



55. Phosphorus has three allotropic forms — (i) white phosphorus (ii) red phosphorus and (iii) black phosphorus. Write the difference between white and red phosphorus on the basis of their structure and reactivity.

**Solution:**

1. White phosphorus

(i) It is soft waxy solid having garlic odour

(ii) It is poisonous

(iii) Has low melting point and boiling point because P<sub>4</sub> molecules are held together by weak van der Waal forces of attraction

2. Red Phosphorus

(i) It is hard, crystalline, odourless solid

(ii) It is non-poisonous

(iii) It has a high melting point because of its polymeric structure.

56. Give an example to show the effect of concentration of nitric acid on the formation of oxidation product.

**Solution:**



57.  $\text{PCl}_5$  reacts with finely divided silver on heating and a white silver salt is obtained, which dissolves on adding excess aqueous  $\text{NH}_3$  solution. Write the reactions involved to explain what happens.

**Solution:**

$\text{PCl}_5$  reacts with silver to form white silver salt ( $\text{AgCl}$ ). Which then dissolves on adding excess aqueous  $\text{NH}_3$  to form a soluble complex



58. Phosphorus forms a number of oxoacids. Out of these oxoacids phosphinic acid has strong

# Chemistry Chapter 12 Review Answers

**T. W. Graham Solomons, Craig B.  
Fryhle, Scott A. Snyder**



# Chemistry Chapter 12 Review Answers

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