

Chemistry Unit 8 Worksheet 4

Samples of Every Kind of Problem

On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedure outlined in class. Be sure to circle your final answer.

1. Calculate the number of moles of potassium chlorate, KClO_3 (s), that must decompose to produce potassium chloride, KCl (s), and 1.8 moles of oxygen gas.

| | | | |
|-----------|---|------|------|
| Equation: | $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$ | | |
| Before: | 1.2 | 0 | 0 |
| Change: | -1.2 | +1.2 | +1.8 |
| After: | 0 | 1.2 | 1.8 |

1.2 moles
 KClO_3 needed

2. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 2.43 g of magnesium?

| | | | |
|-----------|---|--------|--------|
| Equation: | $\text{Mg} + 2 \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$ | | |
| Before: | 0.100 | ? | 0 |
| Change: | -0.100 | -0.200 | +0.100 |
| After: | 0 | ? | 0.100 |

0.200 moles
 HCl needed

3. Ethane, C_2H_6 reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.20 moles of ethane?

| | | | |
|-----------|--|-------|-------|
| Equation: | $2 \text{C}_2\text{H}_6 + 7 \text{O}_2 \rightarrow 4 \text{CO}_2 + 6 \text{H}_2\text{O}$ | | |
| Before: | 2.20 | xs | 0 |
| Change: | -2.20 | -7.70 | +4.40 |
| After: | 0 | xs | 4.40 |

$$\text{O}_2 \text{ needed: } 7.70 \text{ mole} \times \frac{32.0 \text{ g}}{1 \text{ mole}} = 246 \text{ g O}_2 \text{ required}$$

4. Determine the mass of sodium nitrate produced when 0.73 g of nickel (II) nitrate reacts with sodium hydroxide according to the following equation:



$$0.73 \text{ g} \times \frac{1 \text{ mole}}{183 \text{ g}} = .0040 \text{ mole}$$

| | | | |
|-----------|---|---------|---------|
| Equation: | $\text{Ni}(\text{NO}_3)_2 + 2 \text{NaOH} \rightarrow \text{Ni}(\text{OH})_2 + 2 \text{NaNO}_3$ | | |
| Before: | 0.0040 | NS | 0 |
| Change: | -0.0040 | -0.0080 | +0.0040 |
| After: | 0 | NS | 0.0080 |

$$\text{NaNO}_3 \text{ produced: } 0.0080 \text{ mole} \times \frac{85.0 \text{ g}}{1 \text{ mole}} = 0.68 \text{ g}$$

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**Antony C. Wilbraham, Dennis D.
Staley, Michael S. Matta, Edward L.
Waterman**

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Prentice Hall Chemistry Antony C. Wilbraham, 2007

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Student Solutions Manual for General Chemistry John McMurry, Robert Fay, 2013-03-07 By Joseph Topich Virginia Commonwealth University This manual for students contains solutions to selected all in chapter problems and even numbered end of chapter problems

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Student Solution Manual for Introduction to Chemical Principles H. Stephen Stoker, Nancy J. Gardner, 2013-04-22 The Student Solutions Manual includes full solutions to all odd numbered end of chapter problems in the text and answers to all multiple choice practice test questions

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Chemistry 2e Paul Flowers, Richard Langely, William R. Robinson, Klaus Hellmut Theopold, 2019-02-14 Chemistry 2e is designed to meet the scope and sequence requirements of the two semester general chemistry course The textbook provides an important opportunity for students to learn the core concepts of chemistry and understand how those concepts apply to their lives and the world around them The book also includes a number of innovative features including interactive exercises and real world applications designed to enhance student learning The second edition has been revised to incorporate clearer more current and more dynamic explanations while maintaining the same organization as the first edition Substantial improvements have been made in the figures illustrations and example exercises that support the text narrative Changes made in Chemistry 2e are described in the preface to help instructors transition to the second edition

The Pearson Guide To Physical Chemistry For The Aipmt Singhal Atul, 2011-09

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Chemistry O Level Mauritius Edition Pearson Education, Limited, 2007-06-26

Holt McDougal Modern Chemistry Mickey Sarquis, 2012

The Chemistry Maths Book Erich Steiner, 1996 The Chemistry Maths Book is a comprehensive textbook of mathematics for undergraduate students of chemistry Such students often find themselves unprepared and ill equipped to deal with the mathematical content of their chemistry courses Textbooks designed to overcome this problem have so far been too basic for complete undergraduate courses and have been unpopular with students However this modern textbook provides a complete and up to date course companion suitable for all levels of undergraduate chemistry courses All the most useful and important topics are covered with numerous examples of applications in chemistry and some in physics The subject is developed in a logical and consistent way with few assumptions of prior knowledge of mathematics This text is sure to become a widely adopted text and will be highly recommended for all chemistry courses

Chemistry Insights Rex M. Heyworth, J. G. R. Briggs, 2008

Students Study Guide Fredeen, Pearson, 2003-07

Chemical Analysis and Material Characterization by Spectrophotometry Bhim Prasad Kafle, 2019-11-29 Chemical Analysis and Material Characterization by Spectrophotometry integrates and presents the latest known information and examples from the most up to date literature on the use of this method for chemical analysis or materials characterization Accessible to various levels of expertise everyone from students to practicing analytical and industrial chemists the book covers both the fundamentals of spectrophotometry and instrumental procedures for quantitative analysis with spectrophotometric techniques It contains a wealth of examples and focuses on the latest research such as the investigation of optical properties of nanomaterials and thin solid films Covers the basic analytical theory that is essential for understanding spectrophotometry Emphasizes minor trace chemical component analysis Includes the spectrophotometric analysis of nanomaterials and thin solid films Thoroughly describes methods and uses easy to follow practical examples and experiments

Introductory Chemical Engineering Thermodynamics J. Richard Elliott, Carl T. Lira, 2012 In this book two leading experts and long time instructors thoroughly explain thermodynamics taking the molecular perspective that working engineers require This edition contains extensive new coverage of today's fast growing biochemical engineering applications notably biomass conversion to fuels and chemicals It also presents many new MATLAB examples and tools to complement its previous usage of Excel and other software

The Mechanisms of Fast Reactions in Solution Edward Caldin, 2001 Published a few years after the author's death this volume is a sequel to his 1964 book *Fast Reactions in Solution* the material is entirely new extending investigation beyond now well established fast reaction techniques to consider their contribution to understanding events on the molecular scale After an introductory chapter on origins methods mechanisms and rate constants coverage includes the rates of diffusion controlled reactions mathematical theory of diffusion flash photolysis techniques fluorescence quenching Marcus theory involving proton transfer and group transfer reactions in solutions and electron transfer reactions Annotation copyrighted by Book News Inc Portland OR

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