

Thermal Energy- ANSWER KEY

1. A 10.0 g piece of copper wire, sitting in the sun reaches a temperature of 80.0°C. How many Joules are released when the copper cools to 40.0°C? The specific heat capacity of copper is 0.377 J/(g.°C).

First: write down your givens

Mass (m) = 10 g

Initial temperature (T_i) = 80°C

Final temperature (T_f) = 40 °C

Specific heat capacity = 0.377 J/(g.°C)

The formula to find heat (Q) (or thermal energy) = $m c \Delta T$

Remember that $\Delta T = T_f - T_i$

Next: Plug your givens in your formula

$$Q = (10) (0.377) (80 - 40)$$

$$Q = 150.8 \text{ Joules}$$

2. The specific heat capacity of water is 4.184 J/(g.°C). How much thermal energy is required to change the temperature of 700.0g of water from 25.6°C to 75.4 °C?

Givens: $c = 4.184 \text{ J/(g.°C)}$

$m = 700 \text{ g}$

$T_f = 75.4^\circ\text{C}$

$T_i = 25.6^\circ\text{C}$

The formula for thermal energy is

$$Q = m c \Delta T$$

$$Q = (700) (4.184) (75.4 - 25.6)$$

$$Q = 145,854.24 \text{ Joules}$$

Supplemental Problems Thermal Energy Answer Key

DP Hallahan



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