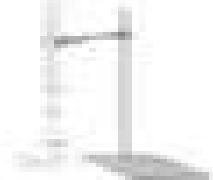




AP Chem Investigation 12 – Redox Titration Lab



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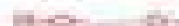
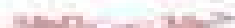
- Check to be sure you have at least 20 mL of $K_2Cr_2O_7$ in your buret. If not, add some $K_2Cr_2O_7$ to your buret.
- Be sure there are not air bubbles in the buret. If there are, stretch some $K_2Cr_2O_7$ out of the buret into a waste beaker until the air bubble comes out. You can dump this solution down the sink with running water.
- Take an initial reading of $K_2Cr_2O_7$ from the buret. Record this value in the data table.
- Measure out exactly 1.00 mL of H_2O_2 using a graduated pipet and add it to a clean flask.
- Add 20 mL of 1M H_2SO_4 to the flask using a graduated cylinder. This is so that the titration can occur in acidic conditions.
- Add water to the flask to bring the solution up to about 50-60 mL of volume. (This is just to improve your ability to see the color change at the endpoint. Adding water will NOT affect the number of moles of H_2O_2 present in the flask, so it won't affect the volume of titrant required to reach the equivalence point.)
- Titrate the H_2O_2 to the endpoint (until a first good color persists). Try not to add even just a drop of $K_2Cr_2O_7$ past the endpoint! Record the volume of titrant.
- Dump your solution down the sink. Leave the leftover $K_2Cr_2O_7$ in your buret for the next class.
- Check the label on the $K_2Cr_2O_7$ bottle and record the molarity of the $K_2Cr_2O_7$ solution in your data section.

Data Table

Volume of H_2O_2 sampled	1.00 mL
Initial reading of $K_2Cr_2O_7$ volume on buret	19.11 mL
Final reading of $K_2Cr_2O_7$ volume on buret	17.41 mL
Concentration of $K_2Cr_2O_7$ solution on bottle	0.00880 M

Questions

- Balance the equation below using the half-reaction method. Show all work.



- What volume of $K_2Cr_2O_7$ was added to reach the endpoint?

$$17.41 - 19.11 = 1.70 \text{ mL}$$

- Knowing the volume and concentration of $K_2Cr_2O_7$, how many moles of $K_2Cr_2O_7$ were added to reach the endpoint?

$$1.70 \text{ mL} \times (0.00880 \text{ mol/L}) \times (1000 \text{ mL/L}) = 0.015 \text{ mol}$$

- Using the balanced equation, how many moles of H_2O_2 were in your sample? (Remember $K_2Cr_2O_7$ is $2Cr^{3+}$ in the balanced equation.)

$$1.00 \text{ mol } Cr^{3+} \text{ and } 2K_2Cr_2O_7 \text{ is } 2 \text{ mol } H_2O_2 \times 0.015 \text{ mol } K_2Cr_2O_7 = 0.030 \text{ mol } H_2O_2$$

- Convert the moles of H_2O_2 into grams.

$$0.030 \text{ mol } H_2O_2 \times (34.02 \text{ g/mol}) = 1.02 \text{ g}$$

- Imagine that your sample of H_2O_2 was mostly water and contained very little H_2O_2 molecules. If it was mostly water, it would be expected to have a density similar to water. The density of water is 1.00 g/mL. Let's assume the sample had a density of 1.00 g/mL. Using the mass of H_2O_2 calculated in the previous question along with the fact that your sample was exactly 10.00 mL, calculate the percentage of your sample that was H_2O_2 .

$$(1.02 \text{ g} / 10.00 \text{ g}) \times 100\% = 10.2\%$$

- The label on the bottle states that the bottle of H_2O_2 is a 3% solution. Compare this value to your calculation in the previous question. Is the label correct? Why or why not? (The 3% is based on H_2O_2 breaking down over time.)

- If necessary to know the exact volume of...

