

## Chemistry Unit 8 Worksheet 4

### Samples of Every Kind of Problem

*On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedure outlined in class. Be sure to circle your final answer.*

1. Calculate the number of moles of potassium chlorate,  $\text{KClO}_3$  (s), that must decompose to produce potassium chloride,  $\text{KCl}$  (s), and 1.8 moles of oxygen gas.

<b>Equation:</b>	$2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$			<b>1.2 moles KClO<sub>3</sub> needed</b>
<b>Before:</b>	1.2	0	0	
<b>Change:</b>	-1.2	+1.2	+1.8	
<b>After:</b>	0	1.2	1.8	

2. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 2.43 g of magnesium?

<b>Equation:</b>	$\text{Mg} + 2 \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$				<b>0.200 moles HCl needed</b>
<b>Before:</b>	0.100	?	0	0	
<b>Change:</b>	-0.100	-0.200 +	0.100 +	0.100	
<b>After:</b>	0	?	0.100	0.100	

3. Ethane,  $\text{C}_2\text{H}_6$  reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.20 moles of ethane?

<b>Equation:</b>	$2 \text{C}_2\text{H}_6 + 7 \text{O}_2 \rightarrow 4 \text{CO}_2 + 6 \text{H}_2\text{O}$			
<b>Before:</b>	2.20	xs	0	0
<b>Change:</b>	-2.20	-7.70	+4.40	+6.60
<b>After:</b>	0	xs	4.40	6.60

$$\text{O}_2 \text{ needed: } 7.70 \text{ mole} \times \frac{32.0 \text{ g}}{1 \text{ mole}} = 246 \text{ g O}_2 \text{ required}$$

4. Determine the mass of sodium nitrate produced when 0.73 g of nickel (II) nitrate reacts with sodium hydroxide according to the following equation:



$$0.73 \text{ g} \times \frac{1 \text{ mole}}{183 \text{ g}} = .0040 \text{ mole}$$

<b>Equation:</b>	$\text{Ni}(\text{NO}_3)_2 + 2 \text{NaOH} \rightarrow \text{Ni}(\text{OH})_2 + 2 \text{NaNO}_3$			
<b>Before:</b>	0.0040	XS	0	0
<b>Change:</b>	-0.0040	-0.0080	+0.0040	+0.0080
<b>After:</b>	0	XS	0.0040	0.0080

$$\text{NaNO}_3 \text{ produced: } 0.0080 \text{ mole} \times \frac{85.0 \text{ g}}{1 \text{ mole}} = 0.68 \text{ g}$$

# **Chemistry Solutions Introduction Name Key**

**Michael Mosher, Paul Kelter**



# Chemistry Solutions Introduction Name Key

Chemistry 2e Paul Flowers, Richard Langely, William R. Robinson, Klaus Hellmut Theopold, 2019-02-14 Chemistry 2e is designed to meet the scope and sequence requirements of the two semester general chemistry course The textbook provides an important opportunity for students to learn the core concepts of chemistry and understand how those concepts apply to their lives and the world around them The book also includes a number of innovative features including interactive exercises and real world applications designed to enhance student learning The second edition has been revised to incorporate clearer more current and more dynamic explanations while maintaining the same organization as the first edition Substantial improvements have been made in the figures illustrations and example exercises that support the text narrative Changes made in Chemistry 2e are described in the preface to help instructors transition to the second edition **Learning**

Chemistry 8 Solution Book (Year 2023-24) ,2024-01-02 *Learning Chemistry 7 Solution Book (Year 2023-24)* ,2024-01-02

*General, Organic, and Biochemistry* Katherine J. Denniston, Joseph J. Topping, Robert L. Caret, 2010-01-01 This edition is designed to help undergraduate health related majors and students of all other majors understand key concepts and appreciate the significant connections between chemistry health disease and the treatment of disease **Learning**

