

Molarity:

- a quantitative description of solution concentration.
- Abbreviated M

$$\text{Molarity} = \frac{\text{moles of solute}}{\text{liters of solution}}$$

Problems: Show all work and circle your final answer.

1. To make a 4.00 M solution, how many moles of solute will be needed if 12.0 liters of solution are required?

$$4.00 \text{ M} = \frac{\text{moles of solute}}{12.0 \text{ L}} \quad \text{moles of solute} = \boxed{48.0 \text{ mol}}$$

2. How many moles of sucrose are dissolved in 250 mL of solution if the solution concentration is 0.150 M?

$$? \text{ L} = 250 \text{ mL} \times \frac{1 \text{ L}}{1000 \text{ mL}} = 0.25 \text{ L}$$

$$0.150 \text{ M} = \frac{\text{moles of solute}}{0.25 \text{ L}} \quad \text{moles of solute} = \boxed{0.038 \text{ mol}}$$

3. What is the molarity of a solution of HNO₃ that contains 12.6 grams HNO₃ in 1.0 L of solution?

$$? \text{ mol HNO}_3 = 12.6 \text{ g HNO}_3 \times \frac{1 \text{ mol HNO}_3}{63.0 \text{ g HNO}_3} = 0.200 \text{ mol HNO}_3$$

$$M = \frac{0.200 \text{ mol HNO}_3}{1.0 \text{ L}} = \boxed{0.200 \text{ M}}$$

4. How many grams of potassium nitrate are required to prepare 0.250 L of a 0.700 M solution?

$$0.700 \text{ M} = \frac{\text{moles of solute}}{0.250 \text{ L}} \quad \text{moles of solute} = 0.175 \text{ mol}$$

$$? \text{ g KNO}_3 = 0.175 \text{ mol KNO}_3 \times \frac{101.1 \text{ g KNO}_3}{1 \text{ mol KNO}_3} = \boxed{17.7 \text{ g KNO}_3}$$

Chemistry Molarity Of Solutions Worksheet Answer Key

**Robert Nelson Smith, Willis Conway
Pierce**



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The Laboratory Barbara Dungey, 2006-01-01

General Chemistry Ralph H. Petrucci, F. Geoffrey Herring, Jeffry D. Madura, Carey Bissonnette, 2010-05

An Introduction to Chemistry Mark Bishop, 2002 This book teaches chemistry at an appropriate level of rigor while removing the confusion and insecurity that impair student success Students are frequently intimidated by

prep chem Bishop s text shows them how to break the material down and master it The flexible order of topics allows unit conversions to be covered either early in the course as is traditionally done or later allowing for a much earlier than usual description of elements compounds and chemical reactions The text and superb illustrations provide a solid conceptual framework and address misconceptions The book helps students to develop strategies for working problems in a series of logical steps The Examples and Exercises give plenty of confidence building practice the end of chapter problems test the student s mastery The system of objectives tells the students exactly what they must learn in each chapter and where to find it

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