

DAT General Chemistry Equation Sheet

General Concepts

Density: $(d = m/v)$

Dilutions: $M_1V_1 = M_2V_2$ (or $C_1V_1 = C_2V_2$)

% Error Calculations

$$\% \text{ Error} = \frac{\text{Actual} - \text{theoretical}}{\text{theoretical}} \times 100$$

Spectrophotometer: $Abs = \epsilon cl$

Abs = Absorbance

ϵ = molar extinction coefficient (molar absorptivity)

l = path length

Atomic Structure

Absorption/Emission Line Spectra: $\Delta E = E_{\text{photon}}$

Energy of a photon: $E_{\text{photon}} = hf = \frac{hc}{\lambda}$

$$c = 3.0 \times 10^8 \text{ m/s}$$

$$h = 6.63 \times 10^{-34} \text{ J} \cdot \text{s}$$

Photoelectric Effect: $KE_{\text{electron}} = E_{\text{photon}} - \phi$

ϕ = work function (minimum energy to ionize electron)

Chemical Kinetics and Equilibrium

General Rate Law: $\text{rate} = k[A]^m[B]^n$

Arrhenius Equation: $k = Ae^{-E_a/RT}$ E_a = activation energy

$R = 0.0821 \text{ L atm/mol K}$

Equilibrium Constant:

$$K_c = \frac{[\text{products}]}{[\text{reactants}]} \quad K_p = \frac{P_{\text{products}}}{P_{\text{reactants}}}$$

Solutions

Concentration

$$\text{Molarity (M)} = \frac{\text{moles}_{\text{solute}}}{L_{\text{solution}}}$$

$$\text{molality (m)} = \frac{\text{moles}_{\text{solute}}}{kg_{\text{solvent}}}$$

Henry's Law: $P_A = k_H[A]$

Colligative Properties

Freezing Point Depression:

$$\Delta T_f = -iK_f m$$

Boiling Point Elevation:

$$\Delta T_b = iK_b m$$

Raoult's Law (* = pure):

$$P_A = X_A P_A^*$$

Osmotic Pressure:

$$\pi = iMRT$$

K_f/K_b = F.P./B.P. depression/elevation constant

X_A = mole fraction of A

i = van't Hoff Factor

$R = 0.0821 \text{ L atm/mol K}$

M = molarity, m = molality

Gases/Gas Laws

Pressure: $P = \text{Force/Area}$

Average Kinetic Energy:

$K.E._{avg} = \frac{3}{2} RT$ (per mole of gas)

Root-mean-square speed: $v = \sqrt{\frac{3RT}{M_m}}$

Gas Laws

Ideal Gas Law:

$$PV = nRT$$

Boyle's Law:

$$P \propto \frac{1}{V}$$

Charles' Law:

$$V \propto T$$

Avogadro's Law:

$$V \propto n$$

Combined Gas Law:

$$\frac{P_1 V_1}{n_1 T_1} = \frac{P_2 V_2}{n_2 T_2}$$

Dalton's Law of Partial Pressures:

$$P_{\text{Tot}} = P_A + P_B + \dots$$

$$P_A = X_A \cdot P_{\text{Tot}}$$

$$X_A = \text{mol fraction } A$$

$R = 0.0821 \text{ L atm/mol K}$

n = moles

M_m = molar mass

r = rate of effusion

a and b are constants specific to a given gas.

Grahams Law of Effusion:

$$\frac{r_1}{r_2} = \sqrt{\frac{M_{m2}}{M_{m1}}}$$

Acid-Base

pH Scale

$$pH = -\log[H^+]$$

$$[H^+] = 10^{-pH}$$

$$pOH = -\log[OH^-]$$

$$[OH^-] = 10^{-pOH}$$

$$pH + pOH = 14$$

$$[H^+][OH^-] = 1 \times 10^{-14}$$

Weak Acids:

$$K_a = \frac{[H_3O^+][A^-]}{[HA]}$$

$$[H^+] = \sqrt{K_a[HA]}$$

Weak Bases:

$$K_b = \frac{[OH^-][HA]}{[A^-]}$$

$$[OH^-] = \sqrt{K_b[B]}$$

$$K_a + K_b = K_w$$

Buffer: $pH = pK_a + \log \frac{[A^-]}{[HA]}$

Titration/Neutralization: $n_a M_a V_a = n_b M_b V_b$

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Various



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