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ec on the basis of this experiment what is the heat of solution per mole of KNO_3 if c_p of H_2O is $4.184 \text{ J/g}^\circ\text{C}$ and the specific heat of water is $4.184 \text{ J/g}^\circ\text{C}$ soln 25.8 g 98.57 g 101.15 g ΔT cal 20.4°C 22.5°C 21°C

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web in a calorimeter the white solid BaSO_4 forms and the temperature of the mixture increases to 28.1°C assuming that the calorimeter absorbs only a negligible quantity of heat and that the specific heat capacity of the solution is $4.18 \text{ J/g}^\circ\text{C}$ and that the density of the final solution is 1.0 g/ml calculate the enthalpy change per mole

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