

Stoichiometry Worksheet # 1

1) Glucose is used as a source of energy by the human body. The overall reaction in the body is



Calculate the number of moles of oxygen needed to oxidize 12.5 g of $\text{C}_6\text{H}_{12}\text{O}_6$.

2) Ammonia is synthesized from hydrogen and nitrogen according to the following equation.



If an excess of nitrogen is reacted with 3.41 moles of hydrogen gas, how many grams of ammonia (NH_3) can be produced?

3) Assume that in the decomposition of potassium chlorate, KClO_3 ,



80.5 moles of O_2 form. How many grams of potassium chlorite, the other product would be formed?

4) In a single displacement reaction,



9.23 moles of aluminum react with excess hydrochloric acid. How many moles of hydrogen will be produced?

5) The compound "Cisplatin" $\text{PtCl}_2(\text{NH}_3)_2$, has been found to be effective in treating some types of cancer. It can be synthesized using the following reaction



A. How many moles of "cisplatin" can be produced from 2.50 g K_2PtCl_4 ?

B. How many moles of NH_3 would be needed if 3.8 moles of K_2PtCl_4 were present?

Chapter 12 Stoichiometry Test B Answers

Kaplan Test Prep



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Chemistry 2e Paul Flowers, Richard Langely, William R. Robinson, Klaus Hellmut Theopold, 2019-02-14 Chemistry 2e is designed to meet the scope and sequence requirements of the two semester general chemistry course. The textbook provides an important opportunity for students to learn the core concepts of chemistry and understand how those concepts apply to their lives and the world around them. The book also includes a number of innovative features including interactive exercises and real world applications designed to enhance student learning. The second edition has been revised to incorporate clearer, more current and more dynamic explanations while maintaining the same organization as the first edition. Substantial improvements have been made in the figures, illustrations, and example exercises that support the text narrative. Changes made in Chemistry 2e are described in the preface to help instructors transition to the second edition.

Cracking the SAT Chemistry Subject Test Princeton Review (Firm), 2014-12-09 EVERYTHING YOU NEED TO HELP SCORE A PERFECT 800 Equip yourself to ace the SAT Chemistry Subject Test with The Princeton Review's comprehensive study guide including 3 full length practice tests, thorough reviews of key chemistry topics, and targeted strategies for every question type. We don't have to tell you how tough SAT Chemistry is or how helpful a stellar exam score can be for your chances of getting into your top choice college. Written by the experts at The Princeton Review, Cracking the SAT Chemistry Subject Test arms you to take on the test and achieve your highest score. Techniques That Actually Work: Tried and true strategies to help you avoid traps and beat the test. Tips for pacing yourself and guessing logically. Essential tactics to help you work smarter, not harder. Everything You Need to Know for a High Score: Expert subject reviews for every test topic. Up to date information on the SAT Chemistry Subject Test. Score conversion tables for accurate self assessment. Practice Your Way to Perfection: 3 full length practice tests with detailed answer explanations. Hands on experience with all three question types in each content chapter. Complete study sheet of core formulas and terms.

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- Chapter 1: Introduction to Chapter 12 Stoichiometry Test B Answers
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- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Chapter 12 Stoichiometry Test B Answers. This chapter will explore what Chapter 12 Stoichiometry Test B Answers is, why Chapter 12 Stoichiometry Test B Answers is vital, and how to effectively learn about Chapter 12 Stoichiometry Test B Answers.

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