

Worksheet 15 - Molecular Shapes

The shapes of molecules can be predicted from their **Lewis structures** by using the **VSEPR (Valence Shell Electron Pair Repulsion)** model, which states that electron pairs around a central atom will assume a geometry that keeps them as far apart from each other as possible.

This is illustrated by the drawings below.



Six groups surrounding a central atom will form an **octahedron**. All of the groups in this structure are at 90° or 180° to each other. All positions are equivalent.



Five groups will form a **trigonal bipyramid**. The two positions pointing up and down are called the **axial** positions. They are at 180° to each other, and at 90° to the other three, **equatorial** positions. The three **equatorial** positions are at 120° to each other. There is more room in the equatorial positions, and large groups will occupy these positions.



Four groups will form a **tetrahedron**. All of the angles in a tetrahedron are 109.5° , and all positions are equivalent.



Three groups will form a flat triangle (**trigonal planar**). Each of the angles is 120° and all positions are equivalent.



Two groups form a straight line (**linear**) with 180° between them.

How does this apply to Chemistry?

The groups occupying these geometric positions will be either **atoms** bonded to the central atom, or **lone pair electrons** on the central atom.

Lone pair electrons occupy **more** space than bonded electrons, so they will take the **equatorial** position in the **trigonal bipyramid**.

Lone pair electrons will also occupy positions that put them as far apart from each other as possible.

Worksheet 15 Molecular Shapes

Elissa Huddart



Worksheet 15 Molecular Shapes

Chemistry 2e Paul Flowers, Richard Langely, William R. Robinson, Klaus Hellmut Theopold, 2019-02-14 Chemistry 2e is designed to meet the scope and sequence requirements of the two semester general chemistry course. The textbook provides an important opportunity for students to learn the core concepts of chemistry and understand how those concepts apply to their lives and the world around them. The book also includes a number of innovative features including interactive exercises and real world applications designed to enhance student learning. The second edition has been revised to incorporate clearer, more current and more dynamic explanations while maintaining the same organization as the first edition. Substantial improvements have been made in the figures, illustrations, and example exercises that support the text narrative. Changes made in Chemistry 2e are described in the preface to help instructors transition to the second edition. *Atoms, Molecules & Elements: What Are Compounds? Gr. 5-8* George Graybill, 2015-10-01 This is the chapter slice What Are Compounds from the full lesson plan Atoms Molecules Elements. Young scientists will be thrilled to explore the invisible world of atoms, molecules, and elements. Our resource provides ready-to-use information and activities for remedial students using simplified language and vocabulary. Students will label each part of the atom, learn what compounds are, and explore the patterns in the periodic table of elements to find calcium (Ca), chlorine (Cl), and helium (He) through hands-on activities. These and more science concepts are presented in a way that makes them more accessible to students and easier to understand. Written to grade and using simplified language and vocabulary, and comprised of reading passages, student activities, crossword, word search, comprehension quiz, and color mini posters, our resource can be used effectively for test prep and your whole class. All of our content is aligned to your State Standards and are written to Bloom's Taxonomy and STEM initiatives. *Atoms, Molecules & Elements: What Are Elements? Gr. 5-8* George Graybill, 2015-10-01 This is the chapter slice What Are Elements from the full lesson plan Atoms Molecules Elements. Young scientists will be thrilled to explore the invisible world of atoms, molecules, and elements. Our resource provides ready-to-use information and activities for remedial students using simplified language and vocabulary. Students will label each part of the atom, learn what compounds are, and explore the patterns in the periodic table of elements to find calcium (Ca), chlorine (Cl), and helium (He) through hands-on activities. These and more science concepts are presented in a way that makes them more accessible to students and easier to understand. Written to grade and using simplified language and vocabulary, and comprised of reading passages, student activities, crossword, word search, comprehension quiz, and color mini posters, our resource can be used effectively for test prep and your whole class. All of our content is aligned to your State Standards and are written to Bloom's Taxonomy and STEM initiatives. **Chemistry** Theodore Lawrence Brown, H. Eugene LeMay, Bruce E. Bursten, Patrick Woodward, Catherine Murphy, 2017-01-03 NOTE This edition features the same content as the traditional text in a convenient three-hole punched loose-leaf version. Books a la

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Chemistry Bruce Averill, Patricia Eldredge, 2007 Emphasises on contemporary applications and an intuitive problem

solving approach that helps students discover the exciting potential of chemical science This book incorporates fresh applications from the three major areas of modern research materials environmental chemistry and biological science

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