

Chapter 11

Properties of Gases

Section 11.2 Gases and Pressure Changes Solutions for Practice Problems Student Edition page 514

Note: Assume that the temperature and amount of gas are constant in all of the following problems.

1. Practice Problem (page 514)

1.00 L of a gas at 1.00 atm pressure is compressed to 0.437 L. What is the new pressure of the gas?

What Is Required?

You need to calculate the final pressure, P_2 , on a sample of gas.

What Is Given?

You know the pressure and volume of the gas for the initial set of conditions and the volume for the final set of conditions:

$$P_1 = 1.00 \text{ atm}$$

$$V_1 = 1.00 \text{ L}$$

$$V_2 = 0.437 \text{ L}$$

Plan Your Strategy

The pressure and volume are changing at constant temperature. Use the equation for Boyle's law: $P_1V_1 = P_2V_2$

Divide each side of the equation by V_2 to isolate the variable P_2 .

Substitute the numbers and units for the known variables in the formula and solve for P_2 .

Act on Your Strategy

Isolation of the variable P_2 :

$$P_1V_1 = P_2V_2$$

$$\frac{P_1V_1}{V_2} = \frac{P_2\cancel{V_2}}{\cancel{V_2}}$$

$$P_2 = \frac{P_1V_1}{V_2}$$

Properties Of Gases Section Review Answer Key

Richard Bailey



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