

AP Chem Chapter 11 Practice Quiz

1. A solution is said to contain 28% phosphoric acid by mass. What does this mean?
 - a) 1 ml of this solution contains 28g of phosphoric acid
 - b) 1 L of this solution has a mass of 28g
 - c) 100g of this solution contains 28g of phosphoric acid
 - d) 1L of this solution contains 28ml of phosphoric acid
 - e) the density of this solution is 2.8 g/ml
2. What is the concentration of chloride ion, in percent, in a solution that contains 35.0 ppm chloride?
 - a) $3.50 \times 10^{-3} \%$
 - b) $3.50 \times 10^{-2} \%$
 - c) $3.50 \times 10^{-1} \%$
 - d) $3.50 \times 10^{-4} \%$
3. Molality is
 - a) moles solute / moles solvent
 - b) moles solute / Liters solution
 - c) moles solute / kg solution
 - d) moles solute / kg solvent
 - e) without units
4. The molality of lead nitrate in 0.726 M $\text{Pb}(\text{NO}_3)_2$ is: (The density of the solution is 1.202 g/ml)
 - a) 0.476 m
 - b) 1.928 m
 - c) 0.755 m
 - d) 0.819 m
 - e) 0.650 m
5. A solution is prepared by dissolving 15.0g of NH_3 in 250g of water. The density of the resulting solution is 0.974 g/ml. Calculate the mole fraction of NH_3 in the solution.
 - a) 0.064
 - b) 0.060
 - c) 0.94
 - d) 0.92
6. A solution is prepared by dissolving 15.0g of NH_3 in 250g of water. The density of the resulting solution is 0.974 g/ml. Calculate the molality of NH_3 in the solution.
 - a) 0.00353 M
 - b) 0.882 M
 - c) 60.0 M
 - d) 3.53 M
7. Calculate the molality of urea (molar mass = 60g/mole) in a solution prepared by dissolving 16g of urea in 39g of water.
 - a) 96 m
 - b) 6.8 m
 - c) 0.68 m
 - d) 0.41 m
8. Which one of the following varies with temperature?
 - a) molarity
 - b) mass percent
 - c) mole fraction
 - d) molality

Ap Chemistry Chapter 11 Practice Test

Neil D. Jespersen, Pamela Kerrigan



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