

CHAPTER 8 REVIEW

Chemical Equations and Reactions

SHORT ANSWER Answer the following questions in the space provided.

2. _____ According to the law of conservation of mass, the total mass of the reacting substances is
- always more than the total mass of the products.
 - always less than the total mass of the products.
 - sometimes more and sometimes less than the total mass of the products.
 - always equal to the total mass of the products.**

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 2. In the reaction described by the word equation

$$\text{sodium} + \text{water} \rightarrow \text{sodium hydroxide} + \text{hydrogen gas}$$
 - sodium hydroxide is a product.
 - sodium is a product.
 - water is a reactant.
 - Both (a) and (c)**
- _____ 3. Which of the following indicates that a chemical equation is balanced?
 - The numbers of atoms of each element are the same on both sides of the equation.**
 - All of the coefficients are the same.
 - The numbers of molecules on each side of the equation are equal.
 - The sums of the coefficients on each side of the equation are equal.
- _____ 4. From a complete and correctly written chemical equation, you can obtain the
 - chemical formulas of the reactants and products.
 - relative amounts of the reactants and products.
 - physical states of the reactants and products.
 - All of the above**
- _____ 5. To balance a chemical equation, you must
 - adjust the subscripts.
 - adjust the number of products.
 - adjust the number of reactants.
 - adjust the coefficients.**
- _____ 6. The formation of a solid product in a chemical reaction is represented by the symbol
 - (aq).
 - (g).
 - (l).
 - (s).**
- _____ 8. In the expression 3CO_2 , the numbers 3 and 2 are, respectively,
 - a subscript and a coefficient.
 - a coefficient and a subscript.**
 - two subscripts.
 - two coefficients.
- _____ 9. What is the total number of molecules that is represented by the following chemical equation? (Add up the coefficients.)

$$\text{NH}_4\text{NO}_2 \rightarrow \text{N}_2 + 2\text{H}_2\text{O}$$
 - 3
 - 4**
 - 8
 - 16
- _____ 10. Which equation below violates the law of conservation of mass?
 - $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 - $\text{KCl} + \text{Br} \rightarrow \text{KBr} + \text{Cl}_2$**
 - $2\text{Fe}_2\text{O}_3 + 3\text{C} \rightarrow 4\text{Fe} + 3\text{CO}_2$
 - $\text{Na}_2\text{CO}_3 + 2\text{HCl} \rightarrow 2\text{NaCl} + \text{H}_2\text{O} + \text{CO}_2$
- _____ 11. Which of the following equations represents the balanced equation for the reaction of iron and oxygen?
 - $2\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$
 - $\text{Fe} + 3\text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$
 - $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$**
 - $3\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$
- _____ 12. Which coefficients correctly balance the formula equation below?

$$\text{KClO}_3(s) \rightarrow \text{KCl}(s) + \text{O}_2(g)$$
 - 1, 1, 1
 - 1, 1, 3
 - 2, 2, 3**
 - 2, 1, 1
- _____ 13. Which equation is *not* balanced?
 - $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 - $4\text{H}_2 + 2\text{O}_2 \rightarrow 4\text{H}_2\text{O}$
 - $\text{H}_2 + \text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}$
 - $2\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$**
- _____ 14. In which type of reaction do two or more compounds react to form one product?
 - synthesis (AKA combination)**
 - decomposition
 - combustion
 - single-replacement
- _____ 15. Which of the following is true regarding single-replacement reactions?
 - Only one product is formed.
 - There is only one reagent.
 - One element replaces another in a compound.**

Chapter Assessment Chemical Reactions Answers

**Ryan Dupont, Kumar Ganesan, Louis
Theodore**



Chapter Assessment Chemical Reactions Answers

Chemistry 2e Paul Flowers, Richard Langely, William R. Robinson, Klaus Hellmut Theopold, 2019-02-14 Chemistry 2e is designed to meet the scope and sequence requirements of the two semester general chemistry course. The textbook provides an important opportunity for students to learn the core concepts of chemistry and understand how those concepts apply to their lives and the world around them. The book also includes a number of innovative features including interactive exercises and real world applications designed to enhance student learning. The second edition has been revised to incorporate clearer, more current and more dynamic explanations while maintaining the same organization as the first edition. Substantial improvements have been made in the figures, illustrations, and example exercises that support the text narrative. Changes made in Chemistry 2e are described in the preface to help instructors transition to the second edition.

