

w % (Na_2SO_4)	w % (NaCl)	w % (Na_2CO_3)	$\Delta_{\text{dis}}S$ ($\text{J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$)	$\Delta_{\text{dis}}H$ ($\text{kJ}\cdot\text{mol}^{-1}$)	R^2
-	-	-	67.99	34.75	0.9963
-	1.01	-	41.71	34.75	0.9977
0.96	-	-	51.69	34.42	0.9982
-	-	1.03	62.68	41.27	0.9918
0.99	-	0.91	72.51	44.44	0.9963
-	1.03	1.00	74.22	45.97	0.9934
-	2.07	1.01	75.22	58.07	0.9963
-	3.13	1.02	113.17	57.95	0.9965
-	4.16	1.00	108.73	57.07	0.9833
-	7.27	1.02	119.18	60.53	0.9795
1.04	1.04	0.97	84.56	48.49	0.9818
2.11	1.03	0.96	115.76	57.88	0.9859
2.99	1.04	1.01	113.51	57.44	0.9884
3.94	1.03	0.97	109.75	56.44	0.9989
8.37	1.05	0.99	115.33	58.41	0.9963

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Thermodynamics of 1:2 Sodium Borate Minerals Laura Suzanne Ruhl,2008 The increase in the equilibrium constant for the dehydration reaction from borax to tinalconite is consistent with an enthalpy of dehydration to water vapor at 323 15K of 55 3 2 0 kJ mol of H₂O This is also consistent with the enthalpy of dehydration determined by HF solution calorimetry of 54 33 0 24 kJ mol of H₂O and by DSC of 55 6kJ 1 1 kJ mol of H₂O The enthalpy of dehydration from borax to kernite and tinalconite to kernite determined by HF calorimetry was 56 8 kJ mol and 1 96 kJ mol respectively These results along with heat capacities for borax tinalconite and kernite determined by DSC and molar volumes determined by XRPD were used to assess the thermodynamic properties of borax dehydration as a function of temperature and pressure The resulting phase diagram is consistent with geologic observations of Na borate stability indicating that in the Na₂B₄O₇ H₂O system borax is the primary precipitate and stable at earth s surface conditions Tinalconite is metastable under most of these conditions and kernite is the result of secondary mineralization Tinalconite holds a small stability field at 298 15K therefore making it stable at surface conditions

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