

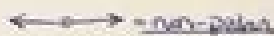
H. 59. Alpha particles with their high ionizing charge but low penetrating ability can best be shielded by?

- a. paper      b. clothing      c. water      d. lead

Alpha  $\alpha$  particles - paper is a sufficient shield

Beta  $\beta$  particles - heavy clothing is sufficient to shield

Gamma rays - lead is required to shield



D. 60. The half-life of Se-121 is 10 days. If you started with 40 grams of this isotope, how much would you have left after 30 days?

- a. 10 g      b. none      c. 15 g      d. 5 g

Mass of Se-121	Number of half-lives	Time
40 g	0	0
20 g	1	10 days
10 g	2	20 days
5 g	3	30 days

- $\oplus \uparrow$  = particles go faster  
 $\oplus \downarrow$  = particles go slower

**PERIODIC TRENDS**

ATOMIC RADIUS: decreased  $\rightarrow$   
 increased  $\downarrow$

METALLIC CHARACTER:

decreases  $\downarrow$   
 increases  $\rightarrow$

IONIZATION  $\oplus$  decreases  $\downarrow$   
 increases  $\rightarrow$



Isotopes = atoms with same # of protons, but different # of neutrons.

EXAMPLE: Ne-20  $\rightarrow$   $^{20}_{10}\text{Ne}$   
 Ne-21  $\rightarrow$   $^{21}_{10}\text{Ne}$   
 Ne-22  $\rightarrow$   $^{22}_{10}\text{Ne}$

Dilution Equation  $M_1 V_1 = M_2 V_2$

•  $\text{mass \%} = \frac{\text{mass solute}}{\text{mass of solution (solute + solvent)}} \times 100\%$

•  $\text{Molarity (M)} = \frac{\text{mol solute}}{\text{L solution}}$

•  $\text{molality (m)} = \frac{\text{mol solute}}{\text{kg solvent}}$

- #protons - #electrons = Charge #
- #neutrons =  $A$  (mass #) -  $Z$  (atomic #)
- if no charge, the #protons + #electrons = same
- if + (positive) charge (cation) = #electrons is less than #protons
- if - (negative) charge (anion) = #electrons is greater than #protons
- Protons  $\neq$  Charge  $\rightarrow$  they give identity to the element.

# **Acs High School Chemistry Exam Study Guide**

**Thomas A. Holme, Kristen Murphy**



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**ACING the ACS** James Coy-Dibley,2017-09-05 Study guide for first year general chemistry ACS Exam

**Biochemistry Education** Assistant Teaching Professor Department of Chemistry and Biochemistry Thomas J Bussey, Timothy J. Bussey, Kimberly Linenberger Cortes, Rodney C. Austin, 2021-01-18 This volume brings together resources from the networks and communities that contribute to biochemistry education Projects authors and practitioners from the American Chemical Society ACS American Society of Biochemistry and Molecular Biology ASBMB and the Society for the Advancement of Biology Education Research SABER are included to facilitate cross talk among these communities Authors offer diverse perspectives on pedagogy and chapters focus on topics such as the development of visual literacy pedagogies and practices and implementation

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**Reagent Chemicals** American Chemical Society, 2015 The American Chemical Society ACS Committee on Analytical Reagents sets the specifications for most chemicals used in analytical testing Currently the ACS is the only organization in the world that sets requirements and develops validated methods for determining the purity of reagent chemicals These specifications have also become the de facto standards for chemicals used in many high purity applications Publications and organizations that set specifications or promulgate analytical testing methods such as the United States Pharmacopeia and the U S Environmental Protection Agency specify that ACS reagent grade purity be used in their test procedures The Eleventh Edition incorporates the supplements accumulated over the past eight years removes some obsolete test methods improves instructions for many existing ones and also introduces some new methods Overall the safety accuracy or ease of use in specifications for about 70 of the 430 listed reagents has been improved and seven new reagents have been added

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