

Part C: Electron Configuration

12. How many electrons can each level hold? 1st = \_\_\_\_\_ 2nd = \_\_\_\_\_ 3rd = \_\_\_\_\_

13. What term is used for the electrons in the outermost shell or energy level? \_\_\_\_\_

14. Scientists use two types of diagrams to show the electron configuration for atoms. Follow your teacher's directions to complete the diagrams.

Sulfur

Atomic # = 16

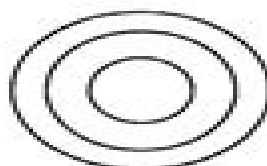
Atomic Mass = 32

Protons = \_\_\_\_\_

Neutrons = \_\_\_\_\_

Electron = \_\_\_\_\_

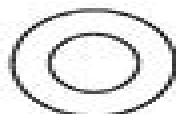
Bohr Diagram  
Shows all electrons



Lewis Structure  
Shows valence electrons

S

15. Calculate the missing information and then draw the Bohr Diagram and Lewis Structure for each element.



Li

Atomic # = 3

Mass # = 7

# of P = \_\_\_\_\_

# of N = \_\_\_\_\_

# of E = \_\_\_\_\_



Ne

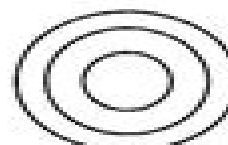
Atomic # = 10

Mass # = 20

# of P = \_\_\_\_\_

# of N = \_\_\_\_\_

# of E = \_\_\_\_\_



Mg

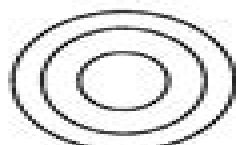
Atomic # = 12

Mass # = 24

# of P = \_\_\_\_\_

# of N = \_\_\_\_\_

# of E = \_\_\_\_\_



Cl

Atomic # = 17

Mass # = 35

# of P = \_\_\_\_\_

# of N = \_\_\_\_\_

# of E = \_\_\_\_\_



He

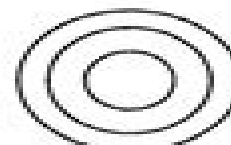
Atomic # = 2

Mass # = 4

# of P = \_\_\_\_\_

# of N = \_\_\_\_\_

# of E = \_\_\_\_\_



Si

Atomic # = 14

Mass # = 28

# of P = \_\_\_\_\_

# of N = \_\_\_\_\_

# of E = \_\_\_\_\_

16. Answer the questions below based on the elements in question #15.

(1) Which elements had a filled outermost shell? \_\_\_\_\_

(2) Which element would be most likely to lose electrons in a chemical bond? \_\_\_\_\_

(3) Which element would be most likely to gain electrons in a chemical bond? \_\_\_\_\_

(4) Which elements are not likely to bond with other elements? \_\_\_\_\_ Why? \_\_\_\_\_

# Chemical Bonds Assessment Answers

**Peipei Pang**



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