

1. Balance the following equations



2. Write a balance equation to correspond to each description

a. When solid potassium chlorate is heated, it decomposes to form solid potassium chloride and oxygen gas.



b. When liquid phosphorus trichloride is added to water it reacts to form aqueous phosphorous acid and aqueous hydrochloric acid



3. Write a balanced reaction for each of the following.

a. Lithium metal is burned in air



b. Aluminum metal is added to a solution of copper (II) chloride



c. Manganese (II) nitrate solution is mixed with sodium hydroxide solution



d. The hydrocarbon hexane, C_6H_{14} , is burned in excess oxygen



e. A sample of calcium carbonate is heated



f. Silver nitrate solution is added to potassium chromate solution



g. Hydrogen peroxide decomposes



h. Solid calcium carbonate is reacted with sulfuric acid



i. Ammonium sulfate solution and potassium hydroxide solution are mixed together



j. Strontium hydroxide is mixed with hydrofluoric acid



Ap Chemistry Solutions

Neil D. Jespersen, Pamela Kerrigan



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CliffsNotes AP Chemistry Bobrow Test Preparation Services,2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter s end Two full length practice tests with detailed answer explanations are included in the book

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Ap Chemistry Solutions Introduction

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