

Concentration – Understanding Molarity and Parameters affecting Molarity Remote Lab

This lab uses the [Concentration](https://phet.colorado.edu/sims/html/concentration/latest/concentration_en.html) and [Molarity](https://phet.colorado.edu/sims/html/molarity/latest/molarity_en.html) simulations from PhET Interactive Simulations at University of Colorado Boulder, under the CC-BY 4.0 license.

https://phet.colorado.edu/sims/html/concentration/latest/concentration_en.html

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Learning Goals

- Describe the relationships between volume and amount of solute to solution concentration.
- Predict how solution concentration will change for any action (or combination of actions) that adds or removes water, solute, or solution, and explain why.

Develop your understanding: Explore the [Concentration](https://phet.colorado.edu/sims/html/concentration/latest/concentration_en.html) simulation.

1. Find 2 ways you can change the concentration of the solution in the beaker and list them below.



Explain your understanding: moles = grams/molar mass

$$\text{Grams} = \text{moles} \times \text{molar mass}$$

1. Mole calculation. Show your work:
 - a. How many moles is 55 g of NaCl?

0.025 g of NaCO₃?
 - b. How many grams is 0.5 moles of NaCl?

2.11 moles of NaCO₃?
2. There are several ways to measure Concentration. The first simulation, called concentration, uses Molarity (mol/L).
 - a. Describe in your own words how mol/L compares to grams/liter. (similarities and differences)

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Richard Bailey



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