



Across: 3. A system, which can exchange energy, but not matter, with its surroundings. 6. Enthalpy is usually measured on the ____-joules scale. 7. Heat and work are not _____ functions. 11. The sample for which energy is studied. 15. The energy of interaction between particles, due to the position of a body (possessing energy) or to the configuration of its parts. 16. A type of change that tends to occur without needing to be driven by an external influence. 17. A process where a substance is converted from its standard state into free atoms. 18. The system and its surroundings. 21. A process, the heat of which is equal in value and opposite in sign to the heat of formation of a given substance. 22. For molar heat capacities of gases, $C_p = C_v + \underline{\hspace{1cm}}$ (two words). 23. Enthalpy change (ΔH) is the heat absorbed/released for a process at constant _____. 24. In the formula, $\Delta \underline{\hspace{1cm}} = \Delta U + \Delta n \times R \times T$, Δn represents the change in the amount of substance of gas in a chemical reaction. 25. Enthalpy of solvation is _____ positive. 26. A reaction enthalpy is the sum of the enthalpies of any sequence of reactions (at the same T and P) into which the overall reaction may be divided (two words).

Down: 1. Free energy _____ with temperature. 2. Standard entropies of the elements are not _____. 4. Molar heat capacity is the amount of heat required to raise the temperature of _____ of a substance by 1 K (or 1°C) (two words). 5. From a more condensed to a less condensed phase, ΔH is always _____. 8. The standard enthalpy of formation is the heat change when 1 mole of a substance is formed from its _____ in their most stable form. 9. The standard enthalpy of combustion is a _____ property. 10. In the formula, $\Delta H = \Delta X + \Delta n \times R \times T$, X represents _____ (two words). 12. _____ bond enthalpy $B(A-B)$ is the average energy required to break a given bond type (averaged over a series of related bonds/compounds). 13. A process where a substance is completely burnt in oxygen. 14. Everything outside the system. 19. A process is spontaneous, if free energy is _____. 20. The change in internal energy of a system equals the heat released/absorbed by the system and the work done on/by the system (two words).

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Kristen Girard Golomb



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