

3. Find the molarity of a solution made from 275 g of CuSO_4 dissolved in enough water to make 4.25 L.

$$\frac{\text{mol}}{\text{dm}^3}$$

$$\frac{1.72 \text{ mol}}{4.25 \text{ dm}^3}$$

$$\text{conc.} = \frac{.405 \text{ mol}}{\text{dm}^3}$$

$$1 \text{ L} = 1 \text{ dm}^3$$

$$\begin{aligned} \text{Cu} &: 63.5 \\ \text{S} &: 32.0 \\ 4 \times \text{O} &: 64.0 \\ \hline &: 169.5 \text{ g/mol} \end{aligned}$$

$$\begin{aligned} 169.5 \text{ g} &= 1 \text{ mol} \\ \frac{275 \text{ g}}{169.5 \text{ g}} &= \frac{x \text{ mol}}{1 \text{ mol}} \\ x &= 1.72 \text{ mol} \\ &\text{CuSO}_4 \end{aligned}$$

4. An alcoholic iodine solution ("tincture" of iodine) is prepared by dissolving 5.15 g of iodine crystals in enough alcohol to make a volume of 225 mL. Calculate the molarity of iodine in the solution.

$$\frac{5.15 \text{ g I}_2}{254 \text{ g I}_2} = \frac{x \text{ mol}}{1 \text{ mol}}$$

$$x = .0203 \text{ mol I}_2$$

$$\frac{\text{mol}}{\text{dm}^3}$$

$$225 \text{ mL} = 225 \text{ cm}^3 =$$

Solution Chemistry Practice Problems

**David Bookin, Darrell D. Ebbing, Steven
D. Gammon**



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