

MIXED GAS LAWS WORKSHEET - SOLUTIONS

1) How many moles of gas occupy 98 L at a pressure of 2.8 atmospheres and a temperature of 292 K?

$$n = \frac{PV}{RT} = \frac{(2.8 \text{ atm})(98 \text{ L})}{(0.0821 \text{ Latm/mol.K})(292 \text{ K})} = 11 \text{ moles of gas}$$

2) If 5.0 moles of O_2 and 3.0 moles of N_2 are placed in a 30.0 L tank at a temperature of $25^\circ C$, what will the pressure of the resulting mixture of gases be? $25^\circ C = 298 \text{ K}$

$$O_2: P = \frac{nRT}{V} = \frac{(5.0 \text{ mol})(0.0821 \text{ Latm/mol.K})(298 \text{ K})}{(30.0 \text{ L})} = 4.1 \text{ atm}$$

$$N_2: P = \frac{nRT}{V} = \frac{(3.0 \text{ mol})(0.0821 \text{ Latm/mol.K})(298 \text{ K})}{(30.0 \text{ L})} = 2.4 \text{ atm}$$

$$P_{\text{tot}} = P_{O_2} + P_{N_2} = 4.1 \text{ atm} + 2.4 \text{ atm} = 6.5 \text{ atm}$$

Alternately you can substitute total moles of gas (8.0) as the total pressure depends on the total moles of gas – what type of gas is irrelevant

3) A balloon is filled with 35.0 L of helium in the morning when the temperature is $20.0^\circ C$. By noon the temperature has risen to $45.0^\circ C$. What is the new volume of the balloon?

$$T_1 = 20.0^\circ C = 293 \text{ K}, V_1 = 35.0 \text{ L}, T_2 = 45.0^\circ C = 318 \text{ K}, V_2 = ?$$

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(35.0 \text{ L})(318 \text{ K})}{(293 \text{ K})} = 38.0 \text{ L}$$

4) A 35 L tank of oxygen is at 315 K with an internal pressure of 190 atmospheres. How many moles of gas does the tank contain?

$$n = \frac{PV}{RT} = \frac{(190 \text{ atm})(35 \text{ L})}{(0.0821 \text{ Latm/mol.K})(315 \text{ K})} = 260 \text{ moles of gas}$$

5) A balloon that can hold 85 L of air is inflated with 3.5 moles of gas at a pressure of 1.0 atmosphere. What is the temperature in $^\circ C$ of the balloon?

$$T = \frac{PV}{nR} = \frac{(1 \text{ atm})(85 \text{ L})}{(3.5 \text{ mol})(0.0821 \text{ Latm/mol.K})} = 296 \text{ K} = 23^\circ C$$

6) $CaCO_3$ decomposes at $1200^\circ C$ to form CO_2 gas and CaO . If 25 L of CO_2 are collected at $1200^\circ C$, what will the volume of this gas be after it cools to $25^\circ C$?

$$T_1 = 1200^\circ C = 1473 \text{ K}, V_1 = 25 \text{ L}, T_2 = 25^\circ C = 298 \text{ K}, V_2 = ?$$

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(25 \text{ L})(298 \text{ K})}{(1473 \text{ K})} = 5.1 \text{ L}$$

7) A helium balloon with an internal pressure of 1.00 atm and a volume of 4.50 L at $20.0^\circ C$ is released. What volume will the balloon occupy at an altitude where the pressure is 0.600 atm and the temperature is $-20.0^\circ C$?

$$P_1 = 1.00 \text{ atm}, V_1 = 4.50 \text{ L}, T_1 = 20.0^\circ C = 293 \text{ K}, P_2 = 0.600 \text{ atm}, V_2 = ?, T_2 = -20.0^\circ C = 253 \text{ K}$$

$$V_2 = \frac{P_1 V_1 T_2}{T_1 P_2} = \frac{(1.00 \text{ atm})(4.50 \text{ L})(253 \text{ K})}{(293 \text{ K})(0.600 \text{ atm})} = 6.48 \text{ L}$$

