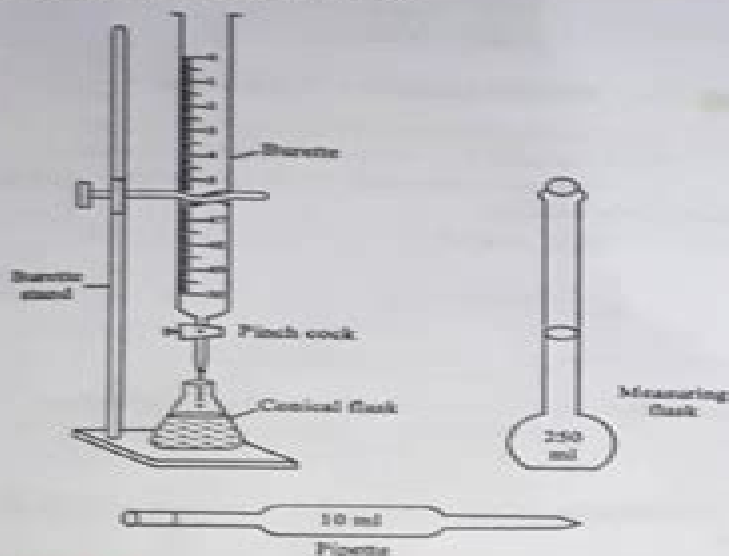


### EXPERIMENT NO. 8-1

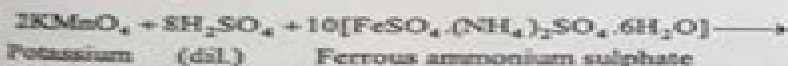
**Object :** Prepare 250 ml standard solution of N/20 Ferrous ammonium sulphate and find out the strength in gram/litre of Potassium permanganate solution of unknown concentration by oxidation-reduction (Redox) single titration.

**Required materials :** Ferrous Ammonium Sulphate, Potassium Permanganate, H<sub>2</sub>SO<sub>4</sub>, Burette, Pipette, Burette stand, Chemical balance, Weight box, Funnel, Test-tube, Measuring flask, Watch glass, Marble tile, Conical flask etc.

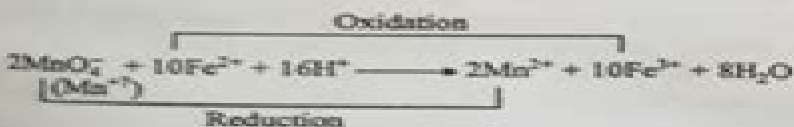


**Principle :** Acidic KMnO<sub>4</sub> acts as a strong oxidising agent and Ferrous ammonium sulphate (Mohr's salt) acts as a strong reducing agent. When KMnO<sub>4</sub> solution is added to a definite volume of Ferrous ammonium sulphate in the presence of dilute H<sub>2</sub>SO<sub>4</sub>, then Ferrous sulphate oxidises to Ferric sulphate. A light and stable pink colour is produced in the solution at the end point. Here KMnO<sub>4</sub> acts as a self indicator.

**Chemical Equation :**



In Ionic form :



**Procedure :** (I) Preparation of standard solution of Ferrous ammonium sulphate of N/20 concentration :

Required quantity of FAS (W) =

$$\frac{\text{Equivalent mass (E)} \times \text{Normality (N)} \times \text{Volume (V)}}{1000}$$

Equivalent mass of FAS (E) = 392 gm,

Normality of FAS (N) =  $\frac{1}{20}$

Volume of FAS (V) = 250 ml.

$$\text{Required quantity of FAS (W)} = \frac{392 \times \frac{1}{20} \times 250}{1000} = 4.9 \text{ gm.}$$

**Observation table :**

S. No.	Reading of Pipette (ml)	Reading of Burette		Volume of used KMnO <sub>4</sub> (ml)	Concordant volume of used KMnO <sub>4</sub>
		Initial reading	Final reading		
1.	10 ml	0.0 ml	9.5 ml	9.7 ml	9.5 ml
2.	10 ml	9.5 ml	19.2 ml	9.5 ml	
3.	10 ml	19.2 ml	28.7 ml	9.5 ml	

**Calculation :** According to the law of equivalence :

$$N_1 V_1 = N_2 V_2$$

Where, N<sub>1</sub> = Normality of known FAS solution =  $\frac{1}{20}$

V<sub>1</sub> = Volume of known FAS solution = 10 ml

N<sub>2</sub> = Normality of unknown KMnO<sub>4</sub> solution = ?

V<sub>2</sub> = Concordant volume of unknown KMnO<sub>4</sub> = 9.5 ml.

By putting the values in,

$$N_1 V_1 = N_2 V_2$$

$$= \frac{1}{20} \times 10 = N_2 \times 9.5$$

$$N_2 = \frac{10}{20 \times 9.5} = 0.052 \text{ N.}$$

Concentration of unknown KMnO<sub>4</sub> solution in gram/litre

$$= \text{Normality (N}_2\text{)} \times \text{Equivalent mass of KMnO}_4$$

$$= 0.052 \times 31.6$$

$$= 1.64 \text{ gm/litre.}$$

Weigh exactly 4.9 gm Ferrous ammonium sulphate in a clean and dry watch glass and transfer it to a 250 ml capacity measuring flask by a dry funnel, in which a little distilled water is already taken, flask is shaken to dissolve the entire crystalline substance in water, level of water is made upto the mark. The solution should be homogenous.

