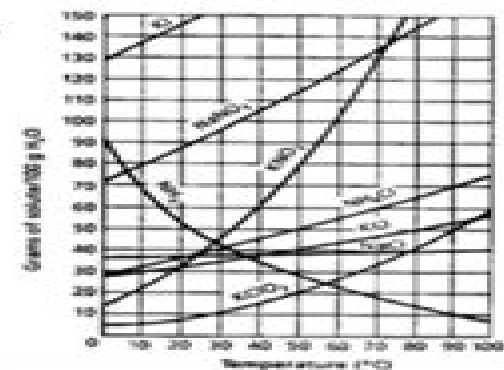


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SOLUTIONS WORKSHEET

A. Solubility curves – refer to the graph to answer the following questions.



1. What is the solubility of NH₄Cl at 70°C? **60 g**
2. Calculate the mass of NH₄Cl that would dissolve in 200 g of water at 90°C.

70 g @ 100g water → 140 g

3. At 0°C, 50.0 g NH₄Cl are added to 100 g water. To what temperature does the solution have to be raised in order to achieve a saturated solution of NH₄Cl_(aq)?

50°C

4. Is NH₃ a (s), (l), or (g)? Explain.

GAS – it's solubility decreases with increase in temperature. Opposite for solids

5. Determine how much water is required to dissolve 35.0 g NH₄Cl at 70°C in order to make a saturated solution. Show your work.

$$\frac{60}{100} = \frac{35}{x} \quad x = 58.3 \text{ g H}_2\text{O}$$

6. If a saturated solution of NH₃ at 15°C is warmed to 40°C, how much ammonia will "fall out" of the solution?

@ 15°C – 60 g soluble @ 40°C – 35 g soluble 25 g will fall out

B. MOLARITY

1. What is molarity a measure of? What is the formula for molarity? What are the units of molarity?

Concentration. $M = n / V$ mol/L or M

2. Calculate the molarity of a 1.5 L solution that has 0.45 mol solute dissolved.

0.30 - M

3. Calculate the grams of sodium sulfate that are required to prepare 1.50 L of 0.25 M solution.

53.3 g Na₂SO₄

4. Calculate the volume of solution required to make 0.024 M solution using 15.0 g of calcium chloride.

V = 5.6 L

5. Describe in detail the steps you would take to prepare 1.50 L of 0.5 M solution of sodium chloride.

n = 0.75 mol → 43.5 g NaCl

- 1) measure out 43.5 g NaCl
- 2) put NaCl into graduated cylinder
- 3) add enough water to make final volume 1.50 L

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Table of Contents Chemistry Solutions Worksheet

1. Understanding the eBook Chemistry Solutions Worksheet
 - The Rise of Digital Reading Chemistry Solutions Worksheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Solutions Worksheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals

3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Solutions Worksheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Solutions Worksheet
 - Personalized Recommendations
 - Chemistry Solutions Worksheet User Reviews and Ratings
 - Chemistry Solutions Worksheet and Bestseller Lists
5. Accessing Chemistry Solutions Worksheet Free and Paid eBooks
 - Chemistry Solutions Worksheet Public Domain eBooks
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 - Chemistry Solutions Worksheet Budget-Friendly Options
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 - Chemistry Solutions Worksheet Compatibility with Devices
 - Chemistry Solutions Worksheet Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Solutions Worksheet
 - Highlighting and Note-Taking Chemistry Solutions Worksheet
 - Interactive Elements Chemistry Solutions Worksheet
8. Staying Engaged with Chemistry Solutions Worksheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Solutions Worksheet
9. Balancing eBooks and Physical Books Chemistry Solutions Worksheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Solutions Worksheet
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions

- Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Solutions Worksheet
 - Setting Reading Goals Chemistry Solutions Worksheet
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Solutions Worksheet
 - Fact-Checking eBook Content of Chemistry Solutions Worksheet
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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