

1 From AQA Chemistry Unit 1 Foundation Chemistry CHEM1 May 2014 (Question 1).

a Table 1.1 shows some data about fundamental particles in an atom.

Table 1.1 Mass of fundamental particles

Particle	proton	neutron	electron
Mass / g	$1.6725 \times 10^{-24}$	$1.6725 \times 10^{-24}$	$0.0009 \times 10^{-24}$

i An atom of hydrogen can be represented as  ${}^1\text{H}$ . Use data from Table 1.1 to calculate the mass of this hydrogen atom.

$$1.6734 \times 10^{-27} \text{ kg}$$

(1 mark)

ii Which one of the following is a fundamental particle that would not be deflected by an electric field?

- A electron  
B neutron  
C proton

Write the correct letter, A, B, or C, on the line below.

B

(1 mark)

b A naturally occurring sample of the element boron has a relative atomic mass of 10.8. In this sample, boron exists as two isotopes,  ${}^{10}\text{B}$  and  ${}^{11}\text{B}$ .

i Calculate the percentage abundance of  ${}^{10}\text{B}$  in this naturally occurring sample of boron.

$$10.8 = (10)(x) + (11)(1-x)$$

$$10.8 = 10x + 11 - 11x$$

$$10.8 - 11 = 10x - 11x$$

$$-0.2 = -x$$

$$x = 0.2 \quad 0.2 \times 100 = 20\%$$

(2 marks)

# **Aqa Exam Style Questions Answers Chemistry A2**

**Alyn G. McFarland, Nora Henry**



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## **Table of Contents Aqa Exam Style Questions Answers Chemistry A2**

1. Understanding the eBook Aqa Exam Style Questions Answers Chemistry A2
  - The Rise of Digital Reading Aqa Exam Style Questions Answers Chemistry A2
  - Advantages of eBooks Over Traditional Books
2. Identifying Aqa Exam Style Questions Answers Chemistry A2
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Aqa Exam Style Questions Answers Chemistry A2
  - User-Friendly Interface
4. Exploring eBook Recommendations from Aqa Exam Style Questions Answers Chemistry A2
  - Personalized Recommendations
  - Aqa Exam Style Questions Answers Chemistry A2 User Reviews and Ratings
  - Aqa Exam Style Questions Answers Chemistry A2 and Bestseller Lists
5. Accessing Aqa Exam Style Questions Answers Chemistry A2 Free and Paid eBooks
  - Aqa Exam Style Questions Answers Chemistry A2 Public Domain eBooks
  - Aqa Exam Style Questions Answers Chemistry A2 eBook Subscription Services
  - Aqa Exam Style Questions Answers Chemistry A2 Budget-Friendly Options
6. Navigating Aqa Exam Style Questions Answers Chemistry A2 eBook Formats
  - ePub, PDF, MOBI, and More
  - Aqa Exam Style Questions Answers Chemistry A2 Compatibility with Devices
  - Aqa Exam Style Questions Answers Chemistry A2 Enhanced eBook Features
7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Aqa Exam Style Questions Answers Chemistry A2
  - Highlighting and Note-Taking Aqa Exam Style Questions Answers Chemistry A2
  - Interactive Elements Aqa Exam Style Questions Answers Chemistry A2
8. Staying Engaged with Aqa Exam Style Questions Answers Chemistry A2
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Aqa Exam Style Questions Answers Chemistry A2
9. Balancing eBooks and Physical Books Aqa Exam Style Questions Answers Chemistry A2
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Aqa Exam Style Questions Answers Chemistry A2
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Aqa Exam Style Questions Answers Chemistry A2
- Setting Reading Goals Aqa Exam Style Questions Answers Chemistry A2
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Aqa Exam Style Questions Answers Chemistry A2
- Fact-Checking eBook Content of Aqa Exam Style Questions Answers Chemistry A2
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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