(II) Titration : Firstly burette is cleaned by rinsing it with water and then with KMnO<sub>4</sub> solution. Then unknown KMnO<sub>4</sub> solution is filled in the burette upto the zero mark and initial reading is noted. A clean pipette is rinsed with FAS solution once, then 10 ml of standard FAS solution is taken in this pipette and poured into a clean conical flask. 5 ml of dilute H<sub>2</sub>SO<sub>4</sub> is mixed in it. Conical

# **Volumetric Analysis Chemistry Practical**

**Harold William Bausor**



# Volumetric Analysis Chemistry Practical

**Practical Volumetric Analysis** Peter McPherson, 2014-09-17 Written by someone who has experienced both teaching and working as a research chemist this textbook will provide the theoretical chemistry associated with volumetric analysis supported by a selection of practicals for undergraduate students taking modules in introductory and analytical chemistry as well as for non specialists teaching chemistry **Practical Chemistry for 'A' Level** Kum Wone Kam, Seet Ngoh Kwan, 1989

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New Reduction Methods in Volumetric Analysis Edmund Knecht, Eva Hibbert, 1910 *Volumetric Analysis* Chirag Fultariya, Jalpa Harsora, 2017-03-04 This book is written for the entry level chemistry students It is most important that students understand the basics behind every experiment they perform to excel in future This book explains the basic principles and practical methods in possible simple way to quantify the concentration of an analyte **Practical Chemistry, Including Simple Volumetric Analysis and Toxicology, Etc** Percy Andrew Ellis RICHARDS, 1904 *Senior Practical Chemistry* Harold William Bausor, 2015-06-25 Excerpt from Senior Practical Chemistry This little book provides a course of experimental work which is designed to meet the requirements of the Senior Cambridge Local Examination in Practical Chemistry It is divided into three parts In Part I Chapter I a number of carefully selected preparations are described with full working details and diagrams where necessary in Chapter II the action of heat on some typical substances is investigated Part II deals with Quantitative Analysis Chapter I consisting of a selection of simple experiments mainly gravimetric whilst Chapter II contains an elementary treatment of Volumetric Analysis Part III is concerned with the Qualitative Analysis of Simple Salts The syllabus of the Examination however states that alternative questions will be set so that a candidate may avoid the Qualitative Analysis altogether if he wishes No apologies are needed for writing a text book to a particular syllabus provided it is a good syllabus As this condition is certainly satisfied the author hopes that the book will be found useful by a wider range of students than those for whom he is ostensibly writing We have to thank the Controller of His Majesty's Stationery Office and also Messrs Macmillan and Co for their kind permission to include in this book the Tables of Logarithms and Antilogarithms published in Examinations in Science and Technology About the Publisher Forgotten Books publishes hundreds of thousands of rare and classic books Find more at [www.forgottenbooks.com](http://www.forgottenbooks.com) This book is a reproduction of an important historical work Forgotten Books uses state of the art technology to digitally reconstruct the work preserving the original format whilst repairing imperfections present in the aged copy In rare cases an imperfection in the original such as a blemish or missing page may be replicated in our edition We do however repair the vast majority of imperfections successfully any imperfections that remain are intentionally left to preserve the state of such historical works **College Practical Chemistry** V K Ahluwalia, Sunita Dhingra, Sunita Dhingra, 2005 *Integrated and Simplified Approaches To*

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investigating rather than simply performing a laboratory experiment to the specifications set by the manual Each experiment begins with a problem scenario and ends with questions requiring feedback on the problem **Analytical Chemistry** E. Hywel Evans, Mike E. Foulkes, 2019 Analytical Chemistry A Practical Approach is the only chemical analysis text with an emphasis on active learning giving students step by step guidance on how the key principles of analytical science are applied in a range of practical real world contexts *Volumetric Analysis* Kum Wone Kam, 1978 **Volumetric Analysis for Students of Pharmaceutical and General Chemistry** Charles Herbert Hampshire, 1912 NEW REDUCTION METHODS IN VOLUM Edmund 1861-1925 Knecht, Eva Hibbert, 2016-08-26 This work has been selected by scholars as being culturally important and is part of the knowledge base of civilization as we know it This work was reproduced from the original artifact and remains as true to the original work as possible Therefore you will see the original copyright references library stamps as most of these works have been housed in our most important libraries around the world and other notations in the work This work is in the public domain in the United States of America and possibly other nations Within the United States you may freely copy and distribute this work as no entity individual or corporate has a copyright on the body of the work As a reproduction of a historical artifact this work may contain missing or blurred pages poor pictures errant marks etc Scholars believe and we concur that this work is important enough to be preserved reproduced and made generally available to the public We appreciate your support of the preservation process and thank you for being an important part of keeping this knowledge alive and relevant

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