

VII. Graham's Law of Diffusion and Effusion - under constant temperature and pressure, gases diffuse at a rate inversely proportional to the square roots of their densities.

$$\frac{v_1}{v_2} = rr = \sqrt{\frac{FM_2}{FM_1}} \quad (\text{use densities also})$$

* diffusion - the process through which two or more different gases spontaneously mix with each other

* effusion - the process through which a gas in a closed container escapes through a small opening

* Basic Rule: lighter gases diffuse and effuse more quickly than heavier gases

I.

1) What is the relative rate of diffusion for Neon (FM = 20.2 g/mol) vs. Argon (FM = 40.0 g/mol)?
Which gas travels faster?

$$rr = \sqrt{\frac{40.0}{20.2}} = 1.47$$

2) If a mystery gas diffuses 1.48 times faster than carbon dioxide (FM = 44.0 g/mol), what is the molecular mass of the mystery gas?

$$\frac{2.19}{1} = \frac{44}{x} \quad \cancel{2.19}x = \frac{44}{2.19} = 20.19 \text{ g/mol}$$

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