book introduces students from a non science background to the fundamentals of chemistry through an array of examples and applications from real life crime scenes Sherlock Holmes stories and authentic accounts of drug deals murders and thefts

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Fundamentals of Chemistry: Laboratory Studies Frank Brescia, 2012-12-02 Fundamentals of Chemistry Laboratory Studies Third Edition is a manual that provides instruction on techniques of chemical laboratory operations Each experiment is discussed in terms of the major objective the experimental approach to the objective the measurements or observations to be made and the calculation and interpretation of results Topics covered include manipulation weights and measures molecular weight acids and bases gravimetric and volumetric stoichiometry and thermochemistry This book is comprised of 43 chapters divided into 14 sections and begins by presenting general information on metric and other units common laboratory equipment and chemical laboratory methods The first chapter introduces the reader to the Bunsen burner and the principles of glass working followed by a discussion on mass and volume measurements including the determination of density The following chapters focus on states of matter molecular weight stoichiometry and intermolecular forces Preparations and syntheses are also considered along with chemical equilibrium and electrochemistry The final section is devoted to qualitative analysis particularly of cations and anions This monograph is intended primarily for students of

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LAKSHMIKUTTY,2006-01-01 This textbook is designed for undergraduate courses in chemical engineering and related disciplines such as biotechnology polymer technology petrochemical engineering electrochemical engineering environmental engineering safety engineering and industrial chemistry The chief objective of this text is to prepare students to make analysis of chemical processes through calculations and also to develop in them systematic problem solving skills The students are introduced not only to the application of law of combining proportions to chemical reactions as the word stoichiometry implies but also to formulating and solving material and energy balances in processes with and without chemical reactions The book presents the fundamentals of chemical engineering operations and processes in an accessible style to help the students gain a thorough understanding of chemical process calculations It also covers in detail the background materials such as units and conversions dimensional analysis and dimensionless groups property estimation P V T behaviour of fluids vapour pressure and phase equilibrium relationships humidity and saturation With the help of examples the book explains the construction and use of reference substance plots equilibrium diagrams psychrometric charts steam tables and enthalpy composition diagrams It also elaborates on thermophysics and thermochemistry to acquaint the students with the thermodynamic principles of energy balance calculations Key Features SI units are used throughout the book Presents a thorough introduction to basic chemical engineering principles Provides many worked out examples and exercise problems with answers Objective type questions included at the end of the book serve as useful review material and also assist the students in preparing for competitive examinations such as GATE

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