

H. 59. Alpha particles with their high ionizing charge but low penetrating ability can best be shielded by?

- a. paper b. clothing c. water d. lead

Alpha α particles - paper is a sufficient shield

Beta β particles - heavy clothing is sufficient to shield

Gamma rays - lead is required to shield

= polas

= polas

= non-polar

= non-polar

= non-polar

D. 60. The half-life of Se-121 is 10 days. If you started with 40 grams of this isotope, how much would you have left after 30 days?

- a. 10 g b. none c. 15 g d. 5 g

Mass of Se-121	Number of half-lives	Time
40 g	0	0
20 g	1	10 days
10 g	2	20 days
5 g	3	30 days

$\oplus \uparrow$ = particles go faster
 $\oplus \downarrow$ = particles go slower

PERIODIC TRENDS

ATOMIC RADIUS: decreased \rightarrow
 increased \downarrow

METALLIC CHARACTER:

decreases \downarrow
 increases \rightarrow

IONIZATION \oplus decreases \downarrow
 increases \rightarrow



Isotopes = atoms with same # of protons, but different # of neutrons.
EXAMPLE: Ne-20 \rightarrow ${}^{20}_{10}\text{Ne}$
 Ne-21 \rightarrow ${}^{21}_{10}\text{Ne}$
 Ne-22 \rightarrow ${}^{22}_{10}\text{Ne}$

Dilution Equation $M_1 V_1 = M_2 V_2$

- mass % = $\frac{\text{mass solute}}{\text{mass of solution (solute + solvent)}} \times 100\%$
- Molarity (M) = $\frac{\text{mol solute}}{\text{L solution}}$
- molality (m) = $\frac{\text{mol solute}}{\text{kg solvent}}$

- #protons - #electrons = Charge #
- #neutrons = A (mass #) - Z (atomic #)
- if no charge, the #protons + #electrons = same
- if + (positive) charge (cation) = #electrons is less than #protons
- if - (negative) charge (anion) = #electrons is greater than #protons
- Protons \neq Charge \rightarrow they give identity to the element.

Acs High School Chemistry Exam Study Guide

Thomas A. Holme, Kristen Murphy



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