

Experiment 15

ANALYSIS OF VINEGAR

FV 9/25/12

MATERIALS: Four 250 mL Erlenmeyer flasks, two 50 mL burets, liquid funnel, one 150 mL beaker, one 50 mL beaker, 5.00 mL pipet, weighing dish, approximately 0.1 M NaOH, $C_6H_4(COOH)COOK$ (potassium hydrogen phthalate), phenolphthalein, grease pencil.

PURPOSE: The purpose of this experiment is two-fold: (1) to standardize a solution of sodium hydroxide; (2) to use the standardized sodium hydroxide solution to titrate the acetic acid in vinegar to determine its percent by mass in vinegar.

LEARNING OBJECTIVE: By the end of this experiment, the midshipman should be able to demonstrate the following proficiencies:

1. Use pipets and burets correctly.
2. Explain the term hygroscopic.
3. Explain the function and necessary qualities of a primary standard.
4. Perform a titration accurately and precisely.
5. Calculate the molarity and the mass of components in a solution from stoichiometry information.
6. Calculate the percent by mass of a component in a solution.

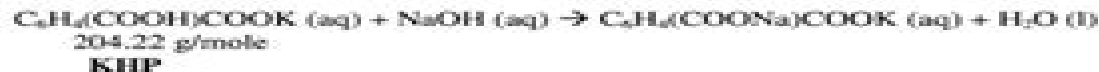
DISCUSSION:

Vinegar is a solution of acetic acid in water. Acetic acid, CH_3COOH , is a weak monoprotic acid with a molar mass of 60.05 g/mole. The percent by mass of acetic acid in vinegar can be determined by titrating a known amount of vinegar with a *standardized* solution of sodium hydroxide, i.e., a sodium hydroxide solution of accurately known concentration. Acetic acid and sodium hydroxide react as shown below:



Sodium hydroxide is a *hygroscopic* solid which means it absorbs water from the air. A weighed quantity of sodium hydroxide therefore contains an unknown mass of water. Therefore, a solution of known molarity cannot be made by dissolving a known mass of solid sodium hydroxide in water. The concentration of a sodium hydroxide solution must be determined experimentally. This is done by titrating the sodium hydroxide solution against a *primary standard*. A primary standard is a substance from which a solution of known concentration can be prepared.

The primary standard used in this experiment is potassium hydrogen phthalate, $C_6H_4(COOH)COOK$, which is referred to by the shorthand notation of "KHP". KHP has several advantages: it does not absorb moisture readily (i.e., is not hygroscopic); it is easily dried; it can be accurately weighed; it can be obtained in very pure form; it has a high molar mass of 204.22 g/mole; and it is very soluble in water. KHP is an acid, which reacts in aqueous solution to neutralize the base, sodium hydroxide, as shown below:



Volumetric Analysis Lab Report Answers Of Vinegar

R Sandford



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Laboratory Experiments for General Chemistry Harold R. Hunt,1998 *Laboratory Exercises for General, Organic, and Biological Chemistry* Sandra Migdalof Kotin,1981 Safety-Scale Lab Exp Biochem 2e Spencer L. Seager,Michael R. Slabaugh,1994-05 **Principles of General Chemistry** Jo A. Beran,2004-04-26 This remarkably popular lab manual has won over users time and time again with its exceedingly clear presentation and broad selection of topics and experiments Now revised and fine tuned this new Seventh Edition features three new experiments Water Analysis Solids Experiment 3 Vitamin C Analysis Experiment 16 and Hard Water Analysis Experiment 30 In addition nearly 90% of the Prelaboratory Assignment Questions and Laboratory Questions are either new or revised *JARQ.* ,2011 **Vinegars of the World** Laura Solieri,Paolo Giudici,2009-08-29 Vinegars can be considered as acidic products of special importance for the enrichment of our diet and resulting from the desired or controlled oxidation of ethanol containing liquid substrates The traditional use and integration of vinegars in numerous cultures can be traced back to ancient times In fact the cultural heritage of virtually every civilization includes one or more vinegars made by the souring action of micro organisms following alcoholic fermentation It has been documented that the Egyptians Sumerians and Babylonians had experience and technical knowledge in making vinegar from barley and any kind of fruit Vinegar was very popular both in ancient Greece and Rome where it was used in food preparations and as remedy against a great number of diseases In Asia the first records about vinegar date back to the Zhou Dynasty 1027-221 BC and probably China's ancient rice wines may have originally been derived from fruit for which malted rice was substituted later The historical and geographical success of vinegars is mainly due to the low technology required for their production and to the fact that several kinds of raw materials rich in sugars may easily be processed to give vinegar In addition vinegars are well known and accepted as safe and stable commodities that can be consumed as beverages health drinks or added to food as preservatives or as flavoring agents **Quantitative Chemical Analysis** Daniel C. Harris,Chuck Lucy,2015-05-29 The gold standard in analytical chemistry Dan Harris Quantitative Chemical Analysis provides a sound physical understanding of the principles of analytical chemistry and their applications in the disciplines **Chemistry 2e** Paul Flowers,Richard Langely,William R. Robinson,Klaus Hellmut Theopold,2019-02-14 Chemistry 2e is designed to meet the scope and sequence requirements of the two semester general chemistry course The textbook provides an important opportunity for students to learn the core concepts of chemistry and understand how those concepts apply to their lives and the world around them The book also includes a number of innovative features including interactive exercises and real world applications designed to enhance student learning The second edition has been revised to incorporate clearer more current and more dynamic explanations while maintaining the same organization as the first edition Substantial improvements have been made in the figures illustrations and example exercises that support the text

narrative Changes made in Chemistry 2e are described in the preface to help instructors transition to the second edition

