

Name _____

Solution Stoichiometry Worksheet

Solve the following solutions Stoichiometry problems:

1. How many grams of silver chromate will precipitate when 150. mL of 0.500 M silver nitrate are added to 100. mL of 0.400 M potassium chromate?



$$\frac{0.150 \text{ L AgNO}_3}{1 \text{ L}} \left| \frac{0.500 \text{ moles AgNO}_3}{1 \text{ L}} \right| \left| \frac{1 \text{ moles Ag}_2\text{CrO}_4}{2 \text{ moles AgNO}_3} \right| \left| \frac{331.74 \text{ g Ag}_2\text{CrO}_4}{1 \text{ moles Ag}_2\text{CrO}_4} \right| = 12.4 \text{ g Ag}_2\text{CrO}_4$$

$$\frac{0.100 \text{ L K}_2\text{CrO}_4}{1 \text{ L}} \left| \frac{0.400 \text{ moles K}_2\text{CrO}_4}{1 \text{ L}} \right| \left| \frac{1 \text{ moles Ag}_2\text{CrO}_4}{1 \text{ moles K}_2\text{CrO}_4} \right| \left| \frac{331.74 \text{ g Ag}_2\text{CrO}_4}{1 \text{ moles Ag}_2\text{CrO}_4} \right| = 13.3 \text{ g Ag}_2\text{CrO}_4$$

2. How many mL of 0.280 M barium nitrate are required to precipitate as barium sulfate all the sulfate ions from 25.0 mL of 0.350 M aluminum sulfate? (93.8 mL barium nitrate)



$$\frac{0.0250 \text{ L Al}_2(\text{SO}_4)_3}{1 \text{ L}} \left| \frac{0.350 \text{ moles Al}_2(\text{SO}_4)_3}{1 \text{ L}} \right| \left| \frac{3 \text{ moles Ba}(\text{NO}_3)_2}{1 \text{ moles Al}_2(\text{SO}_4)_3} \right| \left| \frac{1 \text{ L}}{0.280 \text{ moles Ba}(\text{NO}_3)_2} \right| = 0.0938 \text{ L Ba}(\text{NO}_3)_2$$

3. 25.0 mL of 0.350 M NaOH are added to 45.0 mL of 0.125 M copper (II) sulfate. How many grams of copper (II) hydroxide will precipitate?



$$\frac{0.0250 \text{ L NaOH}}{1 \text{ L NaOH}} \left| \frac{0.350 \text{ moles NaOH}}{1 \text{ L NaOH}} \right| \left| \frac{1 \text{ moles Cu}(\text{OH})_2}{2 \text{ moles NaOH}} \right| \left| \frac{97.57 \text{ g Cu}(\text{OH})_2}{1 \text{ mole Cu}(\text{OH})_2} \right| = 0.427 \text{ g Cu}(\text{OH})_2$$

$$\frac{0.0450 \text{ L CuSO}_4}{1 \text{ L NaOH}} \left| \frac{0.125 \text{ moles CuSO}_4}{1 \text{ L NaOH}} \right| \left| \frac{1 \text{ moles Cu}(\text{OH})_2}{1 \text{ moles CuSO}_4} \right| \left| \frac{97.57 \text{ g Cu}(\text{OH})_2}{1 \text{ mole Cu}(\text{OH})_2} \right| = 0.549 \text{ g Cu}(\text{OH})_2$$

4. What volume of 0.415 M silver nitrate will be required to precipitate as silver bromide all the bromide ion in 35.0 mL of 0.128 M calcium bromide?



$$\frac{0.0350 \text{ L CaBr}_2}{1 \text{ L CaBr}_2} \left| \frac{0.128 \text{ moles CaBr}_2}{1 \text{ L CaBr}_2} \right| \left| \frac{2 \text{ moles AgNO}_3}{1 \text{ moles CaBr}_2} \right| \left| \frac{1 \text{ L AgNO}_3}{0.415 \text{ mole AgNO}_3} \right| = 0.0216 \text{ L AgNO}_3$$

