

TEST DATE: \_\_\_\_\_

**Unit 8 Review Sheet:****Properties of Water, Solutions, Concentration, Acids and Bases****PROPERTIES OF WATER**

- Define the following terms: polarity, surface tension, vapor pressure, specific heat, and capillary action.
- Draw four water molecules. Label the types of bonds, oxygen atoms, hydrogen atoms, and respective charges on the atoms.
- Is water polar or non-polar? Explain.
- What type of bond forms between individual molecules of water?
- What type of bond forms between each hydrogen and the oxygen within a water molecule?
- Why is water considered the universal solvent?
- What are the special properties of water and why do they occur?
- Explain why solid ice is less than liquid water with regard to particle arrangement.
- Why does sugar dissolve in water, but oil does not?

**SOLUTIONS**

- Define the following terms: solution, solvent, solute, dilute, concentrated, dissociate, solubility saturated, supersaturated and unsaturated.
- Give an example of solid, liquid, and gas solution. Identify the solute and solvent.
- What is a solution? Give an example of a solution, and an example of a mixture that is not a solution.
- Describe how temperature and pressure affects the solubility of solid and gas solutes in water.
- A glass of water has 10g of sugar dissolved in it. If more sugar can be added to dissolve in the water, is the solution unsaturated, saturated, or supersaturated?
- How do intermolecular forces affect solvation?
- Why do vinegar and oil not mix?
- On the line at the left, write the letter of the definition that best matches each term.
 

_____ solution	a. measure of how much solute will dissolve in a solvent
_____ solute	b. solution with water as the solvent
_____ solvent	c. substance that is dissolved in a solution
_____ soluble	d. substance that dissolves in water to form a solution that conducts electricity
_____ alloy	e. solid solution containing two or more metals
_____ aqueous	f. homogeneous mixture of two or more substances in a single physical state
_____ electrolyte	g. substance that does the dissolving in a solution
_____ solubility	h. Liquids that are insoluble in each other are considered this
_____ immiscible	i. capable of being dissolved

**CALCULATING CONCENTRATION**

- If 8.7 g of  $\text{Na}_2\text{CO}_3$  is dissolved in 800 mL of water, what is the **molarity** of the solution? (1 mL of water = 1 g of water)
- How many **grams** of  $\text{MgCl}_2$  would be needed to make 1.5 L of a **0.40 M** solution?
- What is the **molarity** of a bleach solution containing 9.5g of  $\text{NaOCl}$  per liter of bleach?
- What is the **percent by volume** of ethanol in a solution that contains 35 mL of ethanol dissolved in 115 mL of water?
- Calculate the **molarity** of 1.60L of a solution containing 1.55 g of dissolved  $\text{KBr}$ .
- If .5mL of blood are added to 10.0 L of water what is the concentration in **PPM**?
- A solution is made up of 123 g  $\text{NaOH}$  and 289 g water. The total volume is 300.0 mL. Determine the following:
 

a. Moles of $\text{NaOH}$	d. Mole fraction
b. Moles of $\text{H}_2\text{O}$	e. Molarity
c. Mass Percent	f. PPM
- What is the **molarity** of a solution that contains 63.0 g  $\text{HNO}_3$  in 0.500 kg  $\text{H}_2\text{O}$ ?
- What **mass of water** is required to dissolve 100. g  $\text{NaCl}$  to prepare a **1.50m** solution?

# Chemistry Review Sheet Unit 8 Answer Key

**Mark Bishop**



# Chemistry Review Sheet Unit 8 Answer Key

**Chemistry 2e** Paul Flowers, Richard Langely, William R. Robinson, Klaus Hellmut Theopold, 2019-02-14 Chemistry 2e is designed to meet the scope and sequence requirements of the two semester general chemistry course The textbook provides an important opportunity for students to learn the core concepts of chemistry and understand how those concepts apply to their lives and the world around them The book also includes a number of innovative features including interactive exercises and real world applications designed to enhance student learning The second edition has been revised to incorporate clearer more current and more dynamic explanations while maintaining the same organization as the first edition Substantial improvements have been made in the figures illustrations and example exercises that support the text narrative Changes made in Chemistry 2e are described in the preface to help instructors transition to the second edition

