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## Chapter 9

### PERIODIC TRENDS

#### Exercises

- 9.2 (a) The similarities in properties between the same oxidation state of an element of Group ( $n$ ) and Period ( $m$ ) with the element of Group ( $n+2$ ) and Period ( $m+1$ ).
- (b) A polyatomic ion that behaves very similarly to that of an ion of an element or groups of elements.

- 9.4 They both form anhydrous chlorides that react with water to release hydrogen chloride gas while their hexahydrates are water soluble. The anhydrous chlorides both function as Friedel-Craft catalysts. In the gas phase, their chlorides exist as dimers,  $Al_2Cl_6$  and  $Fe_2Cl_6$ , respectively. On the other hand, iron(III) oxide is basic, while the oxide of aluminum is amphoteric. Also, iron(III) compounds are colored, while those of aluminum are white.

- 9.6 KH (d.  $417^\circ\text{C}$ );  $CaH_2$  (m.p.  $116^\circ\text{C}$ );  $GaH_3$  (m.p.  $-15^\circ\text{C}$ );  $GeH_4$  (m.p.  $-165^\circ\text{C}$ );  $AsH_3$  (m.p.  $116^\circ\text{C}$ );  $SeH_2$  (m.p.  $-86^\circ\text{C}$ );  $HBr$  (m.p.  $-89^\circ\text{C}$ ) [information from [www.webelements.com](http://www.webelements.com)].

The bonding in the potassium and calcium hydrides is ionic, while that for the germanium, arsenic, selenium, bromine, and krypton compounds is covalent. The gallium compound does not fit the pattern of a network covalent between the ionic and small molecule covalent. It is too low to be ionic and not low enough to match the others in the period. In fact,  $GaH_3$  is a dimer,  $Ga_2H_6$  (analogous to diborane). Thus it is a small molecule covalent species, but with a doubled molecular mass and number of electrons, it has a higher melting point.

# **Descriptive Inorganic Chemistry Solutions Manual Pdf**

**Gary Wulfsberg**



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