



K_c CALCULATIONS 1

answer out to 2 or 3 sig figs
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Simplest questions – equilibrium quantities given

- 1) Calculate K_c for the following equilibrium given the results of an analysis of the equilibrium mixture which show that the concentration of SO₂ = 0.230, O₂ = 1.370, and SO₃ = 0.800 mol dm⁻³.



Answer: $4.479 \text{ dm}^3 \text{ mol}^{-1}$

- 2) Calculate K_c for the equilibrium below if [A] = 0.0050 mol dm⁻³, [B] = 0.100 mol dm⁻³ and [AB] = 0.400 mol dm⁻³.



Answer: $900 \text{ mol}^{-1} \text{ dm}^3$

- 3) Some PCl₅ was heated in a sealed container resulting in the following equilibrium. Analysis of the equilibrium mixture showed that it contained 0.00420 moles of PCl₅, 0.0400 moles of PCl₃ and 0.0400 moles of Cl₂. The volume was 2.00 dm³. Calculate the concentration of each species at equilibrium and then K_c.



Answer: $0.1705 \text{ mol dm}^{-3}$

Next step up – you need to find equilibrium quantities first

- 4) 2.00 moles of PCl₅ vapour are heated to temperature T in a vessel of volume $2.0 \times 10^2 \text{ dm}^3$. The equilibrium mixture contains 1.20 moles of chlorine. Calculate K_c for the equilibrium at temperature T.



Answer: 0.09 mol dm^{-3}

- 5) When 5.0 moles of A is mixed with 5.0 moles of B in a container of volume 12 dm³, an equilibrium is established which contains 3.0 moles of C. Find K_c at this temperature.



Answer: $0.9 \text{ mol}^{-1} \text{ dm}^3$

- 6) 200.0 g of ethyl ethanoate and 7.00 g of water were refluxed together. At equilibrium, the mixture contained 0.250 mol of ethanoic acid. Calculate K_c for the reaction of ethyl ethanoate.



Answer: 4.49

Working back from K_c to find equilibrium quantities

- 7) In the following equilibrium, K_c = 54.1 at a particular temperature. The equilibrium mixture was found to contain H₂ at a concentration of $4.80 \times 10^{-2} \text{ mol dm}^{-3}$ and I₂ at a concentration of $3.50 \times 10^{-2} \text{ mol dm}^{-3}$.



- a) What is the equilibrium concentration of HI? $4.70 \times 10^{-2} \text{ mol dm}^{-3}$
- b) What effect would doubling the concentration of hydrogen have on the equilibrium position and K_c? moves to right, no effect.
- c) What effect would increasing the temperature have on the equilibrium position and K_c?
- 8) K_c for the equilibrium below is 0.00960 at temperature T. If 1.00 mole of H₂ and 1.00 mole of O₂ are allowed to reach equilibrium, what mass of H₂O will be present in the mixture (you do not need to solve a quadratic equation to answer this)?



Answer: 1.8g

Something harder

- 9) The following equilibrium is established at temperature T when 1.00 mol of HI is contained. It is found that 22% of the HI has broken down. Calculate K_c. Error: volume not mentioned. Assume it 1 dm³.



Answer: 0.02

- 10) When 1.00 moles of steam and 1.00 moles of carbon monoxide were allowed to reach equilibrium at 30% of the equilibrium volume is hydrogen. Calculate K_c at this temperature (you will need to think very carefully about equilibrium quantities).



Answer: 0.33

- 11) K_c for the equilibrium below is 10.0 at temperature T. If 1.0 mole of the ester is mixed with 2.0 moles of water and all mixture allowed to reach equilibrium, how many moles of each species will be present at equilibrium (you will need to solve a quadratic equation to answer this)?



Answer: $\frac{10-5\sqrt{5}}{3}$

Kc Calculations 1 Chemsheets

E. N. Ramsden



Kc Calculations 1 Chemsheets

An Introduction to Chemical Kinetics Claire Vallance, 2017-09-28 The book is a short primer on chemical reaction rates based on a six lecture first year undergraduate course taught by the author at the University of Oxford The book explores the various factors that determine how fast or slowly a chemical reaction proceeds and describes a variety of experimental methods for measuring reaction rates The link between the reaction rate and the sequence of steps that makes up the reaction mechanism is also investigated Chemical reaction rates is a core topic in all undergraduate chemistry courses