**MCAT Biology Review**, 2010 The Princeton Review's MCAT Biology Review contains in depth coverage of the challenging biology topics on this important test

**Quantities, Units and Symbols in Physical Chemistry** International Union of Pure and Applied Chemistry. Physical and Biophysical Chemistry Division, 2007 Prepared by the IUPAC Physical Chemistry Division this definitive manual now in its third edition is designed to improve the exchange of scientific information among the readers in different disciplines and across different nations This book has been systematically brought up to date and new sections added to reflect the increasing volume of scientific literature and terminology and expressions being used The Third Edition reflects the experience of the contributors with the previous editions and the comments and feedback have been integrated into this essential resource This edition has been compiled in machine readable form and will be available online

**The Foolish Almanak** Theodore Roosevelt,

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Resources in Education ,2001     *Holt McDougal Modern Chemistry* Mickey Sarquis,2012     **Glencoe Chemistry: Matter and Change, Student Edition** McGraw-Hill Education,2016-06-15     **Teaching Reading in Science** Mary Lee Barton,Deborah L. Jordan,2001 This book suggests that the reading of science text and textbooks requires the same thinking skills that are involved in a hands on science activity and presents the latest research on reading and learning science This supplement also includes suggestions on how to implement appropriate science readings into instruction and help students learn how to construct meaning from science textbooks Contents include 1 Three Interactive Elements of Reading 2 Strategic Processing 3 Strategic Teaching 4 Six Assumptions about Learning and 5 Reading Strategies Contains 54 references YDS

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Introduces the major elements of semantics in a simple step by step fashion Sections of explanation and examples are followed by practice exercises with answers and comment provided *World of Chemistry* Steven S. Zumdahl, Susan L. Zumdahl, Donald J. DeCoste, 2006-08 Our high school chemistry program has been redesigned and updated to give your students the right balance of concepts and applications in a program that provides more active learning more real world connections and more engaging content A revised and enhanced text designed especially for high school helps students actively develop and apply their understanding of chemical concepts Hands on labs and activities emphasize cutting edge applications and help students connect concepts to the real world A new captivating design clear writing style and innovative technology resources support your students in getting the most out of their textbook Publisher [El-Hi Textbooks in Print](#), 1984

**An Introduction to Chemistry** Mark Bishop, 2002 This book teaches chemistry at an appropriate level of rigor while removing the confusion and insecurity that impair student success Students are frequently intimidated by prep chem Bishop s text shows them how to break the material down and master it The flexible order of topics allows unit conversions to be covered either early in the course as is traditionally done or later allowing for a much earlier than usual description of elements compounds and chemical reactions The text and superb illustrations provide a solid conceptual framework and address misconceptions The book helps students to develop strategies for working problems in a series of logical steps The Examples and Exercises give plenty of confidence building practice the end of chapter problems test the student s mastery The system of objectives tells the students exactly what they must learn in each chapter and where to find it

**Chemistry** Steven S. Zumdahl, Susan A. Zumdahl, 2012 Steve and Susan Zumdahl s texts focus on helping students build critical thinking skills through the process of becoming independent problem solvers They help students learn to think like a chemists so they can apply the problem solving process to all aspects of their lives In *CHEMISTRY AN ATOMS FIRST APPROACH* 1e International Edition the Zumdahls use a meaningful approach that begins with the atom and proceeds through the concept of molecules structure and bonding to more complex materials and their properties Because this approach differs from what most students have experienced in high school courses it encourages them to focus on conceptual learning early in the course rather than relying on memorization and a plug and chug method of problem solving that even the best students can fall back on when confronted with familiar material The atoms first organization provides an opportunity for students to use the tools of critical thinkers to ask questions to apply rules and models and to

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### **Chemistry Review Sheet Unit 8 Answer Key Introduction**

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