

Molecular Geometry of Covalent Molecules:  
Lewis Structures and the VSEPR Model

11 Practice Questions

Before beginning this experiment in the laboratory, you should be able to answer the following questions.

1. Draw the Lewis structure for each molecule.

2. What are the electron-pair geometry and molecular geometry for each of the following?



3. What are the electron-pair geometry and molecular geometry for each of the following?



4. What are the formal charges on each atom in the following molecules? (A formal charge of 0 is not shown.)



5. What are the hybrid orbitals on each atom in the following molecules? (A hybrid orbital is not shown.)



6. Calculate the bond angles.

(Do not use the VSEPR model, but use the hybrid orbital model.)

7. Calculate the formal charges of the atoms in  $CO_2$ ,  $NO_2$ , and  $NO_3^-$ .

$CO_2: C = 4 - (0 + 2) = -1$      $O = 6 - (1 + 1) = 1$

$CO_2: C = 4 - (4 + 0) = 0$      $O = 6 - (6) = 0$

$NO_2: N = 5 - (4 + 0) = 1$      $O = 6 - (6) = 0$

$NO_3^-: N = 5 - (1 + 6) = -1$      $O = 6 - (1 + 6) = -1$

# **Experiment 11 Molecular Geometries Covalent Molecules Answers**

**Arshad Iqbal**



## Experiment 11 Molecular Geometries Covalent Molecules Answers

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