

EQUILIBRIUM CONSTANT - WORKSHEET 2_ANSWERS

1. Write the equilibrium constant expression for each of the following reactions:



$$K = \frac{[\text{NO}_2]^2}{[\text{N}_2][\text{O}_2]}$$



$$K = \frac{[\text{C}_2\text{H}_6]}{[\text{H}_2]^3}$$



$$K = \frac{1}{[\text{I}_2(\text{aq})]}$$

2. For the following reaction:



At equilibrium, $[\text{Fe}^{2+}] = 0.50 \text{ M}$, $[\text{Ag}^+] = 1.0 \text{ M}$, $[\text{Fe}^{3+}] = 1.50 \text{ M}$

a) Calculate the equilibrium constant for the reaction.

$$K = 3.0$$

b) If at equilibrium, $[\text{Fe}^{2+}] = 0.20 \text{ M}$ and $[\text{Ag}^+] = 0.30 \text{ M}$, what must be the concentration of Fe^{3+} ?

$$[\text{Fe}^{3+}] = 0.18 \text{ mol/L}$$



Calculate the equilibrium concentration of oxygen gas if the NO and NO_2 concentrations are equal at equilibrium.

$$[\text{O}_2] = 2.0 \times 10^{-6} \text{ mol/L}$$



Which of the above reactions would have the greatest amount of products present at equilibrium compared to reactants?

Reaction b. Since K is larger the amount of products must be larger, because K is products/reactants.

5. For the reaction:



at an equilibrium temperature of 1000°C , a 1.00 L flask contains 0.120 moles of ammonia, 1.03 moles of nitrogen and 1.62 moles of hydrogen. Calculate the equilibrium constant at this temperature.

$$K = 3.29 \times 10^{-2}$$

Equilibrium Constant Answers

**Ralph H. Petrucci, F. Geoffrey
Herring, Jeffry D. Madura, Carey
Bissonnette**

Equilibrium Constant Answers

Chemistry 2e Paul Flowers, Richard Langely, William R. Robinson, Klaus Hellmut Theopold, 2019-02-14 Chemistry 2e is designed to meet the scope and sequence requirements of the two semester general chemistry course. The textbook provides an important opportunity for students to learn the core concepts of chemistry and understand how those concepts apply to their lives and the world around them. The book also includes a number of innovative features including interactive exercises and real world applications designed to enhance student learning. The second edition has been revised to incorporate clearer, more current and more dynamic explanations while maintaining the same organization as the first edition. Substantial improvements have been made in the figures, illustrations, and example exercises that support the text narrative. Changes made in Chemistry 2e are described in the preface to help instructors transition to the second edition. Chemistry Steven S. Zumdahl, Susan A. Zumdahl, 2012 Steve and Susan Zumdahl's texts focus on helping students build critical thinking skills through the process of becoming independent problem solvers. They help students learn to think like a chemist so they can apply the problem solving process to all aspects of their lives. In CHEMISTRY: AN ATOMS FIRST APPROACH 1e International Edition, the Zumdahls use a meaningful approach that begins with the atom and proceeds through the concept of molecules, structure, and bonding to more complex materials and their properties. Because this approach differs from what most students have experienced in high school courses, it encourages them to focus on conceptual learning early in the course rather than relying on memorization and a plug-and-chug method of problem solving that even the best students can fall back on when confronted with familiar material. The atoms first organization provides an opportunity for students to use the tools of critical thinkers to ask questions, to apply rules and models, and to **General Chemistry** Ralph H. Petrucci, F. Geoffrey Herring, Jeffrey D. Madura, Carey Bissonnette, 2010-05 *Principles of Modern Chemistry* David W. Oxtoby, 1998-07-01 PRINCIPLES OF MODERN CHEMISTRY has dominated the honors and high mainstream general chemistry courses and is considered the standard for the course. The fifth edition is a substantial revision that maintains the rigor of previous editions but reflects the exciting modern developments taking place in chemistry today. Authors David W. Oxtoby and H. P. Gillis provide a unique approach to learning chemical principles that emphasizes the total scientific process from observation to application, placing general chemistry into a complete perspective for serious-minded science and engineering students. Chemical principles are illustrated by the use of modern materials comparable to equipment found in the scientific industry. Students are therefore exposed to chemistry and its applications beyond the classroom. This text is perfect for those instructors who are looking for a more advanced general chemistry textbook. The Principles of Chemical Equilibrium Kenneth George Denbigh, 1981-03-26 Sample Text *Chemistry* Bruce Averill, Patricia Eldredge, 2007 Emphasizes on contemporary applications and an intuitive problem solving approach that helps students discover the exciting potential of

chemical science This book incorporates fresh applications from the three major areas of modern research materials environmental chemistry and biological science

