

**Assignment #2 – Compound Names and Formulas**  
**Multi-valent ions only**

A. Write the correct formula for the following compounds, all of which have been named using the modern Roman Numeral Method. The ionic charge is given after the first element (metallic) in Roman Numerals.

- |                              |                                    |                            |                                    |
|------------------------------|------------------------------------|----------------------------|------------------------------------|
| 1. copper (II) oxide         | <u>CuO</u>                         | 11. manganese (III) oxide  | <u>Mn<sub>2</sub>O<sub>3</sub></u> |
| 2. mercury (I) oxide         | <u>Hg<sub>2</sub>O</u>             | 12. vanadium (II) bromide  | <u>VBr<sub>2</sub></u>             |
| 3. gold (III) chloride       | <u>AuCl<sub>3</sub></u>            | 13. niobium(V) oxide       | <u>Nb<sub>2</sub>O<sub>5</sub></u> |
| 4. thallium (III) bromide    | <u>TlBr<sub>3</sub></u>            | 14. titanium (III) oxide   | <u>Ti<sub>2</sub>O<sub>3</sub></u> |
| 5. bismuth (V) oxide         | <u>Bi<sub>2</sub>O<sub>5</sub></u> | 15. titanium (III) nitride | <u>TiN</u>                         |
| 6. terbium (III) oxide       | <u>Tb<sub>2</sub>O<sub>3</sub></u> | 16. iron (II) oxide        | <u>FeO</u>                         |
| 7. uranium (VI) oxide        | <u>UO<sub>3</sub></u>              | 17. cobalt (II) phosphide  | <u>Co<sub>3</sub>P<sub>2</sub></u> |
| 8. protactinium (IV) bromide | <u>PaBr<sub>4</sub></u>            | 18. tin (II) oxide         | <u>SnO</u>                         |
| 9. cerium (III) oxide        | <u>Ce<sub>2</sub>O<sub>3</sub></u> | 19. thulium (II) bromide   | <u>TmBr<sub>2</sub></u>            |
| 10. arsenic (V) sulphide     | <u>As<sub>2</sub>S<sub>5</sub></u> | 20. copper (II) bromide    | <u>CuBr<sub>2</sub></u>            |

B. Name the compound using the Roman Numeral Method and balance the ionic charges.

- |                                   |                             |                                    |                                |
|-----------------------------------|-----------------------------|------------------------------------|--------------------------------|
| 1. SnCl <sub>4</sub>              | <u>Tin (IV) Chloride</u>    | 9. PdF <sub>4</sub>                | <u>Palladium (IV) Fluoride</u> |
| 2. BiBr <sub>5</sub>              | <u>Bismuth (V) Bromide</u>  | 10. Os <sub>2</sub> O <sub>3</sub> | <u>Osmium (III) Oxide</u>      |
| 3. PoO <sub>2</sub>               | <u>Polonium (IV) Oxide</u>  | 11. MoBr <sub>2</sub>              | <u>Molybdenum (II) Bromide</u> |
| 4. PbI <sub>2</sub>               | <u>Lead (II) Iodide</u>     | 12. VCl <sub>5</sub>               | <u>Vanadium (V) Chloride</u>   |
| 5. HgO                            | <u>Mercury (II) Oxide</u>   | 13. Mn <sub>2</sub> O <sub>3</sub> | <u>Manganese (III) Oxide</u>   |
| 6. HgCl                           | <u>Mercury (I) Chloride</u> | 14. CoO                            | <u>Cobalt (II) Oxide</u>       |
| 7. Au <sub>2</sub> O <sub>3</sub> | <u>Gold (III) Oxide</u>     | 15. Np <sub>2</sub> O <sub>3</sub> | <u>Neptunium (III) Oxide</u>   |
| 8. FeCl <sub>2</sub>              | <u>Iron (II) Chloride</u>   | 16. V <sub>2</sub> O <sub>5</sub>  | <u>Vanadium (V) Oxide</u>      |

# **Predicting And Naming Polyatomic Ionic Compounds Answers**

**Spencer L. Seager, Michael R.  
Slabaugh**



## Predicting And Naming Polyatomic Ionic Compounds Answers

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