

# 16 • Chemical Equilibrium

## PRACTICE TEST

1. Consider the reaction system,  
 $\text{CoO}(s) + \text{H}_2(g) \rightleftharpoons \text{Co}(s) + \text{H}_2\text{O}(g)$  *[Products]*  
*[Reactants]*

The equilibrium constant expression is

- a)  $\frac{[\text{CoO}][\text{H}_2]}{[\text{Co}][\text{H}_2\text{O}]}$       d)  $\frac{[\text{H}_2]}{[\text{H}_2\text{O}]}$  *[H<sub>2</sub>O]*  
 b)  $\frac{[\text{Co}][\text{H}_2\text{O}]}{[\text{CoO}][\text{H}_2]}$       **c)  $\frac{[\text{H}_2\text{O}]}{[\text{H}_2]}$**  *[H<sub>2</sub>]*  
 c)  $\frac{[\text{CoO}][\text{H}_2]}{[\text{Co}][\text{H}_2\text{O}]}$   
     [H<sub>2</sub>]

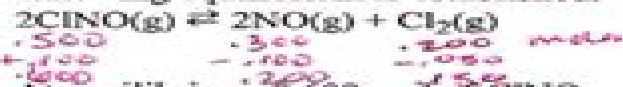
2. Given the equilibrium,  
 $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$ , if this *as soon as*  
 equilibrium is established by beginning with *even 1 m*  
 equal number of moles of SO<sub>2</sub> and O<sub>2</sub> in a *occurs [SO<sub>2</sub>]*  
*< [O<sub>2</sub>]*

1.0 Liter bulb, then the following must be true at equilibrium:

- a) [SO<sub>2</sub>] = [SO<sub>3</sub>]      **d) [SO<sub>2</sub>] < [O<sub>2</sub>]**  
 b) 2[SO<sub>2</sub>] = 2[SO<sub>3</sub>]      e) [SO<sub>2</sub>] > [O<sub>2</sub>]  
 c) [SO<sub>2</sub>] = [O<sub>2</sub>]

Questions 3 & 4 refer to the following:

At a given temperature, 0.300 mole NO, 0.200 mol Cl<sub>2</sub> and 0.500 mol ClNO were placed in a 25.0 Liter container. The following equilibrium is established:



3. At equilibrium, 0.600 mol of ClNO was present. The number of moles of Cl<sub>2</sub> present at equilibrium is

- a) 0.050      d) 0.200  
 b) 0.100      e) 0.250  
**c) 0.150**

4. The equilibrium constant, K<sub>c</sub>, is:

- a)  $4.45 \times 10^{-4}$       d) 0.167  
**b)  $6.67 \times 10^{-4}$**       e) 1500  
 c) 0.111

$$K_c = \frac{\left(\frac{.200}{25}\right)^2 \left(\frac{.150}{25}\right)}{\left(\frac{.600}{25}\right)^2}$$

5. At 985°C, the equilibrium constant for the reaction,



is 1.63. What is the equilibrium constant for the reverse reaction? *K' = 1/k = 1/1.63*

- a) 1.63      **d) 0.613**  
 b) 0.815      e) 1.00  
 c) 2.66

6. What is the relationship between K<sub>p</sub> and K<sub>c</sub> for the reaction, 2ICl(g) ⇌ I<sub>2</sub>(g) + Cl<sub>2</sub>(g)?

- a) K<sub>p</sub> = K<sub>c</sub>(RT)<sup>-1</sup>      **d) K<sub>p</sub> = K<sub>c</sub> Δn=0**  
 b) K<sub>p</sub> = K<sub>c</sub>(RT)      e) K<sub>p</sub> = K<sub>c</sub>(2RT)  
 c) K<sub>p</sub> = K<sub>c</sub>(RT)<sup>2</sup>

7. For the reaction  $2\text{NO}_2(g) \rightleftharpoons \text{N}_2\text{O}_4(g)$ , K<sub>p</sub> at 25°C is 7.3, when all partial pressures are expressed in atmospheres. What is K<sub>c</sub> for this reaction? *Δn = 1 - 2 = -1*  
*298 K* [R=0.0821 L·atm·mol<sup>-1</sup>·K<sup>-1</sup>]

- a) 4270      **d) 179** *K<sub>p</sub> = K<sub>c</sub>(RT)<sup>Δn</sup>*  
 b) 0.0119      e) 2.06 *K<sub>c</sub> = K<sub>p</sub> / (RT)<sup>Δn</sup>*  
 c) 0.291

8. 0.200 mol NO is placed in a one liter flask at 2273 K. After equilibrium is attained, 0.0863 mol N<sub>2</sub> and 0.0863 mol O<sub>2</sub> are present. What is K<sub>c</sub> for this reaction?



- a) 9.92**      d) 39.7  
 b) 3.15      e) 0.576  
 c) 0.0372

.200 M	0	0
-2(.0863)	+.0863	+.0863
.0274	.0863	.0863

$$K_c = \frac{(0.0863)(0.0863)}{(0.0274)^2} = 9.92$$

# Workbook Answers For Chemical Equilibrium

**DJ Losen**



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