

Specific Heat Capacity – Worksheet 1

SAC
1 of 4

1. In kJ, how much heat is required to raise the temperature of 1.23 kg of water from 23.6 °C to 84.9 °C? (317 kJ)

$$Q = ?$$

$$m = 1.23 \text{ kg}$$

$$c = 4200 \text{ J/kg}^\circ\text{C}$$

$$T_i = 23.6^\circ\text{C}$$

$$T_f = 84.9^\circ\text{C}$$

$$Q = mc\Delta T$$

$$Q = 1.23 \text{ kg} (4200 \text{ J/kg}^\circ\text{C}) (84.9^\circ\text{C} - 23.6^\circ\text{C})$$

$$Q = 316675.8 \text{ J}$$

$$Q = 317 \text{ kJ}$$

2. How much heat is needed to warm up 3.6 kg of hydrogen gas from 12.0 °C to 48.0 °C?

$$Q = ?$$

$$m = 3.6 \text{ kg}$$

$$T_i = 12.0^\circ\text{C}$$

$$T_f = 48.0^\circ\text{C}$$

$$c = 14400 \text{ J/kg}^\circ\text{C}$$

$$Q = mc\Delta T$$

$$Q = 3.6 \text{ kg} (14400 \text{ J/kg}^\circ\text{C}) (48.0^\circ\text{C} - 12.0^\circ\text{C})$$

$$Q = 1.8 \times 10^6 \text{ J}$$

3. How much heat is released when a 260.0 g plate of copper cools from $5.30 \times 10^2 \text{ K}$ to 20.0 °C? (-24.0 kJ)

$$Q = ?$$

$$m = 260.0 \text{ g} = 0.2600 \text{ kg}$$

$$c = 390 \text{ J/kg}^\circ\text{C}$$

$$T_i = 5.30 \times 10^2 \text{ K}$$

$$T_f = 20.0^\circ\text{C} = 293 \text{ K}$$

$$Q = mc\Delta T$$

$$Q = 0.2600 \text{ kg} (390 \text{ J/kg}^\circ\text{C}) (293 \text{ K} - 5.30 \times 10^2 \text{ K})$$

$$Q = -24071.8 \text{ J}$$

$$Q = -24.0 \text{ kJ}$$

4. 3.4 kJ of heat is added to a 1.4 kg rod of uranium. What is the change in temperature the rod undergoes?

$$Q = 3.4 \text{ kJ} = 3400 \text{ J}$$

$$m = 1.4 \text{ kg}$$

$$\Delta T = ?$$

$$c = 210 \text{ J/kg}^\circ\text{C}$$

$$Q = mc\Delta T$$

$$\frac{Q}{mc} = \Delta T$$

$$\frac{3400 \text{ J}}{1.4 \text{ kg} (210 \text{ J/kg}^\circ\text{C})} = \Delta T$$

$$12^\circ\text{C} \approx \Delta T$$

5. 24 000 J of heat is released when 640 g of silver drops in temperature. If the silver was initially at -10.0 °C, what is the final temperature of the silver? (-170 °C)

$$Q = -24000 \text{ J}$$

$$m = 640 \text{ g} = 0.640 \text{ kg}$$

$$T_i = -10.0^\circ\text{C}$$

$$T_f = ?$$

$$c = 230 \text{ J/kg}^\circ\text{C}$$

$$Q = mc(T_f - T_i)$$

$$\frac{Q}{mc} = T_f - T_i$$

$$\frac{Q}{mc} + T_i = T_f$$

$$\frac{-24000 \text{ J}}{0.640 \text{ kg} (230 \text{ J/kg}^\circ\text{C})} + (-10.0^\circ\text{C}) = T_f$$

$$-170^\circ\text{C} \approx T_f$$

Thermal Energy Practice Problems With Answer Key

**V.K. Sally ,R.N. Das Gupta,Goyal
Brothers Prakashan**



Thermal Energy Practice Problems With Answer Key

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Table of Contents Thermal Energy Practice Problems With Answer Key

1. Understanding the eBook Thermal Energy Practice Problems With Answer Key
 - The Rise of Digital Reading Thermal Energy Practice Problems With Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermal Energy Practice Problems With Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermal Energy Practice Problems With Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermal Energy Practice Problems With Answer Key
 - Personalized Recommendations
 - Thermal Energy Practice Problems With Answer Key User Reviews and Ratings
 - Thermal Energy Practice Problems With Answer Key and Bestseller Lists
5. Accessing Thermal Energy Practice Problems With Answer Key Free and Paid eBooks
 - Thermal Energy Practice Problems With Answer Key Public Domain eBooks
 - Thermal Energy Practice Problems With Answer Key eBook Subscription Services
 - Thermal Energy Practice Problems With Answer Key Budget-Friendly Options
6. Navigating Thermal Energy Practice Problems With Answer Key eBook Formats

- ePub, PDF, MOBI, and More
 - Thermal Energy Practice Problems With Answer Key Compatibility with Devices
 - Thermal Energy Practice Problems With Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermal Energy Practice Problems With Answer Key
 - Highlighting and Note-Taking Thermal Energy Practice Problems With Answer Key
 - Interactive Elements Thermal Energy Practice Problems With Answer Key
 8. Staying Engaged with Thermal Energy Practice Problems With Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermal Energy Practice Problems With Answer Key
 9. Balancing eBooks and Physical Books Thermal Energy Practice Problems With Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermal Energy Practice Problems With Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Thermal Energy Practice Problems With Answer Key
 - Setting Reading Goals Thermal Energy Practice Problems With Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Thermal Energy Practice Problems With Answer Key
 - Fact-Checking eBook Content of Thermal Energy Practice Problems With Answer Key
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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