5. What volume of 0.496 M HCl is required to neutralize 20.0 mL of 0.809 M sodium hydroxide?



$$\frac{0.0200 \text{ L NaOH}}{1 \text{ L NaOH}} \left| \frac{0.809 \text{ mole NaOH}}{1 \text{ L NaOH}} \right| \left| \frac{1 \text{ mole HCl}}{1 \text{ mole NaOH}} \right| \left| \frac{1 \text{ L HCl}}{0.496 \text{ mole HCl}} \right| = 0.0326 \text{ L HCl}$$

Stoichiometry Worksheet Answer Key

Penny Commons



Stoichiometry Worksheet Answer Key

Chemistry 2e Paul Flowers, Richard Langely, William R. Robinson, Klaus Hellmut Theopold, 2019-02-14 Chemistry 2e is designed to meet the scope and sequence requirements of the two semester general chemistry course The textbook provides an important opportunity for students to learn the core concepts of chemistry and understand how those concepts apply to their lives and the world around them The book also includes a number of innovative features including interactive exercises and real world applications designed to enhance student learning The second edition has been revised to incorporate clearer more current and more dynamic explanations while maintaining the same organization as the first edition Substantial improvements have been made in the figures illustrations and example exercises that support the text narrative Changes made in Chemistry 2e are described in the preface to help instructors transition to the second edition

Chemistry for the IB Diploma Workbook with CD-ROM Jacqueline Paris, 2017-04-06 Chemistry for the IB Diploma Second edition covers in full the requirements of the IB syllabus for Chemistry for first examination in 2016 This workbook is specifically for the IB Chemistry syllabus for examination from 2016 The Chemistry for the IB Diploma Workbook contains straightforward chapters that build learning in a gradual way first outlining key terms and then providing students with plenty of practice questions to apply their knowledge Each chapter concludes with exam style questions This structured approach reinforces learning and actively builds students confidence using key scientific skills handling data evaluating information and problem solving This helps empower students to become confident and independent learners Answers to all of the questions are on the CD ROM

Chemistry 2e Paul Flowers, Klaus Theopold, Richard Langley, Edward J. Neth, William R. Robinson, 2019-02-14 Chemistry 2e is designed to meet the scope and sequence requirements of the two semester general chemistry course The textbook provides an important opportunity for students to learn the core concepts of chemistry and understand how those concepts apply to their lives and the world around them The book also includes a number of innovative features including interactive exercises and real world applications designed to enhance student learning The second edition has been revised to incorporate clearer more current and more dynamic explanations while maintaining the same organization as the first edition Substantial improvements have been made in the figures illustrations and example exercises that support the text narrative Changes made in Chemistry 2e are described in the preface to help instructors transition to the second edition

Chemistry Theodore Lawrence Brown, H. Eugene LeMay, Bruce E. Bursten, Patrick Woodward, Catherine Murphy, 2017-01-03 NOTE This edition features the same content as the traditional text in a convenient three hole punched loose leaf version Books a la Carte also offer a great value this format costs significantly less than a new textbook Before purchasing check with your instructor or review your course syllabus to ensure that you select the correct ISBN Several versions of MyLab tm and Mastering tm platforms exist for each title including customized versions for

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STOICHIOMETRY AND PROCESS CALCULATIONS K. V. NARAYANAN, B. LAKSHMIKUTTY, 2006-01-01

This textbook is designed for undergraduate courses in chemical engineering and related disciplines such as biotechnology polymer technology petrochemical engineering electrochemical engineering environmental engineering safety engineering and industrial chemistry The chief objective of this text is to prepare students to make analysis of chemical processes through calculations and also to develop in them systematic problem solving skills The students are introduced not only to the application of law of combining proportions to chemical reactions as the word stoichiometry implies but also to formulating and solving material and energy balances in processes with and without chemical reactions The book presents

the fundamentals of chemical engineering operations and processes in an accessible style to help the students gain a thorough understanding of chemical process calculations It also covers in detail the background materials such as units and conversions dimensional analysis and dimensionless groups property estimation P V T behaviour of fluids vapour pressure and phase equilibrium relationships humidity and saturation With the help of examples the book explains the construction and use of reference substance plots equilibrium diagrams psychrometric charts steam tables and enthalpy composition diagrams It also elaborates on thermophysics and thermochemistry to acquaint the students with the thermodynamic principles of energy balance calculations Key Features SI units are used throughout the book Presents a thorough introduction to basic chemical engineering principles Provides many worked out examples and exercise problems with answers Objective type questions included at the end of the book serve as useful review material and also assist the students in preparing for competitive examinations such as GATE