Basic Concepts of Chemistry, Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

Holt McDougal Modern Chemistry Mickey Sarquis, 2012 A Framework for K-12 Science Education National Research Council, Division of Behavioral and Social Sciences and Education, Board on Science Education, Committee on a Conceptual Framework for New K-12 Science Education Standards, 2012-02-28 Science, engineering, and technology permeate nearly every facet of modern life and hold the key to solving many of humanity's most pressing current and future challenges. The United States' position in the global economy is declining in part because U.S. workers lack fundamental knowledge in these fields. To address the critical issues of U.S. competitiveness and to better prepare the workforce, A Framework for K-12 Science Education proposes a new approach to K-12 science education that will capture students' interest and provide them with the necessary foundational knowledge in the field. A Framework for K-12 Science Education outlines a broad set of expectations for students in science and engineering in grades K-12. These expectations will inform the development of new standards for K-12 science education and subsequently revisions to

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World of Chemistry Steven S. Zumdahl, Susan L. Zumdahl, Donald J. DeCoste, 2006-08 Our high school chemistry program has been redesigned and updated to give your students the right balance of concepts and applications in a program that provides more active learning more real world connections and more engaging content A revised and enhanced text designed especially for high school helps students actively develop and apply their understanding of chemical concepts Hands on labs and activities emphasize cutting edge applications and help students connect concepts to the real world A new captivating design clear writing style and innovative technology resources support your students in getting the most out of their textbook Publisher

Basic Concepts of Chemistry Leo J. Malone, Theodore Dolter, 2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features Each section

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L3 Karen Morrison, 2008 **Pearson Chemistry** Antony C. Wilbraham, Dennis D. Staley, Michael S. Matta, Edward L. Waterman, 2012-01-01 Glencoe Chemistry: Matter and Change, Student Edition McGraw-Hill Education, 2016-06-15

Thermal Safety of Chemical Processes Francis Stoessel, 2021-05-24 Vollständig bearbeitete aktualisierte 2. Auflage mit 12 neuen Kapiteln und den aktuellen IUPAC Normen Beschäftigt sich mit thermischen Risiken in der chemischen Prozessindustrie Jedes Kapitel beginnt mit einer Fallstudie und den aus Zwischenfällen gewonnenen Erkenntnissen

Pollution Prevention Ryan Dupont, Kumar Ganesan, Louis Theodore, 2016-11-18 This new edition has been revised throughout and adds several sections including lean manufacturing and design for the environment low impact development and green infrastructure green science and engineering and sustainability It presents strategies to reduce waste from the source of materials development through to recycling and examines the basic concepts of the physical chemical and biological properties of different pollutants It includes case studies from several industries such as pharmaceuticals pesticides metals electronics petrochemicals refineries and more It also addresses the economic considerations for each pollution prevention approach Elements of Chemical Reaction Engineering H. Scott Fogler, 1999 The fourth edition of Elements of Chemical Reaction Engineering is a completely revised version of the book It combines authoritative coverage of the principles of chemical reaction engineering with an unsurpassed focus on critical thinking and creative problem solving employing open ended questions and stressing the Socratic method Clear and organized it integrates text visuals and computer simulations to help readers solve even the most challenging problems through reasoning rather than by memorizing equations **BOOK JACKET**

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web conclusion the purpose of the experiment was to identify the unknown solid acid 17 by calculating its molar mass this was done by titration of the acid with sodium hydroxide in order to find the volume of sodium hydroxide required to reach the end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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 $HCl(aq) + NaOH(aq) \rightarrow NaCl(aq) + H_2O(l)$ one mole of HCl would be fully neutralized by one mole of $NaOH$

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web conclusion the identification of an unknown solid acid the purpose of this experiment was to determine the standardization of sodium hydroxide $NaOH$ and oxalic acid $H_2C_2O_4 \cdot 2H_2O$ and the molar mass of an unknown solid acid using the molar mass of the standardized sodium hydroxide solution

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak pK_a of an unknown acid or pK_b of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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Table of Contents Chapter Assessment Chemical Reactions Answers

1. Understanding the eBook Chapter Assessment Chemical Reactions Answers
 - The Rise of Digital Reading Chapter Assessment Chemical Reactions Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter Assessment Chemical Reactions Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter Assessment Chemical Reactions Answers

- User-Friendly Interface
- 4. Exploring eBook Recommendations from Chapter Assessment Chemical Reactions Answers
 - Personalized Recommendations
 - Chapter Assessment Chemical Reactions Answers User Reviews and Ratings
 - Chapter Assessment Chemical Reactions Answers and Bestseller Lists
- 5. Accessing Chapter Assessment Chemical Reactions Answers Free and Paid eBooks
 - Chapter Assessment Chemical Reactions Answers Public Domain eBooks
 - Chapter Assessment Chemical Reactions Answers eBook Subscription Services
 - Chapter Assessment Chemical Reactions Answers Budget-Friendly Options
- 6. Navigating Chapter Assessment Chemical Reactions Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter Assessment Chemical Reactions Answers Compatibility with Devices
 - Chapter Assessment Chemical Reactions Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter Assessment Chemical Reactions Answers
 - Highlighting and Note-Taking Chapter Assessment Chemical Reactions Answers
 - Interactive Elements Chapter Assessment Chemical Reactions Answers
- 8. Staying Engaged with Chapter Assessment Chemical Reactions Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter Assessment Chemical Reactions Answers
- 9. Balancing eBooks and Physical Books Chapter Assessment Chemical Reactions Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter Assessment Chemical Reactions Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter Assessment Chemical Reactions Answers
 - Setting Reading Goals Chapter Assessment Chemical Reactions Answers

- Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter Assessment Chemical Reactions Answers
 - Fact-Checking eBook Content of Chapter Assessment Chemical Reactions Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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