

- ii.) 234000
- iii.) 8008
- iv.) 500.0
- v.) 6.0012

Q7. Define Avogadro number and mole. What is their importance?

Q8. In the combustion of methane in the air, what is the limiting reactant and why?

Q9. Is the law of constant composition true for all types of compounds? Explain why or why or why not.

Q10. Sulphuric acid is generally available in the market as an 18.0 M solution. How would you prepare 250 mL of 0.50 M aqueous H_2SO_4 ?

Q11. Calculate:

- a.) Mass of 1.5 g atoms of calcium
- b.) Gram atoms in 12.8 g of oxygen

Q12. The vapour density of a mixture of NO_2 and N_2O_4 is 38.3 at $26.7^\circ C$. Calculate the number of moles of NO_2 in 100 g of the mixture.

Q13. Balance the following equations by hit and trial method:

- a.) $KMnO_4 + HCl \rightarrow KCl + MnCl_2 + H_2O + Cl_2$
- b.) $Mg_3N_2 + H_2O \rightarrow Mg(OH)_2 + NH_3$
- c.) $Al_4C_3 + H_2O \rightarrow Al(OH)_3 + CH_4$
- d.) $FeS_2 + O_2 \rightarrow Fe_2O_3 + SO_2$
- e.) $K_2Cr_2O_7 + H_2SO_4 \rightarrow K_2SO_4 + Cr_2(SO_4)_3 + H_2O + O_2$

Q14. Lithium oxide removes water from the air according to the reaction:



If 72 kg of water is to be removed and 35 kg of Li_2O is available,

- i.) Which reactant is limiting?
- ii.) How many kg of excess reactant is left?

Q15. A compound (molecular mass = 246) has the following data:

Element	% Composition	Relative No. of Atoms
A	9.76	0.406
B	13.01	0.406
C	26.01	1.625
D	51.22	2.846

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C Cleary



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