CHEMISTRY: LIMITING REAGENT Practice Problems

Please show your work on a separate piece of paper!

Assume standard temperature and pressure, STP conditions for all reactions.

At high temperatures, sulfur combines with iron to form the brown-black iron (II) sulfide:

$$Fe_{(0)} + S_{(0)} \rightarrow FeS_{(0)}$$

In one experiment, 7.62 g of Fe are allowed to react with 8.67 g of S. a. What is the limiting reagent, and what is the reactant in excess?

b. Calculate the mass of FeS formed.

 Arcylonitrile, C 3H3N, is the starting material for the production of a kind of synthetic fiber (acrylics) and can be made from propylene, C3H6, by reaction with nitric oxide, NO, as follows:

$$4 C_3H_{6 (g)} + 6 NO_{(g)} \rightarrow 4 C_3H_3N_{(s)} + 6 H_2O_{(l)} + N_{2 (g)}$$

What mass of C₃H₃N can be made when 21.6 g of C₃H₆ react with 21.6 g of nitric oxide?

- Calculate the percent yield for the reaction: P_{4 to} + 6 Cl_{2 tg} → 4 PCl_{3 (t)} if 75.0 g of phosphorus reacts with excess chlorine gas to produce 111.0 g of phosphorus trichloride.
- 4. Formic acid, HCHO2, burns in oxygen to form carbon dioxide and water as follows:

If a 3.15-g sample of formic acid was burned in 2.0 L of oxygen, what volume of carbon dioxide would be produced?

Zinc metal reacts with hydrochloric acid to produce zinc chloride and hydrogen gas.
 a. Balance the following reaction: Zn_(i) + HCl_(aq) → ZnCl_{2 (aq)} + H_{2 (g)}
 b. A 3.50-g sample of zinc metal is allowed to react with 2.50 g of hydrochloric acid.

Complete the following table:

reactants/products	Zn (grams)	HCl (grams)	ZnCl ₂ (grams)	$H_2(L)$
Before reaction				
After reaction	1.26 g			

Limiting Reagent Problems Answers

P.S. Raghavan

Limiting Reagent Problems Answers

Green Organic Chemistry Kenneth M. Doxsee, James E. Hutchison, 2004 This lab text describes the tools and strategies of green chemistry and the lab experiments that allow investigation of organic chemistry concepts and techniques in a greener laboratory setting Students acquire the tools to assess the health and environmental impacts of chemical processes and the strategies to improve develop new processes that are less harmful to human health and the environment The curriculum introduces a number of state of the art experiments and reduces reliance on expensive environmental controls such as fume Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan hoods Provided by publisher Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields Dr. Salm's Chemistry Problem <u>Drill Book Robert J. Salm, Robert J. Salm Robert J. Salm Ph.D., 2005 This drill book contains many common problem types that</u> are asked in General Chemistry classes in High School and College This work will give you practice with the major problem types as you prepare for finals and standardized tests U Can: Chemistry I For Dummies John T. Moore, Chris Hren, Peter J. Mikulecky, 2015-08-10 Now you can score higher in chemistry Every high school requires a course in chemistry for graduation and many universities require the course for majors in medicine engineering biology and various other sciences U Can Chemistry I For Dummies offers all the how to content you need to enhance your classroom learning simplify complicated topics and deepen your understanding of often intimidating course material Plus you ll find easy to follow examples and hundreds of practice problems as well as access to 1 001 additional Chemistry I practice problems online As more and more students enroll in chemistry courses the need for a trusted and accessible resource to aid in study has never been greater That's where U Can Chemistry I For Dummies comes in If you re struggling in the classroom this hands on

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STOICHIOMETRY AND PROCESS CALCULATIONS K. V. NARAYANAN, B. LAKSHMIKUTTY, 2006-01-01 This textbook is designed for undergraduate courses in chemical engineering and related disciplines such as biotechnology polymer technology petrochemical engineering electrochemical engineering environmental engineering safety engineering and industrial chemistry. The chief objective of this text is to prepare students to make analysis of chemical processes through calculations and also to develop in them systematic problem solving skills The students are introduced not only to the application of law of combining proportions to chemical reactions as the word stoichiometry implies but also to formulating and solving material and energy balances in processes with and without chemical reactions The book presents the fundamentals of chemical engineering operations and processes in an accessible style to help the students gain a thorough understanding of chemical process calculations It also covers in detail the background materials such as units and conversions dimensional analysis and dimensionless groups property estimation P V T behaviour of fluids vapour pressure and phase equilibrium relationships humidity and saturation With the help of examples the book explains the construction and use of reference substance plots equilibrium diagrams psychrometric charts steam tables and enthalpy composition diagrams It also elaborates on thermophysics and thermochemistry to acquaint the students with the thermodynamic principles of energy balance calculations Key Features SI units are used throughout the book Presents a thorough introduction to basic chemical engineering principles Provides many worked out examples and exercise problems with answers Objective type questions included at the end of the book serve as useful review material and also assist the students in preparing for competitive examinations such as GATE Chemistry All-in-One For Dummies (+ Chapter Quizzes Online) Christopher Hren, John T. Moore, Peter J. Mikulecky, 2022-11-23 Everything you need to crush chemistry with confidence Chemistry All in One For Dummies arms you with all the no nonsense how to content you ll need to pass your chemistry class with flying colors You ll find tons of practical examples and practice problems and you ll get access to an online guiz for every chapter Reinforce the concepts you learn in the classroom and beef up your understanding of all the chemistry topics covered in the standard curriculum Prepping for the AP Chemistry exam Dummies has your back with plenty of review before test day With clear definitions concise explanations and plenty of helpful information on everything from matter and molecules to moles and measurements Chemistry All in One For Dummies is a one stop resource for chem students of all valences Review all the topics covered in a full year high school chemistry course or one semester of college

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