

## Experiment 7 Lab Summary

### Experiment 7 – Hydrolysis of Salts and pH of Buffer Solutions

**Purpose** – To learn about the concept of hydrolysis and to gain familiarity with acid-base indicators and the behavior of buffer solutions.

#### Procedure –

##### A. Hydrolysis of Salts

1. Boil **50ml of distilled water** for 10 minutes. (this expels  $\text{CO}_2$ )
2. Cover + allow water to cool to room temp.
3. Add **5ml of distilled water** (not boiling this time) to six test tubes. Add **three drops of a different indicator** to each of the test tubes (one indicator per tube) and record the colors on the report sheet.
4. From the identified colors and data, use the chart in your manual to identify the pH of the nonboiled water. (Note: pH should be  $>7$  because of the undissolved  $\text{CO}_2$ )
5. Empty the contents of the six test tubes and rinse them out thoroughly.
6. Pour **5ml of boiled distilled water** into each of the test tubes.
7. Add **three drops of each indicator** into each of the six test tubes. (one indicator per tube)
8. Record the colors and determine the pH.
9. Empty the contents of the six test tubes + rinse them each three times with distilled (boiling?) water.
10. Repeat the same procedure to 5ml each of the following solutions:
  - 0.1M  $\text{NaCl}$
  - 0.1M  $\text{NaC}_2\text{H}_3\text{O}_2$
  - 0.1M  $\text{NH}_4\text{Cl}$
  - 0.1M  $\text{ZnCl}_2$
  - 0.1M  $\text{KAl(SO}_4)_2$
  - 0.1M  $\text{Na}_2\text{CO}_3$(Be sure to rinse three times between each solution change.)
11. From the pH values you determined above, calculate the hydroxide and hydrogen ion concentrations for each solution.
12. Complete the tables on the report sheet, calculate the  $K_a$  or  $K_b$ , and dispose of the chemicals as instructed.

##### B. pH of Buffer Solutions

###### Part I. Preparation of Acetic Acid-Sodium Acetate Buffer

13. Determine + Record the exact mass of **3.5g  $\text{NaC}_2\text{H}_3\text{O}_2 \cdot 3\text{H}_2\text{O}$**  (to the nearest .01g)
14. Add the 3.5g  $\text{NaC}_2\text{H}_3\text{O}_2 \cdot 3\text{H}_2\text{O}$  to a 150ml beaker.
15. Using a 20ml graduated cylinder, measure **8.5ml of 3.0M acetic acid** + add it to the beaker containing the sodium acetate.

# Chem 1108 Lab Manual Answers

**David Harvey**



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