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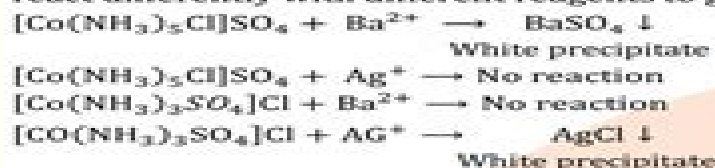
(Chapter 9) (Coordination Compound) (Intext Questions)  
(Class 12)

## Question 9.4:

Give evidence that  $[\text{Co}(\text{NH}_3)_5\text{Cl}]\text{SO}_4$  and  $[\text{Co}(\text{NH}_3)_5\text{SO}_4]\text{Cl}$  are ionization isomers.

### Answer 9.4:

When ionization isomers are dissolved in water, they ionize to give different ions. The ions then react differently with different reagents to give different products.



## Question 9.5:

Explain on the basis of valence bond theory that  $[\text{Ni}(\text{CN})_4]^{2-}$  ion with square planar structure is diamagnetic and the  $[\text{NiCl}_4]^{2-}$  ion with tetrahedral geometry is paramagnetic.

### Answer 9.5:

Ni is in the +2 oxidation state i.e., in  $d^8$  configuration.



There are 4  $\text{CN}^-$  ions. Thus, it can either have a tetrahedral geometry or square planar geometry. Since  $\text{CN}^-$  ion is a strong field ligand, it causes the pairing of unpaired 3d electrons.



It now undergoes  $dsp^2$  hybridization. Since all electrons are paired, it is diamagnetic. In case of  $[\text{NiCl}_4]^{2-}$ ,  $\text{Cl}^-$  ion is a weak field ligand. Therefore, it does not lead to the pairing of unpaired 3d electrons. Therefore, it undergoes  $sp^3$  hybridization.



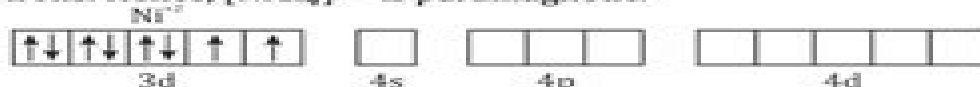
Since there are 2 unpaired electrons in this case, it is paramagnetic in nature.

## Question 9.6:

$[\text{NiCl}_4]^{2-}$  is paramagnetic while  $[\text{Ni}(\text{CO})_4]$  is diamagnetic though both are tetrahedral. Why?

### Answer 9.6:

Though both  $[\text{NiCl}_4]^{2-}$  and  $[\text{Ni}(\text{CO})_4]$  are tetrahedral, their magnetic characters are different. This is due to a difference in the nature of ligands.  $\text{Cl}^-$  is a weak field ligand and it does not cause the pairing of unpaired 3d electrons. Hence,  $[\text{NiCl}_4]^{2-}$  is paramagnetic.



In  $[\text{Ni}(\text{CO})_4]$ , Ni is in the zero oxidation state i.e., it has a configuration of  $3d^8 4s^2$ .

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# Chemistry 12 Intext Questions With Solution

**Prem Kumar**



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1. Understanding the eBook Chemistry 12 Intext Questions With Solution
  - The Rise of Digital Reading Chemistry 12 Intext Questions With Solution
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry 12 Intext Questions With Solution
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry 12 Intext Questions With Solution
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry 12 Intext Questions With Solution
  - Personalized Recommendations
  - Chemistry 12 Intext Questions With Solution User Reviews and Ratings
  - Chemistry 12 Intext Questions With Solution and Bestseller Lists

5. Accessing Chemistry 12 Intext Questions With Solution Free and Paid eBooks
  - Chemistry 12 Intext Questions With Solution Public Domain eBooks
  - Chemistry 12 Intext Questions With Solution eBook Subscription Services
  - Chemistry 12 Intext Questions With Solution Budget-Friendly Options
6. Navigating Chemistry 12 Intext Questions With Solution eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry 12 Intext Questions With Solution Compatibility with Devices
  - Chemistry 12 Intext Questions With Solution Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry 12 Intext Questions With Solution
  - Highlighting and Note-Taking Chemistry 12 Intext Questions With Solution
  - Interactive Elements Chemistry 12 Intext Questions With Solution
8. Staying Engaged with Chemistry 12 Intext Questions With Solution
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry 12 Intext Questions With Solution
9. Balancing eBooks and Physical Books Chemistry 12 Intext Questions With Solution
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry 12 Intext Questions With Solution
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry 12 Intext Questions With Solution
  - Setting Reading Goals Chemistry 12 Intext Questions With Solution
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry 12 Intext Questions With Solution
  - Fact-Checking eBook Content of Chemistry 12 Intext Questions With Solution
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning



- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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