

5. The relative molecular mass of lead (II) trioxonitrate (V) is (Pb = 207, N = 14, O = 16)
- 237.
 - 267.
 - 269.
 - 323.
 - 331.
6. The empirical formula of a hydrocarbon containing 0.12 mole of carbon and 0.36 mole of hydrogen is
- CH₂.
 - CH₃.
 - C₂H₂.
 - C₂H₄.
 - C₂H₆.
7. How many valence electrons does an oxygen atom have?
- 2
 - 6
 - 8
 - 16
 - 32
8. How many orbitals are associated with the p-sub orbital energy level?
- 2
 - 3
 - 4
 - 5
 - 6
9. The following are precautions during titration **except**
- avoid parallax error in reading meniscus level.
 - avoid spillage of solution from the pipette when transferring.
 - blow the last drop at the tip of the pipette.
 - remove air bubbles from the burette or pipette.
 - rinse the burette with the solution of the acid.
10. What is the function of conc. H₂SO₄ in the reaction below?
- $$\text{C}_{12}\text{H}_{22}\text{O}_{11} \xrightarrow{\text{Conc. H}_2\text{SO}_4} 12\text{C} + 11\text{H}_2\text{O}$$
- It acts as a/an
- catalyst.
 - dehydrating agent.
 - oxidizing agent.
 - reducing agent.
 - strong acid.
11. The crystal layers of graphite are held together by
- covalent bond.
 - electrovalent bond.
 - electrostatic force.
 - hydrogen bond.
 - van der Waals' forces.
12. An element Q forms a compound QCl₅. In which group of the periodic table is Q?
- I
 - III
 - V
 - VI
 - VII

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JA Banks



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Table of Contents Chemistry Question And Answer Theory Objectives

1. Understanding the eBook Chemistry Question And Answer Theory Objectives
 - The Rise of Digital Reading Chemistry Question And Answer Theory Objectives
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Question And Answer Theory Objectives
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Question And Answer Theory Objectives

- User-Friendly Interface
- 4. Exploring eBook Recommendations from Chemistry Question And Answer Theory Objectives
 - Personalized Recommendations
 - Chemistry Question And Answer Theory Objectives User Reviews and Ratings
 - Chemistry Question And Answer Theory Objectives and Bestseller Lists
- 5. Accessing Chemistry Question And Answer Theory Objectives Free and Paid eBooks
 - Chemistry Question And Answer Theory Objectives Public Domain eBooks
 - Chemistry Question And Answer Theory Objectives eBook Subscription Services
 - Chemistry Question And Answer Theory Objectives Budget-Friendly Options
- 6. Navigating Chemistry Question And Answer Theory Objectives eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Question And Answer Theory Objectives Compatibility with Devices
 - Chemistry Question And Answer Theory Objectives Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Question And Answer Theory Objectives
 - Highlighting and Note-Taking Chemistry Question And Answer Theory Objectives
 - Interactive Elements Chemistry Question And Answer Theory Objectives
- 8. Staying Engaged with Chemistry Question And Answer Theory Objectives
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Question And Answer Theory Objectives
- 9. Balancing eBooks and Physical Books Chemistry Question And Answer Theory Objectives
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Question And Answer Theory Objectives
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Question And Answer Theory Objectives
 - Setting Reading Goals Chemistry Question And Answer Theory Objectives

- Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Question And Answer Theory Objectives
 - Fact-Checking eBook Content of Chemistry Question And Answer Theory Objectives
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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