

Chapter 1 The atomic nature of matter

Section 1.1 Nanomaterials and nanoparticles

Worked example: Try yourself 1.1.1

CONVERTING KILOMETRES TO NANOMETRES

Convert 2.43 km into nanometres.	
Thinking	Working
Write the length in metres.	$2.43 \text{ km} = 2430 \text{ m}$
Write the length in scientific notation.	$2430 \text{ m} = 2.43 \times 10^3 \text{ m}$
Multiply by 10^9 to convert to nanometres. (Hint: to do this, you just need to add 9 to the index on the number 10.)	$(2.43 \times 10^3) \times 10^9 = 2.43 \times 10^{3+9}$ $= 2.43 \times 10^{12} \text{ nm}$

1.1 Key questions

1 $8.35 \text{ cm} = 0.0835 \text{ m}$
 $0.0835 \text{ m} = 8.35 \times 10^{-2} \text{ m}$
 $(8.35 \times 10^{-2}) \times 10^9 = 8.35 \times 10^{-2+9}$
 $= 8.35 \times 10^7 \text{ nm}$

2 **a** $1.35 \text{ cm} = 1.35 \times 10^{-2} \text{ m}$
 $1.35 \times 10^{-2} \text{ m} \times 10^9 = 1.35 \times 10^7 \text{ nm}$
b $4.2 \text{ mm} = 4.2 \times 10^{-3} \text{ m}$
 $4.2 \times 10^{-3} \times 10^9 = 4.2 \times 10^6 \text{ nm}$

3 $0.050 \text{ mm} = 5 \times 10^{-5} \text{ m}$
 $5 \times 10^{-5} \times 10^9 = 5.0 \times 10^4 \text{ nm}$
 $5.0 \times 10^4 \div 20 = 2500 \text{ times larger}$

4 **a**

Shape	Surface area	Volume	$\frac{\text{Surface area}}{\text{Volume}}$
Cube $2 \text{ cm} \times 2 \text{ cm} \times 2 \text{ cm}$	24 cm^2	8 cm^3	3
Sphere of radius 1.38 cm	23.9 cm^2	11 cm^3	2.17
Tube (cylinder) Radius of 1 cm Height of 2 cm	19 cm^2	6.3 cm^3	3

- b** Sphere
c Sphere
d A large SA : V ratio is required for rapid cooling (e.g. on a hot day people spread out). A small SA : V ratio is required to minimise loss of heat when cold. Hence, people huddle or wrap their arms around themselves to keep warm.
- 5 There are two reasons why their size makes nanoparticles useful for transporting medicine. First, their small size makes it easy for the nanoparticles to get into the body by inhalation, injection or absorption through the skin. Second, their small size gives nanoparticles a large surface area compared to their volume. This means that a large number of molecules can be adsorbed onto the surface of a relatively small volume of nanoparticles.

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Richard Bailey



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