

1. How many neutrons and protons are there in the nuclei of the following atoms:

- (a) ${}^7\text{Li}$,
- (b) ${}^{24}\text{Mg}$,
- (c) ${}^{135}\text{Xe}$,
- (d) ${}^{209}\text{Bi}$,
- (e) ${}^{222}\text{Rn}$?

(a) $Z = 3, N = 4$; (b) $Z = 12, N = 12$; (c) $Z = 54, N = 81$; (d) $Z = 83, N = 126$; (e) $Z = 86, N = 136$.

3. How many atoms are there in 10 g of ${}^{12}\text{C}$?

$10\text{ g} = 10/12$ mole and contains $(10/12) \times 0.6022 \times 10^{24} = 0.502 \times 10^{24}$ atoms.

5. When H_2 gas is formed from naturally occurring hydrogen, what percentages of the molecules have molecular weights of approximately 2, 3, and 4?

The molecules in question are ${}^1\text{H}{}^1\text{H}$, ${}^1\text{H}{}^2\text{H}$ or ${}^2\text{H}{}^1\text{H}$, and ${}^2\text{H}{}^2\text{H}$. Let $x(1)$ and $x(2)$ be the fractions of naturally-occurring hydrogen that are ${}^1\text{H}$ and ${}^2\text{H}$. Combining atoms pulled successively from a large volume of hydrogen, the fraction that has weight 2 is $x(1) \times x(1) = (0.99985)^2 = 0.9997 = 99.97\%$. The fraction with weight 4 is $x(2) \times x(2) = 2.25 \times 10^{-8} = 2.25 \times 10^{-6}\%$. Weight 3 can be formed in two ways. The fraction of 3 is $x(1)x(2) + x(2)x(1) = 2x(1)x(2) = 2 \times 0.99985 \times 0.00015 = 3.00 \times 10^{-4} = 0.03\%$. Note that the fraction that have molecular weights of either 2, 3, or 4 is $x^2(1) + 2x(1)x(2) + x^2(2) = (x(1) + x(2))^2 = 1$.

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