Mixed Gas Laws Worksheet Answers

U. S. Customs and Border Protection



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GAS LAWS NARAYAN CHANGDER,2024-04-01 THE GAS LAWS MCQ MULTIPLE CHOICE QUESTIONS SERVES AS A VALUABLE RESOURCE FOR INDIVIDUALS AIMING TO DEEPEN THEIR UNDERSTANDING OF VARIOUS COMPETITIVE EXAMS CLASS TESTS QUIZ COMPETITIONS AND SIMILAR ASSESSMENTS WITH ITS EXTENSIVE COLLECTION OF MCQS THIS BOOK EMPOWERS YOU TO ASSESS YOUR GRASP OF THE SUBJECT MATTER AND YOUR PROFICIENCY LEVEL BY ENGAGING WITH THESE MULTIPLE CHOICE QUESTIONS YOU CAN IMPROVE YOUR KNOWLEDGE OF THE SUBJECT IDENTIFY AREAS FOR IMPROVEMENT AND LAY A SOLID FOUNDATION DIVE INTO THE GAS LAWS MCQ TO EXPAND YOUR GAS LAWS KNOWLEDGE AND EXCEL IN QUIZ COMPETITIONS ACADEMIC STUDIES OR PROFESSIONAL ENDEAVORS THE ANSWERS TO THE QUESTIONS ARE PROVIDED AT THE END OF EACH PAGE MAKING IT EASY FOR PARTICIPANTS TO VERIFY THEIR ANSWERS AND PREPARE EFFECTIVELY

Mixed Gas Research Investigation
American Gas Association. Laboratories, *Williams & Meyers Oil and Gas Law* Howard R. Williams,2010 **Oil and Gas Law in a Nutshell** John S. Lowe,2003 *Chemistry 2e* Paul Flowers,Richard Langely,William R. Robinson,Klaus Hellmut Theopold,2019-02-14 Chemistry 2e is designed to meet the scope and sequence requirements of the two semester general chemistry course The textbook provides an important opportunity for students to learn the core concepts of chemistry and understand how those concepts apply to their lives and the world around them The book also includes a number of innovative features including interactive exercises and real world applications designed to enhance student learning The second edition has been revised to incorporate clearer more current and more dynamic explanations while maintaining the same organization as the first edition Substantial improvements have been made in the figures illustrations and example exercises that support the text narrative Changes made in Chemistry 2e are described in the preface to help instructors transition to the second edition *University Physics* Samuel J. Ling,Jeff Sanny,William Moebs,2017-12-19 University Physics is designed for the two or three semester calculus based physics course The text has been developed to meet the scope and sequence of most university physics courses and provides a foundation for a career in mathematics science or engineering The book provides an important opportunity for students to learn the core concepts of physics and understand how those concepts apply to their lives and to the world around them Due to the comprehensive nature of the material we are offering the book in three volumes for flexibility and efficiency Coverage and Scope Our University Physics textbook adheres to the scope and sequence of most two and three semester physics courses nationwide We have worked to make physics interesting and accessible to students while maintaining the mathematical rigor inherent in the subject With this objective in mind the content of this textbook has been developed and arranged to provide a logical progression from fundamental to more advanced concepts building upon what students have already learned and emphasizing connections between topics and

between theory and applications The goal of each section is to enable students not just to recognize concepts but to work with them in ways that will be useful in later courses and future careers The organization and pedagogical features were developed and vetted with feedback from science educators dedicated to the project

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General Chemistry Ralph H. Petrucci, F. Geoffrey Herring, Jeffrey D. Madura, Carey Bissonnette, 2010-05

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Chemistry 2e Paul Flowers, Klaus Theopold, Richard Langley, Edward J. Neth, William R. Robinson, 2019-02-14 Chemistry 2e is designed to meet the scope and sequence requirements of the two semester general chemistry course The textbook provides an important opportunity for students to learn the core concepts of chemistry and understand how those concepts apply to their lives and the world around them The book also includes a number of innovative features including interactive exercises and real world applications designed to enhance student learning The second edition has been revised to incorporate clearer more current and more dynamic explanations while maintaining the same organization as the first edition Substantial improvements have been made in the figures illustrations and example exercises that support the text narrative Changes made in Chemistry 2e are described in the preface to help instructors transition to the second edition

Quantities, Units and Symbols in Physical Chemistry International Union of Pure and Applied Chemistry. Physical and Biophysical Chemistry Division, 2007 Prepared by the IUPAC Physical Chemistry Division this definitive manual now in its third edition is designed to improve the exchange of scientific information among the readers in different disciplines and across different nations This book has been systematically brought up to date and new sections added to reflect the increasing volume of scientific literature and terminology and expressions being used The Third Edition reflects the experience of the contributors with the previous editions and the comments and feedback have been integrated into this essential resource This edition has been compiled in

machine readable form and will be available online *Chemical Engineering Design* Gavin Towler, Ray Sinnott, 2012-01-25

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