Green Chemistry Laboratory Manual for General Chemistry Sally A. Henrie, 2015-03-18 Green chemistry involves designing novel ways to create and synthesize products and implement processes that will eliminate or greatly reduce negative environmental impacts Providing educational laboratory materials that challenge students with the customary topics found in a general chemistry laboratory manual this lab manual enables students to see how green chemistry principles can be applied to real world issues Following a consistent format each lab experiment includes objectives prelab questions and detailed step by step procedures for performing the experiments Additional questions encourage further research about how green chemistry principles compare with traditional more hazardous experimental methods **Journal of the Society of Chemical Industry** Society of Chemical Industry (Great Britain), 1890 Includes list of members 1882 1902 and proceedings of the annual meetings and various supplements **Merck's Report** Theodore Weicker, 1902 **The Merck Report**, 1903

California's Health, 1971 **District Laboratory Practice in Tropical Countries, Part 2** Monica Cheesbrough, 2006-03-02 This new edition includes an update on HIV disease AIDS recently developed HIV rapid tests to diagnose HIV infection and screen donor blood and current information on antiretroviral drugs and the laboratory monitoring of antiretroviral therapy Information on the epidemiology and laboratory investigation of other pathogens has also been brought up to date Several new rapid simple to perform immunochromatographic tests to assist in the diagnosis of infectious diseases are described including those for brucellosis cholera dengue leptospirosis syphilis and hepatitis Recently developed IgM antibody tests to investigate typhoid fever are also described The new classification of salmonellae has been introduced Details of manufacturers and suppliers now include website information and e mail addresses The haematology and blood transfusion chapters have been updated including a review of haemoglobin measurement methods in consideration of the high prevalence of anaemia in developing countries [Standardization of Potassium Permanganate Solution by Sodium Oxalate](#) Russell Smith McBride, 1913 [Comprehensive Organic Chemistry Experiments for the Laboratory Classroom](#) Carlos A. M. Afonso, Nuno R. Candeias, Dulce Pereira Simão, Alexandre F. Trindade, Jaime A. S. Coelho, Bin Tan, Robert Franzén, 2016-12-16 This expansive and practical textbook contains organic chemistry experiments for teaching in the laboratory at the undergraduate level covering a range of functional group transformations and key organic reactions The editorial team have collected contributions from around the world and standardized them for publication Each experiment will explore a modern chemistry scenario such as sustainable chemistry application in the pharmaceutical industry catalysis and material sciences to name a few All the experiments will be complemented with a set of questions to challenge the students and a section for the instructors concerning the results obtained and advice on getting the best outcome from the experiment A section covering practical aspects with tips and advice for the instructors together with the results obtained in the laboratory by students has been compiled for each experiment Targeted at professors and lecturers in chemistry this

useful text will provide up to date experiments putting the science into context for the students

Illustrated Guide to Home Chemistry Experiments Robert Bruce Thompson, 2012-02-17 For students DIY hobbyists and science buffs who can no longer get real chemistry sets this one of a kind guide explains how to set up and use a home chemistry lab with step by step instructions for conducting experiments in basic chemistry not just to make pretty colors and stinky smells but to learn how to do real lab work Purify alcohol by distillation Produce hydrogen and oxygen gas by electrolysis Smelt metallic copper from copper ore you make yourself Analyze the makeup of seawater bone and other common substances Synthesize oil of wintergreen from aspirin and rayon fiber from paper Perform forensics tests for fingerprints blood drugs and poisons and much more From the 1930s through the 1970s chemistry sets were among the most popular Christmas gifts selling in the millions But two decades ago real chemistry sets began to disappear as manufacturers and retailers became concerned about liability em The Illustrated Guide to Home Chemistry Experiments steps up to the plate with lessons on how to equip your home chemistry lab master laboratory skills and work safely in your lab The bulk of this book consists of 17 hands on chapters that include multiple laboratory sessions on the following topics Separating Mixtures Solubility and Solutions Colligative Properties of Solutions Introduction to Chemical Reactions Stoichiometry Reduction Oxidation Redox Reactions Acid Base Chemistry Chemical Kinetics Chemical Equilibrium and Le Chatelier s Principle Gas Chemistry Thermochemistry and Calorimetry Electrochemistry Photochemistry Colloids and Suspensions Qualitative Analysis Quantitative Analysis Synthesis of Useful Compounds Forensic Chemistry With plenty of full color illustrations and photos Illustrated Guide to Home Chemistry Experiments offers introductory level sessions suitable for a middle school or first year high school chemistry laboratory course and more advanced sessions suitable for students who intend to take the College Board Advanced Placement AP Chemistry exam A student who completes all of the laboratories in this book will have done the equivalent of two full years of high school chemistry lab work or a first year college general chemistry laboratory course This hands on introduction to real chemistry using real equipment real chemicals and real quantitative experiments is ideal for the many thousands of young people and adults who want to experience the magic of chemistry

Chemical Abstracts ,1928

Food Analysis Laboratory Manual S. Suzanne Nielsen, 2010-03-20 This second edition laboratory manual was written to accompany Food Analysis Fourth Edition ISBN 978 1 4419 1477 4 by the same author The 21 laboratory exercises in the manual cover 20 of the 32 chapters in the textbook Many of the laboratory exercises have multiple sections to cover several methods of analysis for a particular food component of characteristic Most of the laboratory exercises include the following introduction reading assignment objective principle of method chemicals reagents precautions and waste disposal supplies equipment procedure data and calculations questions and references This laboratory manual is ideal for the laboratory portion of undergraduate courses in food analysis

Modern Analytical Chemistry David Harvey, 2000 This introductory text covers both traditional and contemporary topics relevant to analytical chemistry Its flexible approach allows instructors

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