

LOHS Honors Chemistry

Balancing Chemical Equations

Part I: Balance these chemical equations. Record what type of reaction on the blank in the right margin.

- | | |
|--|-----------|
| 1. $\text{BaCl}_2 + 2\text{KClO}_3 \rightarrow \text{Ba}(\text{ClO}_3)_2 + 2\text{KCl}$ | <u>DR</u> |
| 2. $\text{Pb}(\text{OH})_2 \rightarrow \text{PbO} + \text{H}_2\text{O}$ | <u>D</u> |
| 3. $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$ | <u>SR</u> |
| 4. $\text{Pb}(\text{NO}_3)_2 + \text{K}_2\text{CrO}_4 \rightarrow \text{PbCrO}_4 + 2\text{KNO}_3$ | <u>DR</u> |
| 5. $\text{Al}_2(\text{SO}_4)_3 + 3\text{Ca}(\text{OH})_2 \rightarrow 2\text{Al}(\text{OH})_3 + 3\text{CaSO}_4$ | <u>DR</u> |
| 6. $\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$ | <u>SR</u> |
| 7. $\text{Sn} + 2\text{KOH} \rightarrow \text{K}_2\text{SnO}_2 + \text{H}_2$ | <u>SR</u> |
| 8. $2\text{PaI}_5 \rightarrow 2\text{Pa} + 5\text{I}_2$ | <u>D</u> |
| 9. $3\text{SiF}_4 + 3\text{H}_2\text{O} \rightarrow 2\text{H}_2\text{SiF}_6 + \text{H}_2\text{SiO}_3$ | <u>SR</u> |
| 10. $\text{Cl}_2\text{O}_7 + \text{H}_2\text{O} \rightarrow 2\text{HClO}_4$ | <u>S</u> |
| 11. $16\text{Na} + \text{S}_8 \rightarrow 8\text{Na}_2\text{S}$ | <u>S</u> |
| 12. $\text{N}_2\text{O}_3 + \text{H}_2\text{O} \rightarrow 2\text{HNO}_2$ | <u>S</u> |

Part II: Write balanced chemical equations for each of the reactions below. Assume that all ionic compounds are dissolved in water (aqueous). Also be sure include physical states where appropriate.

1. Iron metal combines with oxygen gas to form iron (III) oxide solid



2. Chromium metal and sulfuric acid (H_2SO_4) react to form chromium (III) sulfate solution and hydrogen gas.



3. Barium chromate solution and aqueous hydrochloric acid (HCl) react to produce barium chloride solution and hydrogen chromate (solid).



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