

C $62.2 \text{ g HF} \times \frac{1}{178.5 \text{ g}} \times 0.348 = 1 \text{ HF}$
 $37.4 \text{ g Cl} \times \frac{1}{35.5 \text{ g}} = \frac{1.054}{.348} = 3 \text{ Cl}$

B $.05 \text{ mol Au}_2\text{S}_3 \times \frac{2 \text{ Au}}{1 \text{ Au}_2\text{S}_3} \times \frac{197 \text{ g}}{1 \text{ Au}} = 19.7$

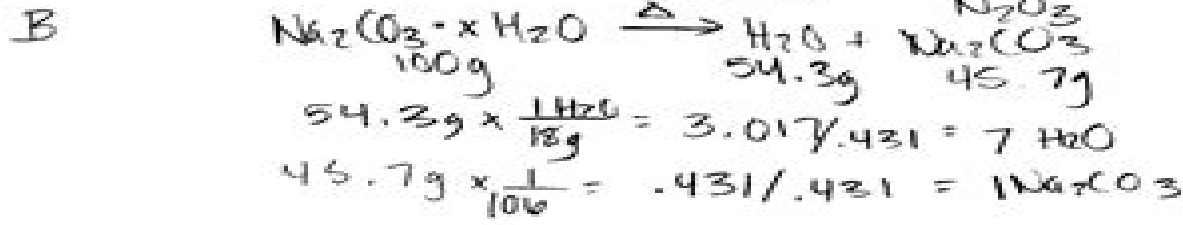
D $4 \text{ P} + 5 \text{ O}_2 \rightarrow \text{P}_4\text{O}_{10}$
 $14.2 \text{ g P}_4\text{O}_{10} \times \frac{1 \text{ mol}}{284 \text{ g}} \times \frac{5 \text{ O}_2}{1 \text{ P}_4\text{O}_{10}} = 0.25 \text{ mol}$

D $88 \text{ g CO}_2 \times \frac{1}{44 \text{ g}} \times \frac{1 \text{ C}}{1 \text{ C}_2} = 2 \text{ C}$
 $27 \text{ g H}_2\text{O} \times \frac{1}{18 \text{ g}} \times \frac{2 \text{ H}}{1} = 3 \text{ H}$
 $\text{C}_2\text{H}_3 \rightarrow \text{C}_4\text{H}_6$

E $24 \text{ g O} \times \frac{1 \text{ O}}{16 \text{ g}} \times \frac{1 \text{ Ca(NO}_3)_2}{6 \text{ oxygens}} \times \frac{164 \text{ g}}{1 \text{ Ca(NO}_3)_2} = 41 \text{ g}$

D $36.8 \text{ g N} \times \frac{1}{14} = 2.63 \text{ mol N} \rightarrow 1 \text{ N}_2 \times 2$

D $63.7 \text{ g O} \times \frac{1}{16} = 3.95 \text{ mol O} = 1.50 \times 2$



D $14.9 - 12.4 = 2.5$
 $d = \frac{19.6 \text{ ml}}{2.5} = 7.84 \text{ g} \rightarrow 7.8$

C Vol is affected by Temp $d = \frac{M}{L}$, $M = \frac{\text{mol}}{L}$
 a. $0.2 \times 2 = 0.4$ c. $0.2 \times 1 = 0.2$ m. $\frac{\text{mol}}{\text{g}}$
 v. $0.2 \times 2 = 0.4$ d. $0.2 \times 2 = 0.4$ e. $0.2 \times 3 = 0.6$

B PbBr_2 , FeBr_2 , CuBr_2

$2 \text{ HCl} \times (\text{Ca(OH)}_2 \rightarrow 2 \text{ H}_2\text{O} + \text{CaCl}_2$
 $.025 \text{ L} \times \frac{.15}{1 \text{ L}} \times \frac{2 \text{ HCl}}{1 \text{ Ca(OH)}_2} \times \frac{1 \text{ mol}}{.45 \text{ mol}} = .040$
 E

(b) E $m = \frac{3}{.05467} = 5$

.3 mol toluene
 .7 mol benzene $\times \frac{78.1}{1 \text{ mol}} = \frac{54.6}{1000}$

(7) D $b = \frac{x}{.05}$ $x = .3 \text{ mol} \times \frac{12}{1 \text{ r}}$
 $x = 52.2 \text{ g}$

(8) D
 (9) $M_1 V_1 = M_2 V_2$ A
 $12.2 (1) = 4x$
 $x = 3.05$

(10) $\Delta T = 1 \text{ km}$ C
 $m = \frac{\text{mol}}{\text{kg soln}}$ MM $\frac{1}{\text{mol}}$

Atomic

- (1) C 3 valence e⁻
- (2) C - no 2d
- (3) E - has d
- (4) B - halogen, -1, add 1 e⁻
- (5) B - Alkali earth, +2

(6) E
 (7) D $1s^2 2s^2 2p^4 3s^2 3p^2$
 Valence e⁻ = 3

(8) C
 (9) E
 (10) B
 (11) A
 (12) A
 (13) C

(15) D
 (16) C sp^1
 $n=5, p=1, l$
 $ML = \text{Fl to } 1 \text{ mg}$

(17) C $n=2, l=3$ (Mn²⁺) all same

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