**Chemistry 6 Solution Book (Year 2023-24)** ,2024-01-02 **Student Solutions Manual for Introduction to General, Organic, and Biochemistry** Morris Hein, 2001-07-21 Experienced authors and educators Hein Pattison Best and Arena

recognize that both science and mathematics can be daunting subjects They skillfully anticipate areas of difficulty and pace the text accordingly Particular emphasis is placed on the understanding of how chemical principles relate to the readers lives and future careers **A Problem-Solving Approach to Aquatic Chemistry** James N. Jensen, 2023-01-12 A Problem Solving Approach to Aquatic Chemistry Enables civil and environmental engineers to understand the theory and application of aquatic equilibrium chemistry The second edition of A Problem Solving Approach to Aquatic Chemistry provides a detailed introduction to aquatic equilibrium chemistry calculation methods for systems at equilibrium applications of aquatic chemistry and chemical kinetics The text directly addresses two required ABET program outcomes in environmental engineering chemistry including stoichiometry equilibrium and kinetics and material and energy balances fate and transport of substances in and between air water and soil phases The book is very student centered with each chapter beginning with an introduction and ending with a summary that reviews the chapter s main points To aid in reader comprehension important terms are defined in context and key ideas are summarized Many thought provoking discussion questions worked examples and end of chapter problems are also included Each part of the text begins with a case study a portion of which is addressed in each subsequent chapter illustrating the principles of that chapter In addition each chapter has an Historical Note exploring connections with the people and cultures connected to topics in the text A Problem Solving Approach to Aquatic

Chemistry includes Fundamental concepts such as concentration units thermodynamic basis of equilibrium and manipulating equilibria Solutions of chemical equilibrium problems including setting up the problems and algebraic graphical and computer solution techniques Acid base equilibria including the concepts of acids and bases titrations and alkalinity and acidity Complexation including metals ligands equilibrium calculations with complexes and applications of complexation chemistry Oxidation reduction equilibria including equilibrium calculations graphical approaches and applications Gas liquid and solid liquid equilibrium with expanded coverage of the effects of global climate change Other topics including chemical kinetics of aquatic systems surface chemistry and integrative case studies For advanced senior undergraduates and first year graduate students in environmental engineering courses A Problem Solving Approach to Aquatic Chemistry serves as an invaluable learning resource on the topic with a variety of helpful learning elements included throughout to ensure information retention and the ability to apply covered concepts in practical settings

*An Introduction to Chemistry* Michael Mosher, Paul Kelter, 2023-03-18 This textbook is written to thoroughly cover the topic of introductory chemistry in detail with specific references to examples of topics in common or everyday life It provides a major overview of topics typically found in first year chemistry courses in the USA The textbook is written in a conversational question based format with a well defined problem solving strategy and presented in a way to encourage readers to think like a chemist and to think outside of the box Numerous examples are presented in every chapter to aid students and provide helpful self learning tools The topics are arranged throughout the textbook in a traditional approach to the subject with the primary audience being undergraduate students and advanced high school students of chemistry

*Chemistry* Theodore Lawrence Brown, H. Eugene LeMay, Bruce E. Bursten, Patrick Woodward, Catherine Murphy, 2017-01-03 NOTE This edition features the same content as the traditional text in a convenient three hole punched loose leaf version Books a la Carte also offer a great value this format costs significantly less than a new textbook Before purchasing check with your instructor or review your course syllabus to ensure that you select the correct ISBN Several versions of MyLab™ and Mastering™ platforms exist for each title including customized versions for individual schools and registrations are not transferable In addition you may need a Course ID provided by your instructor to register for and use MyLab and Mastering products For courses in two semester general chemistry Accurate data driven authorship with expanded interactivity leads to greater student engagement Unrivaled problem sets notable scientific accuracy and currency and remarkable clarity have made Chemistry The Central Science the leading general chemistry text for more than a decade Trusted innovative and calibrated the text increases conceptual understanding and leads to greater student success in general chemistry by building on the expertise of the dynamic author team of leading researchers and award winning teachers In this new edition the author team draws on the wealth of student data in Mastering™ Chemistry to identify where students struggle and strives to perfect the clarity and effectiveness of the text the art and the exercises while addressing student misconceptions and encouraging thinking about the practical real

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