Illustrated Guide to Home Chemistry Experiments Robert Bruce Thompson, 2012-02-17 For students DIY hobbyists and science buffs who can no longer get real chemistry sets this one of a kind guide explains how to set up and use a home chemistry lab with step by step instructions for conducting experiments in basic chemistry not just to make pretty colors and stinky smells but to learn how to do real lab work Purify alcohol by distillation Produce hydrogen and oxygen gas by electrolysis Smelt metallic copper from copper ore you make yourself Analyze the makeup of seawater bone and other common substances Synthesize oil of wintergreen from aspirin and rayon fiber from paper Perform forensics tests for fingerprints blood drugs and poisons and much more From the 1930s through the 1970s chemistry sets were among the most popular Christmas gifts selling in the millions But two decades ago real chemistry sets began to disappear as manufacturers and retailers became concerned about liability em The Illustrated Guide to Home Chemistry Experiments steps up to the plate with lessons on how to equip your home chemistry lab master laboratory skills and work safely in your lab The bulk of this book consists of 17 hands on chapters that include multiple laboratory sessions on the following topics Separating Mixtures Solubility and Solutions Colligative Properties of Solutions Introduction to Chemical Reactions Stoichiometry Reduction Oxidation Redox Reactions Acid Base Chemistry Chemical Kinetics Chemical Equilibrium and Le Chatelier s Principle Gas Chemistry Thermochemistry and Calorimetry Electrochemistry Photochemistry Colloids and Suspensions Qualitative Analysis Quantitative Analysis Synthesis of Useful Compounds Forensic Chemistry With plenty of full color illustrations and photos Illustrated Guide to Home Chemistry Experiments offers introductory level sessions suitable for a middle school or first year high school chemistry laboratory course and more advanced sessions suitable for students who intend to take the College Board Advanced Placement AP Chemistry exam A student who completes all of the laboratories in this book will have done the equivalent of two full years of high school chemistry lab work or a first year college general chemistry laboratory course This hands on introduction to real chemistry using real equipment real chemicals and real quantitative experiments is ideal for the many thousands of young

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General Chemistry Ralph H. Petrucci, F. Geoffrey Herring, Jeffrey D. Madura, Carey Bissonnette, 2010-05 Holt Chemistry R. Thomas Myers, 2006 *Modern Analytical Chemistry* David Harvey, 2000 This introductory text covers both traditional and contemporary topics relevant to analytical chemistry Its flexible approach allows instructors to choose their favourite topics of discussion from additional coverage of subjects such as sampling kinetic method and quality assurance

Chemistry Steven S. Zumdahl, Susan A. Zumdahl, 2012 Steve and Susan Zumdahl s texts focus on helping students build critical thinking skills through the process of becoming independent problem solvers They help students learn to think like a chemists so they can apply the problem solving process to all aspects of their lives In CHEMISTRY AN ATOMS FIRST APPROACH 1e International Edition the Zumdahls use a meaningful approach that begins with the atom and proceeds through the concept of molecules structure and bonding to more complex materials and their properties Because this approach differs from what most students have experienced in high school courses it encourages them to focus on conceptual learning early in the course rather than relying on memorization and a plug and chug method of problem solving that even the best students can fall back on when confronted with familiar material The atoms first organization provides an opportunity for students to use the tools of critical thinkers to ask questions to apply rules and models and to

Pearson Chemistry 11 New South Wales Skills and Assessment Book Elissa Huddart, 2017-11-30 The write in Skills and Assessment Activity Books focus on working scientifically skills and assessment They are designed to consolidate concepts learnt in class Students are also provided with regular opportunities for reflection and self evaluation throughout the book

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