

Reviewing Content

42. The solvent is the substance in which the solute is dissolved.
43. Random collisions of the solvent molecules with the solute particles provide enough force to overcome gravity.
44. **solubility:** the amount of a substance that dissolves in a given quantity of solvent at specified conditions of temperature and pressure to produce a saturated solution.
saturated solution: a solution containing the maximum amount of solute for a given amount of solvent at a constant temperature and pressure. **unsaturated solution:** a solution that contains less solute than a saturated solution at a given temperature and pressure. **miscible:** describes liquids that dissolve in each other. **immiscible:** describes liquids that are insoluble in each other.
45. Particles of solute crystallize.
46. No, if there were undissolved solute, the excess solute would come out of a supersaturated solution.
47. 5.55×10^2 g AgNO_3
48. Solubility increases with pressure.
49. a. 1.6×10^{-2} g/L
b. 4.7×10^{-2} g/L
50. Dilute and concentrated are relative terms and are not quantitative. Molarity provides the exact number of moles of solute per liter of solution.
51. Molarity is the number of moles of solute dissolved in one liter of solution.
a. 1.5M KCl
b. 5.3×10^{-2} M MgCl_2
52. 2.00×10^2 mol
53. a. 5.0×10^{-2} mol NaCl, 20 g NaCl
b. 1.0 mol KNO_3 , 1.0×10^2 g KNO_3
c. 2.5×10^2 mol CaCl_2 , 2.0 g CaCl_2
54. a. 2.3×10^2 g NaCl
b. 2.0 g MgCl_2
55. a. 10% (v/v) ethanol
b. 63.6% (v/v) isopropyl alcohol
56. Colligative properties are properties of a solution that depend only on the number of solute particles; boiling-point elevation, freezing-point depression, and vapor-pressure lowering. Boiling points are elevated because shells of solvent form around solute particles, reducing the amount of solvent molecules that have sufficient energy to escape the solution; relative to the pure solvent, the amount of energy required to cause vaporization or boiling increases. Solutes disrupt the ordering of the solvent structure, so more kinetic energy must be withdrawn from a solution for it to solidify. This lowers the freezing point of the solution.
57. a. sea water
b. 1.5M KNO_3
c. 0.100M MgCl_2
58. The effective molality of the $\text{Ca}(\text{NO}_3)_2$ solution is 3m. The effective molality of the NaCl solution is 2m.
59. When vapor pressure is lowered relative to pure solvent, more energy must be supplied to reach the boiling point; thus the boiling point is increased relative to pure solvent.
60. The salt lowers the freezing point of the ice-water cooling mixture.
61. 1M solution: 1 mol of solute in 1 L of solution; 1m solution: 1 mol of solute in 1000 g of solvent
62. Add 27.0 g H_2O to 32.0 g $\text{C}_2\text{H}_5\text{OH}$.
63. a. 100.26°C
b. 101.54°C
64. a. -4.40°C
b. -2.2°C
65. a. -1.1°C
b. -0.24°C
c. -1.2°C

Understanding Concepts

66. a. The freezing-point depression is twice as great for solute B; solute B must provide twice as many particles in solution.

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Basic Concepts of Chemistry, Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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Table of Contents Ch 16 Assessment Answer Chemistry

1. Understanding the eBook Ch 16 Assessment Answer Chemistry
 - The Rise of Digital Reading Ch 16 Assessment Answer Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Ch 16 Assessment Answer Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ch 16 Assessment Answer Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ch 16 Assessment Answer Chemistry
 - Personalized Recommendations
 - Ch 16 Assessment Answer Chemistry User Reviews and Ratings

- Ch 16 Assessment Answer Chemistry and Bestseller Lists
- 5. Accessing Ch 16 Assessment Answer Chemistry Free and Paid eBooks
 - Ch 16 Assessment Answer Chemistry Public Domain eBooks
 - Ch 16 Assessment Answer Chemistry eBook Subscription Services
 - Ch 16 Assessment Answer Chemistry Budget-Friendly Options
- 6. Navigating Ch 16 Assessment Answer Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Ch 16 Assessment Answer Chemistry Compatibility with Devices
 - Ch 16 Assessment Answer Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ch 16 Assessment Answer Chemistry
 - Highlighting and Note-Taking Ch 16 Assessment Answer Chemistry
 - Interactive Elements Ch 16 Assessment Answer Chemistry
- 8. Staying Engaged with Ch 16 Assessment Answer Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ch 16 Assessment Answer Chemistry
- 9. Balancing eBooks and Physical Books Ch 16 Assessment Answer Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ch 16 Assessment Answer Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ch 16 Assessment Answer Chemistry
 - Setting Reading Goals Ch 16 Assessment Answer Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ch 16 Assessment Answer Chemistry
 - Fact-Checking eBook Content of Ch 16 Assessment Answer Chemistry
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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