

## PART I : CHEMISTRY

### SECTION – I (Total Marks : 24)

(Single Correct Answer Type)

This section contains 8 multiple choice questions. Each question has four choices (A), (B), (C) and (D) out of which **ONLY ONE** is correct.

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1. Oxidation states of the metal in the minerals haematite and magnetite, respectively, are
- (A) II, III in haematite and III in magnetite  
(B) II, III in haematite and II in magnetite  
(C) II in haematite and II, III in magnetite  
(D) III in haematite and II, III in magnetite

**ANSWER : D**

2. Among the following complexes (K–P),  
 $K_3[Fe(CN)_6]$  (K),  $[Co(NH_3)_6]Cl_3$  (L),  $Na_2[Co(oxalate)_3]$  (M),  $[Ni(H_2O)_6]Cl_2$  (N),  
 $K_2[Pt(CN)_4]$  (O) and  $[Zn(H_2O)_6](NO_3)_2$  (P)  
the diamagnetic complexes are
- (A) K, L, M, N      (B) K, M, O, P      (C) L, M, O, P      (D) L, M, N, O

**ANSWER : C**

3. Passing  $H_2S$  gas into a mixture of  $Mn^{2+}$ ,  $Ni^{2+}$ ,  $Cu^{2+}$  and  $Hg^{2+}$  ions in an acidified aqueous solution precipitates
- (A) CuS and HgS      (B) MnS and CuS  
(C) MnS and NiS      (D) NiS and HgS

**ANSWER : A**

4. Consider the following cell reaction:



At  $[Fe^{2+}] = 10^{-3} \text{ M}$ ,  $P(O_2) = 0.1 \text{ atm}$  and  $pH = 3$ , the cell potential at  $25^\circ\text{C}$  is

- (A) 1.47 V      (B) 1.77 V      (C) 1.87 V      (D) 1.57 V

**ANSWER : D**

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