Calculations in AS/A Level Chemistry Jim Clark, 2000-01-01 Suitable for all examination specifications for students over 16 this friendly and reliable guide leads students through examples of each problem

Medicinal Chemistry and Pharmacological Potential of Fullerenes and Carbon Nanotubes Franco Cataldo, Tatiana da Ros, 2008-09-08 Fullerenes and nanotubes are two classes of carbon structures or allotropes which were discovered about 17 years ago Since that time many chemical derivatives have been synthesized using fullerenes and nanotubes as building blocks Particularly promising was the theory that the chemical properties of fullerenes and certain derivatives made them likely candidates for anticancer drugs inhibitors of viruses such as HIV or even as anti bacterials Their cytotoxicity can also be controlled by specific circumstances In addition the functionalization of nanotubes has not only produced relatively simple derivatives but also complex hybrids with biological macromolecules which show unique supramolecular architecture and which are promising in many medical applications The application of fullerenes and nanotubes in medicine is at the frontier of our knowledge thus the work in this field represents the basis for future novel developments

Longman GCSE Chemistry Jim Clark, 2002 One of a series of books designed for middle to higher ability students following the revised GCSE Double and Triple Award specifications Exam style questions and practice questions aim to facilitate progression from GCSE to AS S level

A-level Chemistry E. N. Ramsden, 2000 This textbook has been updated to cover the new specifications for AS and A2 Chemistry and improved with new features and rewritten material to enhance learning and increase accessibility It covers all the main specifications for the English and Welsh Awarding Bodies and should be particularly suitable for students approaching A Level from GCSE Science Double Award This answer key is designed to support the core book and contains suggested answers worked solutions to the checkpoints and examination questions in the core book also synoptic questions for further practice complete with suggested answers and worked solutions to help develop confidence

Edexcel IGCSE Chemistry Jim Clark, 2009 Offers complete coverage of the new specification Includes free student ActiveBook CD ROM Links to additional support and teacher support are provided online directly from Edexcel

Organic Analysis John Mitchell, Izaak Maurits Kolthoff, Erich S. Proskauer, 1953

The Chemistry Maths Book Erich Steiner, 1996 The Chemistry Maths Book is a comprehensive textbook of mathematics for undergraduate students of chemistry Such students often find themselves

unprepared and ill equipped to deal with the mathematical content of their chemistry courses Textbooks designed to overcome this problem have so far been too basic for complete undergraduate courses and have been unpopular with students However this modern textbook provides a complete and up to date course companion suitable for all levels of undergraduate chemistry courses All the most useful and important topics are covered with numerous examples of applications in chemistry and some in physics The subject is developed in a logical and consistent way with few assumptions of prior knowledge of mathematics This text is sure to become a widely adopted text and will be highly recommended for all chemistry courses

Life's Other Secret Ian Stewart,1998 Until the middle of this century it was completely unclear whether life had any kind of inorganic basis The discovery of the first secret of life the molecular structure of DNA solved that particular riddle

Maths Skills for a Level Chemistry Dan McGowan,Emma Poole,2013-04-24 This handbook specifically targets the mathematical elements of A Level Science whichever specification you re following Includes plenty of practice questions in different contexts to increase confidence worked examples and model answers for revision and exam preparation Plus hints and tips for the exam and how to avoid common errors made in mathematical science questions

Chemistry in Context - Laboratory Manual Graham Hill,John Stranger Holman,John Holman,2001 The laboratory manual and study guide supports your teaching with a broad range of practicals emphasising safety and risk assessment It is an essential companion to Chemistry in Context and can also be used alongside other Advanced Chemistry books It offers practicals with detailed instructions for openended investigations and opportunities for assessed practical work in the four skill areas of planning implementing analysing and evaluating

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Chemistry Data Book J. G. Stark,H. G. Wallace,1982 This text is a standard reference book for A Level and equivalent examinations

Advanced Practical Chemistry Ann Lainchbury,Alec Thompson,1997 An advanced level course of laboratory practical work in chemistry derived from the second edition of ILPAC books 1 12 revised to reflect the syllabus requirements The 95 experiments in this book cover all aspects of Advanced level chemistry For each the student is provided with aims a brief introduction equipment requirements detailed procedures explanations of how to treat the results and a series of follow up questions Where necessary there is guidance on the calculations

Modern Chemical Techniques C. B. Faust,1992-01-01 This book is based on a series of symposia that enabled individuals to update their chemical skills and learn about the newest methods techniques and instrumentation available

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