Chemistry 2e Paul Flowers, Klaus Theopold, Richard Langley, Edward J. Neth, William R. Robinson, 2019-02-14 Chemistry 2e is designed to meet the scope and sequence requirements of the two semester general chemistry course The textbook provides an important opportunity for students to learn the core concepts of chemistry and understand how those concepts apply to their lives and the world around them The book also includes a number of innovative features including interactive exercises and real world applications designed to enhance student learning The second edition has been revised to incorporate clearer more current and more dynamic explanations while maintaining the same organization as the first edition Substantial improvements have been made in the figures illustrations and example exercises that support the text narrative Changes made in Chemistry 2e are described in the preface to help instructors transition to the second edition

The Iron(III) Thiocyanate Reaction Kevin C. de Berg, 2019-11-13 This Brief presents an historical investigation into the reaction between ferric ions and thiocyanate ions which has been viewed in different ways throughout the last two centuries Historically the reaction was used in chemical analysis and to highlight the nature of chemical reactions the laws of chemistry models and theories of chemistry chemical nomenclature mathematics and data analysis and instrumentation which are important ingredients of what one might call the nature of chemistry Using the history of the iron III thiocyanate reaction as a basis the book s main objective is to explore how chemistry develops its own knowledge base how it assesses the reliability of that base and how some important tools of the trade have been brought to bear on a chemical reaction to achieve understanding a worthwhile goal of any historical investigation

The Thermodynamics of Phase and Reaction Equilibria Ismail Tosun, 2012-10-17 This book provides you with a sound foundation for understanding abstract concepts eg physical properties such as fugacity etc or chemical processes ie distillation etc of phase and reaction equilibria and shows you how to apply these concepts to solve practical problems using numerous and clear examples

General Chemistry Ralph H. Petrucci, Ralph Petrucci, F. Geoffrey Herring, Jeffry Madura, Carey Bissonnette, 2017 The most trusted general chemistry text in Canada is back in a thoroughly revised 11th edition General Chemistry Principles and Modern Applications is the most trusted book on the market recognized for its superior problems lucid writing and precision of argument and precise and detailed and treatment of the subject The 11th edition offers enhanced hallmark features new innovations and revised discussions that that respond to key market needs for detailed and modern treatment of organic chemistry embracing the power of visual learning and conquering the challenges of effective problem solving and assessment Note You are purchasing a standalone product MasteringChemistry does not come packaged with this content Students if interested in purchasing this title with MasteringChemistry ask your instructor for the correct package ISBN and Course ID Instructors contact your Pearson representative for more information If you would like to purchase both the physical text and MasteringChemistry search for 0134097327 9780134097329 General Chemistry

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Chemistry, Life, the Universe and Everything Melanie Cooper, Michael Klymkowsky, 2014-06-27 As you can see this molecular formula is not very informative it tells us little or nothing about their structure and suggests that all proteins are similar which is confusing since they carry out so many different roles *Kaplan MCAT General Chemistry Review* Kaplan, 2015-07-07 More people get into medical school with a Kaplan MCAT course than all major courses combined Now the same results are available with Kaplan s *MCAT General Chemistry Review* This book features thorough subject review more questions than any competitor and the highest yield questions available The commentary and instruction come directly from Kaplan MCAT experts and include targeted focus on the most tested concepts plus more questions than any other guide *Kaplan s MCAT General Chemistry Review* offers UNPARALLELED MCAT KNOWLEDGE The Kaplan MCAT team has spent years studying every document related to the MCAT available In conjunction with our expert psychometricians the Kaplan team is able to ensure the accuracy and realism of our practice materials THOROUGH SUBJECT REVIEW Written by top rated award winning Kaplan instructors All material has been vetted by editors with advanced science degrees and by a medical doctor EXPANDED CONTENT THROUGHOUT While the MCAT has continued to develop this book has been updated

