

1. Assign oxidation numbers to all atoms in the equation.

2. Identify the atoms that are oxidized and the atoms that are reduced.

3. Determine the change in oxidation number for the atoms that are oxidized and for the atoms that are reduced.

4. Make the change in oxidation numbers equal in magnitude by adjusting coefficients in the equation.

5. If necessary, use the conventional method to balance the remainder of the equation.

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Glencoe Chemistry Chapter 19 :