a) H_2O_2 sample added to the flask? 1.70 mL

b) water added to the flask? 40 mL

c) $K_2Cr_2O_7$ added to the flask? 1.70 mL

Analysis Of Hydrogen Peroxide Prelab Answers

British Standards Institute Staff



Analysis Of Hydrogen Peroxide Prelab Answers

The Analysis of Aqueous Solutions of Hydrogen Peroxide Charles E. Huckaba,1947 The Analysis of Hydrogen Peroxide Solutions BECCO Chemical Division. Food Machinery and Chemical Corporation, Buffalo Electro-Chemical Company,1954 **Theoretical Analysis of the Decomposition of High Strength Hydrogen Peroxide Solutions** Herbert S. Isbin,1947 **Analysis of a Hydrogen Peroxide Solution** George Deckey,1987-01-01 *Analysis and Experiments of Hydrogen Peroxide Vaporizations and Decomposition* J. S. Mok,2003 Analysis of Hydrogen Peroxide in Los Angeles Atmosphere Isaac R. Kaplan,H. Sakugawa,California. Air Resources Board,1987 **Hydrogen Peroxide for Industrial Use. Method for Determination of Hydrogen Peroxide Content** British Standards Institute Staff,1992-04-01 Hydrogen peroxide Industrial Determination of content Safety measures Volumetric analysis Concentration chemical Testing conditions Chemical analysis and testing Samples *Determination of Hydrogen Peroxide in Reactor Moderator Solutions by Flow Injection Analysis (U)* M. J. Whitaker,Savannah River Site (Westinghouse Savannah River Company),1991 **Fundamental Modeling and Analysis for Thermal Decomposition of Hydrogen Peroxide** ,2004 *A Volumetric Method for Titrimetric Analysis of Hydrogen Peroxide* Walter G. Cox,NAVAL SURFACE WEAPONS CENTER SILVER SPRING MD.,1985 This report documents an oxidation reduction scheme for the volumetric analysis of hydrogen peroxide in the approximate range of 30 percent but it is not limited to this range The purpose of the proposed method is to provide a titrimetric noninstrumental method for the analysis of hydrogen peroxide Keywords Hydrogen Peroxide Potassium Dichromate Ferrous Ammonium Sulfate Hexahydrate Titrimetric Analysis Quantitative Analysis Volumetric Analysis With proper care application of the method described in this technical report provides excellent precision Determination of Hydrogen Peroxide in Reactor Moderator Solutions by Flow Injection Analysis ,1991 A flow injection analysis FIA method for the determination of hydrogen peroxide in reactor moderator water was developed and installed at the Savannah River Site SRS Water Quality Laboratory The technique has an analytical range of 0 10 to 2 50 ppm ug mL with a sampling rate of 40 samples per hour The calibration curve is linear with a correlation coefficient of 0 999 and the precision is excellent with relative standard deviations at the 0 50% level for both 0 10 and 2 50 ppm standards When the automated FIA procedure is compared to the manual method it demonstrates a twenty minute reduction in analysis time per sample and the total liquid waste generated per sample analyzed is reduced by roughly 95 mL 6 refs 10 figs 6 tabs *The Stability of Concentrated Solutions of Hydrogen Peroxide* Walter Cecil Schumb,1947 **Analysis of Hydrogen Peroxide in Los Angeles Atmosphere** ,1987 **Instructors Manual to Lab Manual** Ralph Petrucci,William Harwood,Geoffrey Herring,2001 **Standardization of Potassium Permanganate Solution by Sodium Oxalate** Russell Smith McBride,1913 **Green Chemistry Laboratory Manual for General Chemistry** Sally A. Henrie,2015-03-18 Green chemistry involves designing

novel ways to create and synthesize products and implement processes that will eliminate or greatly reduce negative environmental impacts Providing educational laboratory materials that challenge students with the customary topics found in a general chemistry laboratory manual this lab manual enables students to see how green chemistry principles can be applied to real world issues Following a consistent format each lab experiment includes objectives prelab questions and detailed step by step procedures for performing the experiments Additional questions encourage further research about how green chemistry principles compare with traditional more hazardous experimental methods

Modern Analytical Chemistry David Harvey,2000 This introductory text covers both traditional and contemporary topics relevant to analytical chemistry Its flexible approach allows instructors to choose their favourite topics of discussion from additional coverage of subjects such as sampling kinetic method and quality assurance Addison-Wesley Science Insights ,1996

Exploring General Chemistry in the Laboratory Colleen F. Craig, Kim N. Gunnerson,2017-02-01 This laboratory manual is intended for a two semester general chemistry course The procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life This lab manual covers topics such as composition of compounds reactivity stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes By the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science Science Insights ,1999

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Table of Contents Analysis Of Hydrogen Peroxide Prelab Answers

1. Understanding the eBook Analysis Of Hydrogen Peroxide Prelab Answers
 - The Rise of Digital Reading Analysis Of Hydrogen Peroxide Prelab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Analysis Of Hydrogen Peroxide Prelab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms

- Features to Look for in an Analysis Of Hydrogen Peroxide Prelab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Analysis Of Hydrogen Peroxide Prelab Answers
 - Personalized Recommendations
 - Analysis Of Hydrogen Peroxide Prelab Answers User Reviews and Ratings
 - Analysis Of Hydrogen Peroxide Prelab Answers and Bestseller Lists
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 - Adjustable Fonts and Text Sizes of Analysis Of Hydrogen Peroxide Prelab Answers
 - Highlighting and Note-Taking Analysis Of Hydrogen Peroxide Prelab Answers
 - Interactive Elements Analysis Of Hydrogen Peroxide Prelab Answers
 8. Staying Engaged with Analysis Of Hydrogen Peroxide Prelab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Analysis Of Hydrogen Peroxide Prelab Answers
 9. Balancing eBooks and Physical Books Analysis Of Hydrogen Peroxide Prelab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Analysis Of Hydrogen Peroxide Prelab Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Analysis Of Hydrogen Peroxide Prelab Answers

- Setting Reading Goals Analysis Of Hydrogen Peroxide Prelab Answers
- Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Analysis Of Hydrogen Peroxide Prelab Answers
 - Fact-Checking eBook Content of Analysis Of Hydrogen Peroxide Prelab Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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