

Introduction to Hazard Mitigation

Background:

Hand warmers are popular cold weather gear used to provide warmth to frigid fingers. Many commercial hand warmers consist of a plastic package containing a solid and an inner pouch filled with water. When the pack is activated, the solid dissolves in water and produces a large temperature change.

The energy or enthalpy change associated with a solute dissolving is called the heat of solution (ΔH_{sol}). At constant pressure, this enthalpy change is equal to the magnitude of heat loss or gain to the surroundings. In the case of an ionic solid dissolving in water the overall enthalpy change is the sum of three processes:

1. The energy required to break electrostatic forces in the ionic compound & separate the ions (ΔH_1),
2. The energy required to break the intermolecular forces between water molecules (ΔH_2), and
3. The energy released when the ions are attracted to the water molecules (ΔH_3)

The overall process can be represented by the following equation:

$$\Delta H_{sol} = \Delta H_1 + \Delta H_2 + \Delta H_3$$

If the amount of energy released when the ions are attracted to the water molecules (ΔH_3) is greater than ΔH_1 and ΔH_2 together, then the ΔH_{sol} will be negative and the solution process exothermic. If the amount of energy released when the ions are attracted to the water molecules (ΔH_3) is less than ΔH_1 and ΔH_2 together, then the ΔH_{sol} will be positive and the solution process endothermic.

Heats of solution and other enthalpy changes are generally measured in an insulated vessel called a calorimeter that reduces or prevents heat loss to the atmosphere. When using a calorimeter, the reagents being studied are mixed directly in the calorimeter and the temperature is recorded both before and after the reaction has occurred. The amount of heat transfer (q) may be calculated using the following equation:

$$q = m C_p \Delta T$$

where m is the total mass of the solution, C_p is the specific heat of the solution, and ΔT is the observed temperature change. The specific heat of the solution is generally assumed to be the same as water, 4.18 J/g°C.

Course:

The purpose of this lab is to design an effective heat-warmer that is inexpensive, nontoxic, and safe for the environment. Working in groups of four, each group will be provided nine different solids, along with their cost and Molecular Weight (MW) (Safety Data Sheet (SDS)). Determining the heat of solution for each solid and analyzing the cost and safety information to propose a design for the best all-around hand warmer.

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