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Introduction to Atmospheric Chemistry Daniel J. Jacob, 1999 Atmospheric chemistry is one of the fastest growing fields in the earth sciences Until now however there has been no book designed to help students capture the essence of the subject in a brief course of study Daniel Jacob a leading researcher and teacher in the field addresses that problem by presenting the first textbook on atmospheric chemistry for a one semester course Based on the approach he developed in his class at Harvard Jacob introduces students in clear and concise chapters to the fundamentals as well as the latest ideas and findings in the field Jacob's aim is to show students how to use basic principles of physics and chemistry to describe a complex system such as the atmosphere He also seeks to give students an overview of the current state of research and the work that led to this point Jacob begins with atmospheric structure design of simple models atmospheric transport and the continuity equation and continues with geochemical cycles the greenhouse effect aerosols stratospheric ozone the oxidizing power of the atmosphere smog and acid rain Each chapter concludes with a problem set based on recent scientific literature This is a novel approach to problem set writing and one that successfully introduces students to the prevailing issues This is a major contribution to a growing area of study and will be welcomed enthusiastically by students and teachers alike

Modern Engineering Thermodynamics - Textbook with Tables Booklet Robert T. Balmer, 2011-01-03 Modern Engineering Thermodynamics Textbook with Tables Booklet offers a problem solving approach to basic and applied engineering thermodynamics with historical vignettes critical thinking boxes and case studies throughout to help relate abstract concepts to actual engineering applications It also contains applications to modern engineering issues This textbook is designed for use in a standard two semester engineering thermodynamics course sequence with the goal of helping students develop engineering problem solving skills through the use of structured problem solving techniques The first half of the text contains material suitable for a basic Thermodynamics course taken by engineers from all majors The second half of the text is suitable for an Applied Thermodynamics course in mechanical engineering programs The Second Law of Thermodynamics is introduced through a basic entropy concept providing students a more intuitive understanding of this key course topic Property Values are discussed before the First Law of Thermodynamics to ensure students have a firm understanding of

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The Law of Mass Action Andrei B. Koudriavtsev,Reginald F. Jameson,Wolfgang Linert,2011-06-27 Why are atoms so small asks naive physicist in Erwin Schrodinger s book What is Life The Physical Aspect of the Living Cell The question is wrong answers the author the actual problem is why we are built of such an enormous number of these particles The idea that everything is built of atoms is quite an old one It seems that Democritus himself borrowed it from some obscure Phoenician source The arguments for the existence of small indivisible units of matter were quite simple 2 According to Lucretius observable matter would disappear by wear and tear the world exists for a sufficiently long if not infinitely long time unless there are some units which cannot be further split into parts th However in the middle of the 19 century any reference to the atomic structure of matter was considered among European physicists as a sign of extremely bad taste and provinciality The hypothesis of the ancient Greeks for Lucretius had translated Epicurean philosophy into Latin hexameters was at that time seen as bringing nothing positive to exact science The properties of gaseous liquid and solid bodies as well as the behaviour of heat and energy were successfully described by the rapidly developing science of thermodynamics

Concept Development Studies in Chemistry John S. Hutchinson,2009-09-24 This is an on line textbook for an Introductory General Chemistry course Each module develops a central concept in Chemistry from experimental observations and inductive reasoning This approach complements an interactive or active learning teaching approach Additional multimedia resources can be found at <http://cnx.org/content/col10264/1.5>

Pearson Chemistry 12 New South Wales Skills and Assessment Book Penny Commons,2018-10-15 The write in Skills and Assessment Activity Books focus on working scientifically skills and assessment They are designed to consolidate concepts learnt in class Students are also provided with regular opportunities for reflection and self evaluation throughout the book

Chemistry Edward J. Neth,Pau Flowers,Klaus Theopold,William R. Robinson,Richard Langley,2016-06-07 Chemistry Atoms First is a

peer reviewed openly licensed introductory textbook produced through a collaborative publishing partnership between OpenStax and the University of Connecticut and UConn Undergraduate Student Government Association This title is an adaptation of the OpenStax Chemistry text and covers scope and sequence requirements of the two semester general chemistry course Reordered to fit an atoms first approach this title introduces atomic and molecular structure much earlier than the traditional approach delaying the introduction of more abstract material so students have time to acclimate to the study of chemistry Chemistry Atoms First also provides a basis for understanding the application of quantitative principles to the chemistry that underlies the entire course Open Textbook Library **World of Chemistry** Steven S. Zumdahl, Susan L. Zumdahl, Donald J. DeCoste, 2006-08 Our high school chemistry program has been redesigned and updated to give your students the right balance of concepts and applications in a program that provides more active learning more real world connections and more engaging content A revised and enhanced text designed especially for high school helps students actively develop and apply their understanding of chemical concepts Hands on labs and activities emphasize cutting edge applications and help students connect concepts to the real world A new captivating design clear writing style and innovative technology resources support your students in getting the most out of their textbook Publisher

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